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Carbon(sp³)–Fluorine Bond-Forming Reductive Elimination from Palladium(IV) Complexes**

Joy M. Racowski, J. Brannon Gary, and Melanie S. Sanford*

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General Procedures

NMR spectra were obtained on a Varian vnmrs 700 (699.76 MHz for ¹H; 175.95 MHz for ¹³C), a Varian Inova 400 (399.96 MHz for ¹H; 376.34 MHz for ¹⁹F; 100.57 MHz for ¹³C), a Varian vnmr500 (500.09 MHz for ¹H; 470.56 MHz for ¹⁹F; 125.75 MHz for ¹³C), or a Varion MR400 (400.53 MHz for ¹H; 376.87 MHz for ¹⁹F; 100.71 MHz for ¹³C) spectrometer. ¹H, ¹⁹F and ¹³C chemical shifts are reported in parts per million (ppm) relative to TMS, with the residual solvent peak used as an internal reference. ¹⁹F NMR spectra are referenced on a unified scale, where the single primary reference is the frequency of the residual solvent peak in the ¹H NMR spectrum.¹ ¹H and ¹⁹F multiplicities are reported as follows: singlet (s), doublet (d), doublet of doublets (dd), quartet (q), multiplet (m), and broad resonance (br). Mass spectral data were obtained on a Micromass magnetic sector mass spectrometer or on a Micromass LCT mass spectrometer in electrospray ionization mode.

Materials and Methods

Bipyridine (bpy) and 2-methyl-2-phenylpropyl magnesium chloride were obtained from Aldrich. 1-Fluoro-2,4,6-trimethylpyridinium triflate (NFTPT) and 1-fluoro-2,4,6-trimethylpyridinium tetrafluoroborate (NFTPB) were obtained from TCI America. Unless otherwise noted, all reagents were used as received. NMR solvents were obtained from Cambridge Isotope Laboratories. All other solvents were obtained from Fisher Chemicals. Tetrahydrofuran was purified using an Innovative Technologies (IT) solvent purification system consisting of a copper catalyst, activated alumina, and molecular sieves.

Synthesis of Pd^{II} Complex 1



Pd^{II}(CH₂CMe₂-*o*-C₆H₄)(COD)² (720 mg, 2.08, 1.0 equiv) was combined with 2,2'-bipyridine (325 mg, 2.08 mmol, 1.0 equiv) in CH₂Cl₂ (200 mL), and the reaction mixture allowed to stir for 30 min. The solution was concentrated under vacuum (to 5 mL), and hexanes (30 mL) was added to precipitate the product. Complex **1** was isolated as a bright yellow solid (737 mg, 90% yield). ¹H NMR (500 MHz [D₃]chloroform, 25 °C): δ = 9.25 (d, *J* = 5 Hz, 1H), 8.80 (d, *J* = 5 Hz, 1H), 8.07-8.05 (multiple peaks, 2H), 7.99-7.93 (multiple peaks, 2H), 7.56-7.53 (multiple peaks, 2H), 7.47 (t, *J* = 7 Hz, 1H), 7.01-6.99 (multiple peaks, 2H), 6.87 (t, *J* = 5 Hz, 1H), 2.46 (s, 2H), 1.43 (s, 6H). ¹³C NMR (125 MHz [D₃]chloroform, 25 °C): δ = 168.71, 158.54, 154.73, 154.21, 150.43, 149.37, 137.11, 134.35, 128.18, 125.42, 125.13, 123.72, 122.76, 121.34, 121.04, 121.00, 46.96, 44.97, 33.35 (two overlapping carbons). HRMS-electrospray (m/z): [M + H]⁺ calcd for C₂₀H₂₀N₂Pd, 395.0734; Found, 395.0746.

Synthesis of Pd^{IV} Complexes 2-5



Compound **1** (70 mg, 0.18 mmol, 1.0 equiv) and NFTPT (52 mg, 0.18 mmol, 1.0 equiv) were combined in CH_2Cl_2 (15 mL), and this mixture was stirred for 15 min. The solvent was removed by rotary evaporation, and the resulting yellow oil was washed with diethyl ether (15 mL). The solid material was then dissolved in CH_2Cl_2 (1 mL), and diethyl ether (10 mL) was added to precipitate the product. The precipitate was collected and dried under vacuum to afford **2** as a light yellow solid (96 mg, 94% yield). ¹H NMR (400 MHz [D₃]acetonitrile, 25 °C): δ = 8.96 (d, *J* = 5 Hz, 1H), 8.58-8.56 (multiple peaks, 2H), 8.43 (t, *J* = 8 Hz, 1H), 8.34 (t, *J* = 8 Hz, 1H), 8.03-7.99 (multiple peaks, 2H), 7.78 (d, *J* = 8 Hz, 1H), 7.61 (t, *J* = 6 Hz, 1H), 7.34 (d, *J* = 8 Hz, 1H), 7.29 (d, *J* = 8 Hz, 1H), 7.13 (d, *J* = 8 Hz, 1H), 4.80 (dd, *J* = 15, 5 Hz, 1H), 4.24 (app. br. s, 1H), 1.45 (s, 3H), 1.12 (s, 3H). ¹⁹F NMR (376 MHz [D₃]acetonitrile, 25 °C): δ = -79.11 (s, 3F), -336.17 (d, *J* = 15 Hz, 1F). ¹³C NMR data could not be obtained due to the instability of the complex over the timescale required for the experiment. HRMS-ESI (m/z): [M – OTf]⁺ calcd for C₂₀H₂₀FN₂Pd 413.0640; Found, 413.0644.



