

SYNTHESIS OF FLUORINE-18 LABELED 1,1-DIFLUORO-2,2-DICHLOROETHYL ARYL ETHERS BY ^{18}F -FOR- ^{19}F EXCHANGE

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SUMMARY

Fluorine-18 labeled 1,1-difluoro-2,2-dichloroethyl aryl ethers have been prepared by a facile ^{18}F -for- ^{19}F isotopic exchange reaction. The isotopic exchange proceeds in good to excellent yields (up to 85%), and is dependent on reaction time, temperature, and concentration of reactants (aryl ether or fluoride ion). Substitution even occurs at room temperature in DMSO, and at elevated temperature (100 to 155 °C) in DMSO or acetonitrile containing up to 20% water. Specific activities of 0.5-55 Ci/mol have been achieved, with the products obtained in pure radiochemical and chemical form.

KEY WORDS: [^{18}F]fluoride ion, isotopic exchange,
nucleophilic substitution, perfluoroethers

INTRODUCTION

[^{18}F]Fluoride ion has become a readily available, widely used precursor for radiopharmaceutical synthesis, and has been applied to numerous aliphatic and aromatic nucleophilic substitution reactions (1,2). As part of our interests in labeling of organofluorine compounds with fluorine-18, we report here the simple,

unprecedented synthesis of carrier-added fluorine-18 labeled 1,1-difluoro-2,2-dichloroethyl aryl ethers by a remarkably facile ^{18}F -for- ^{19}F exchange reaction.

EXPERIMENTAL

Materials and Methods. No-carrier-added [^{18}F]fluoride ion was produced by proton irradiation of 95-97% oxygen-18 enriched water (Isotec) held in an all-silver cyclotron target (3). Acetonitrile and dimethylsulfoxide were obtained in anhydrous form (Aldrich Chem. Co.) and used without further purification. The 1,1-difluoro-2,2-dichloroethyl aryl ethers (1-4) were prepared by condensation of the appropriate sodium phenoxide with 1,1-difluoro-2,2-dichloroethylene under phase transfer conditions (4,5). TLC was done using plastic-backed silica gel plates (Merck) and one of two solvent systems: System A = hexane; System B = 7/3 hexane/ethylacetate. HPLC was done using a Phenomenex C18 column, 0.45 x 15 cm, 60:40 CH_3CN :0.065 M NH_4OAc , 1 ml/min: detection by UV (264 nm) and radioactivity (Beckman Model 170 flow detector).

General Procedure for [^{18}F]Fluorinations. Aqueous [^{18}F]fluoride ion was combined with potassium carbonate (1.8 mg) and Kryptofix-2,2,2 (10 mg) and the water evaporated (100 °C, N_2 flow). Acetonitrile was added and evaporated to remove traces of water. The residue was dissolved in 1 ml of anhydrous solvent, and 100 μl aliquots used in subsequent [^{18}F]fluorination reactions. The [^{18}F]fluorinations were performed by adding 100 μl of a solution of ether (1-4) to the [^{18}F]fluoride solution, then heating (various temperatures and reaction times: see Table 1). Reactions were quenched by addition of 2 ml water, and the aqueous mixture extracted with 4 ml of diethyl ether or 4/1 diethyl ether/ethylacetate. The organic layers were then analyzed by TLC and the yields of fluoride exchange (radioactivity in organic phase/total radioactivity) corrected for [^{18}F]fluoride ion (if any) extracted into the organic solution. Identity and purity (chemical and radiochemical) were also determined by injection of the isolated products on HPLC. In some instances, aliquots from reactions done in acetonitrile were directly spotted on TLC plates, or were diluted with 60:40 CH_3CN : H_2O and directly injected on the HPLC.

The R_f values (TLC) and R_t times (HPLC) for the four ethers were as follows:

(1): TLC, A, $R_f = 0.28$; B $R_f = 0.61$: HPLC, $R_t = 13.9$ min.

(2): TLC, A, $R_f = 0.27$; B $R_f = 0.43$: HPLC, $R_t = 14.0$ min.

(3): TLC, A, $R_f = 0.05$; B $R_f = 0.37$: HPLC, $R_t = 12.8$ min.