Compound **1** (120 mg, 0.3 mmol, 1.0 equiv), NFTPT (88 mg, 0.3 mmol, 1.0 equiv), and $C_{5}H_{5}N$ (49 µL, 0.6 mmol, 2.0 equiv) were combined in CH₂Cl₂ (15 mL), and this mixture was stirred for 15 min. The solvent was removed by rotary evaporation, and the resulting yellow oil was washed with diethyl ether (10 mL). The solid material was then dissolved in CH₂Cl₂ (2 mL), and diethyl ether (25 mL) was added to precipitate the product. The precipitate was collected and dried under vacuum to afford **3** as an off-white solid (125 mg, 63% yield). ¹H NMR (700 MHz [D₂]dichloromethane, 25 °C): δ = 8.93-8.91 (multiple peaks, 2H), 8.72 (d, *J* = 8 Hz, 1H), 8.48 (t, *J* = 8 Hz, 1H), 8.29 (t, *J* = 8 Hz, 1H), 8.24 (app. br. s, 2H), 8.07 (d, *J* = 6 Hz, 1H), 7.88 (t, *J* = 8 Hz, 1H), 7.77 (t, *J* = 6 Hz, 1H), 7.64 (t, *J* = 6 Hz, 1H), 7.47-7.44 (multiple peaks, 2H), 7.24 (t, *J* = 7 Hz, 1H), 7.05 (d, *J* = 8 Hz, 1H), 7.01 (t, *J* = 8 Hz, 1H), 6.66 (d, *J* = 8 Hz, 1H), 4.72 (dd, *J* = 15, 6 Hz, 1H), 3.89 (app. br. s, 1H), 1.50 (s, 3H), 1.13 (s, 3H). ¹⁹F NMR (376 MHz [D₂]dichloromethane, 25 °C):: δ = -78.97 (s, 3F), -324.80 (d, *J* = 15 Hz, 1F). ¹³C NMR data could not be obtained due to the instability of the complex over the timescale required for the experiment. HRMS-ESI (m/z): [M + H]⁺ calcd for C₂₆H₂₄F₄N₃O₃PdS 642.0660; Found, 642.0669.



Compound 2 (11 mg, 0.02 mmol) was dissolved in acetone (1.5 mL) in a 3 mL vial. The vial was then placed in a 20 mL vial containing pentane and sealed with a Teflon-lined cap. The vial was placed in a – 35 °C freezer until yellow crystals formed. The solvent was decanted, and the crystals were washed with pentanes (3 mL), CH_2CI_2 (3 mL), and acetone (3 mL). The crystals were then collected and dried under vacuum to afford 4 as a bright yellow solid (5.2 mg, 53% yield). ¹H NMR (500 MHz [D₆]dimethyl sulfoxide, 25 °C): δ = 9.04 (d, *J* = 5 Hz, 1H), 8.91-8.89 (multiple peaks, 2H), 8.56 (t, *J* = 8 Hz, 1H), 8.47 (t, *J* = 8 Hz, 1H), 8.10 (t, *J* = 6 Hz, 1H), 7.86 (d, *J* = 7 Hz, 1H), 7.78 (t, *J* = 6 Hz, 1H), 7.72 (d, *J* = 7 Hz, 1H), 7.31 (d, *J* = 8 Hz, 1H), 7.27 (d, *J* = 8 Hz, 1H), 7.14 (d, *J* = 7 Hz, 1H), 4.61 (dd, *J* = 15, 5 Hz, 1H), 4.23 (d, *J* = 5 Hz, 1H), 3.31 (broad s) 1.45 (s, 3H), 1.13 (s, 3H). ¹⁹F NMR (376 MHz [D₆]dimethyl sulfoxide, 25 °C): δ = – 77.91 (s, 3F), –328.96 (d, *J* = 15 Hz, 1F). ¹³C NMR data could not be obtained due to the instability of the complex over the timescale required for the experiment. HRMS-ESI (m/z): [M – H₂O – OTf]⁺ calcd for $C_{20}H_{20}FN_2Pd$ 413.0640; Found, 413.0656.