(4): TLC A, $R_f = 0.05$; B $R_f = 0.43$: HPLC, $R_t = 12.9$ min.

For some reactions (see Table 1), the [¹⁸F]fluoride ion was prepared as the tetrabutylammonium fluoride salt (substitution of TBAF for K₂CO₃/Kryptofix 2,2,2) or as a polymer-bound species, using an aminopyridine-substituted polystyrene polymer as previously described (6).

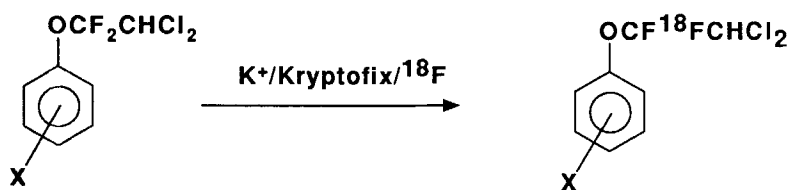
RESULTS AND DISCUSSION

The reaction of [¹⁸F]fluoride ion (as the soluble K⁺/Kryptofix salt, tetrabutylammonium salt, or resin-bound) with 1,1-difluoro-2,2-dichloroethyl aryl ethers 1-4 proceeds with ¹⁸F-for-¹⁹F substitution to give moderate to excellent yields of the [¹⁸F]ethers (Table 1). The isotopic exchange reaction is dependent on temperature, length of reaction, concentration of substrate (aryl ether), and substituent on the aryl ring. Highest yields, and greatest reactivity, was exhibited by the 3-nitrophenyl ether 3, with lowest yields obtained with the 4-fluorophenyl ether 2.

Dipolar aprotic solvents appear most useful for this exchange reaction. Addition of non-polar solvents, such as chloroform or mesitylene, reduce [¹⁸F]fluorination yields. Most surprisingly, however, were the nearly 50% yields obtained in solutions of DMSO or acetonitrile containing up to 20% (v/v) of water. Exchange of [¹⁸F]fluoride with 3 also takes place in the absence of any solvent, with the yield (44%, 30 min, 155 °C) perhaps limited by the solubility (unknown) of the K[¹⁸F]F/Kryptofix reagent in the neat ether solution. The exchange reaction is not dependent on the nature of the cation (K⁺, tetrabutylammonium, or resin) as essentially identical results were obtained with all three sources of [¹⁸F]fluoride ion.

The remarkable ease of ¹⁸F-for-¹⁹F exchange in these 1,1-difluoro-2,2-dichloroethyl aryl ethers was surprising. Although isotopic fluorine exchange is preceded with both alkyl and aryl fluorides, previous attempts to effect isotopic fluorine exchanges in perfluoroalkyl compounds have met limited success and have

Table 1. Yields of 1,1-[¹⁸F]difluoro-2,2,-dichloroethyl aryl ethers obtained by isotopic fluorine exchange. Yields are decay-corrected.



X	¹⁹ F (mM)	conc (mM)	solvent	T, °C	time (min)	yield (%)
H (1)	-	7	CH ₃ CN	100	30	12
	-	7	DMSO	155	30	45
	-	70	DMSO	155	30	55,57
	-	70	10% H ₂ O in DMSO	155	30	55
	-	70	20% H ₂ O in DMSO	155	30	51
4-Fluoro (2)	-	7	CH ₃ CN	100	30	2.5
	-	7	DMSO	155	30	25
	-	70	DMSO	155	30	23
3-Nitro (3)	-	66	CH ₃ CN	25	30	5
	-	6.6	CH ₃ CN	100	30	70
	-	66	CH ₃ CN	100	10	55*
	-	66	CH ₃ CN	100	15	80
	-	0.0066	DMSO	155	30	<5
	-	0.066	DMSO	155	30	31
	-	0.66	DMSO	155	30	57,46
	-	6.6	DMSO	155	5	41
	-				10	59
	-				15	67
	-				30	71
	-	66	DMSO	155	30	85,82
	10	6.6	DMSO	155	15	44**
	100	6.6	DMSO	155	30	17**
	-	6.6	DMSO	100	30	53
-	6.6	2/1 DMSO/CHCl ₃	100	30	31	
-	6.6	2/1 DMSO/CH ₃ CN	100	30	78	
-	6.6	2/1 DMSO/mesitylene	100	30	27	
3-Nitro (3)	-	66	10% H ₂ O in DMSO	155	30	43
	-	66	20% H ₂ O in DMSO	155	75	47
	-	66	10% H ₂ O in CH ₃ CN	100	30	49

Table 1, cont'd.

X	¹⁹ F (mM)	conc (mM)	solvent	T, °C	time (min)	yield (%)
	-	66	20% H ₂ O in CH ₃ CN	100	30	12
4-Nitro (4)	-	6.6	CH ₃ CN	100	30	31
	-	6.6	DMSO	155	5	23
	-				10	27
	-				15	34
	-	66	DMSO	155	30	30
	-	66	DMSO	155	30	55

* [¹⁸F]fluoride prepared in polymer-supported form**[¹⁸F]fluoride prepared as tetrabutylammonium salt

required high reaction temperatures (7,8,9). The isotopic exchange reported here occurs even at room temperature, albeit in low yield. Assignment of the products as due to isotopic fluorine exchange reaction, rather than ring substitutions of nitro or fluoro groups, was based on the reactivity of the unsubstituted aryl ether **1**, as well as TLC and HPLC analysis showing the labeled product from reaction of 4-nitrophenyl ether **4** was identical to the unlabeled ether, and quite separate from the alternative product involving ring fluorination (compound **3**). Perfluoroalkyl ethers have been previously examined for effects on ring reactivity, and have been classified as halogen-like (10), and thus would not be expected to be activating towards nucleophilic substitutions of phenyl rings. Fluorine-18 substitution for the fluorine, rather than one of the chlorines, follows from the reactivity of α -fluoroalkyl ethers towards AlCl₃ to produce α,α,α -trichloroperfluoroethers (11), substitution of 1,1-difluoro-2,2-dichloroethyl aryl ethers by bromide ion (AlBr₃) to yield 1,1-dibromo-2,2-dichloroethyl ethers (4), and nucleophilic attack by sodium ethylate on ethoxy-1,1-difluoro-2,2-dichloroethane to give the orthoester (substitution of ethoxide for fluoride, rather than chloride (12)). [¹⁸F]Fluoride ion substitution of a chlorine is preceded with alkyl chlorides (13), but the presence of the two fluorine substituents on the alpha position would be expected to significantly decrease the reactivity of the chlorines toward simple nucleophilic substitution. In all cases the radiolabeled products were identical chromatographically to the starting materials, and we could not identify any products which might have arisen by F-for-Cl substitution, nor any products which might arise by rearrangement reactions.

Although the 1,1-difluoro-2,2-dichloroethyl ethers are subject to both rearrangement and hydrolysis under strongly basic reaction conditions (4,5), we did not find chemical or radiochemical by-products which would support such reactions under the conditions utilized here for the isotopic exchange. Our preparations of [^{18}F]fluoride ion in organic solvents (using $\text{K}_2\text{CO}_3/\text{Kryptofix}$, tetrabutylammonium hydroxide, or anion exchange resin) are all quite suitable for use with substrates which are sensitive to basic conditions, as evidenced by good yields in the fluorination of the triflate precursor to 2-deoxy-2- [^{18}F]fluoro-D-glucose.

The reactions described here yield, of course, the labeled aryl ethers in carrier-added form, with specific activities in the range of 0.5 - 55 Ci/mol. Most experiments in Table 1 were done with small amounts (100 μCi or less) of [^{18}F]fluoride ion, although we have prepared multimillicurie amounts of the compounds **1** and **3** in fluorine-18 form. The use of high initial levels of [^{18}F]fluoride ion, coupled with low amounts of aryl ether (0.1 μmol) might lead to fluorine-18 labeled products of higher (50-100 Ci/mmol) specific activity. The products are obtained in very high chemical purity (i.e., no decomposition or side reactions) and can be easily separated from unreacted [^{18}F]fluoride ion (silica gel column). 1,1-Difluoro-2,2-dichloroethyl ethers are interesting chemical structures. Recently investigated as phenol protecting groups, they are remarkably stable to a wide variety of reaction conditions and reagents (4) The possible uses of such aryl ethers to prepare radiopharmaceuticals (for those cases where high specific activity is not necessary) are under investigation.

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