Compound **1** (552 mg, 1.4 mmol, 1.0 equiv), NFTPB (259 mg, 1.4 mmol, 1.0 equiv), and C_5H_5N (225 µL, 2.8 mmol, 2.0 equiv) were combined in CH₂Cl₂ (50 mL), and this mixture was stirred for 15 min. The solvent was removed by rotary evaporation, and the resulting yellow oil was washed with diethyl ether (20 mL). The solid material was then dissolved in CH₂Cl₂ (3 mL), and diethyl ether (25 mL) was added to precipitate the product. The precipitate was collected and dried under vacuum to afford **3-BF**₄ as an off-white solid (486 mg, 58% yield). ¹H NMR (400 MHz [D₂]dichloromethane, 25 °C): δ = 8.91 (d, *J* = 5 Hz, 1H), 8.79 (d, *J* = 8 Hz, 1H), 8.62 (d, *J* = 8 Hz, 1H), 8.47 (d, *J* = 8 Hz, 1H), 8.29 (d, *J* = 8 Hz, 1H), 8.23 (app. br. s, 2H), 8.07 (d, *J* = 6 Hz, 1H), 7.05 (d, *J* = 7 Hz, 1H), 7.01 (t, *J* = 8 Hz, 1H), 6.65 (d, *J* = 8 Hz, 1H), 4.72 (dd, *J* = 15, 6 Hz, 1H), 3.88 (app. br. s, 1H), 1.49 (s, 3H), 1.13 (s, 3H). ¹⁹F NMR (376 MHz [D₂]dichloromethane, 25 °C): δ = -152.48 (s, 4F), -324.80 (d, *J* = 15 Hz, 1F). ¹³C NMR data could not be obtained due to the instability of the complex over the timescale required for the experiment. HRMS-ESI (m/z): [M – pyridine – BF₄]⁺ calcd for C₂₀H₂₀FN₂Pd 413.0640; Found, 413.0645.



Compound **2** (310 mg, 0.54 mmol, 1.0 equiv) and Me₄NF (52 mg, 0.86 mmol, 1.6 equiv) were combined in dry CH₂Cl₂ (15 mL), and this mixture was stirred for 15 min in the glovebox. The solution turned dark orange with a dark solid precipitate. The reaction was filtered through celite, and the solvent was removed under vacuum. The resulting yellow oil was dissolved in DCE (4 mL), and pentane (30 mL) was added to precipitate the product. The precipitate was collected and dried under vacuum to afford **5** as a light yellow solid (190 mg, 77% yield, along with 19 % Me₄NBF₄ as determined by ¹H NMR). ¹H NMR (500 MHz [D₂]dichloromethane, 25 °C): $\delta = \delta$ 9.06 (d, J = 5 Hz, 1H), 8.43 (d, J = 8 Hz, 1H), 8.40 (d, J = 8 Hz, 1H), 8.14-8.08 (multiple peaks, 2H), 8.01 (t, J = 8 Hz, 1H), 7.95 (d, J = 8 Hz, 1H), 7.74 (d, J = 6 Hz, 1H), 7.31 (t, J = 6 Hz, 1H), 7.24-7.16 (multiple peaks, 2H), 6.95 (d, J = 8 Hz, 1H), 4.11 (m, 1H), 3.51 (dd, J = 6, 3 Hz, 1H), 1.39 (s, 3H), 1.02 (s, 3H). ¹⁹F NMR (470 MHz [D₂]dichloromethane, 25 °C): $\delta = -201.42$ (d, J =51 Hz, 1F), -336.71 (dd, J = 51, 14 Hz, 1F). HRMS-ESI (m/z): [M – F]⁺ calcd for C₂₀H₂₀FN₂Pd 413.0645; Found, 413.0640.

Synthesis of Pd^{II} Reductive Elimination Products



Compound **3** (30 mg, 0.06 mmol) was dissolved CH₂Cl₂ (10 mL). The reaction was stirred for 30 min at 80 °C. The solvent was removed by rotary evaporation, and the resulting yellow oil was dissolved in CH₂Cl₂ (1 mL), and precipitated with pentane (15 mL). The precipitate was collected and dried under vacuum to afford **6** as a tacky yellow solid (28 mg, 93% yield). ¹H NMR (500 MHz [D₃]chloroform, 25 °C): δ = 8.88 (d, *J* = 8 Hz, 2H), 8.56 (d, *J* = 8 Hz, 1H), 8.53 (d, *J* = 8 Hz, 1H), 8.22 (t, *J* = 8 Hz, 1H), 8.16 (t, *J* = 8 Hz, 1H), 8.04-7.99 (multiple peaks, 2H), 7.77 (d, *J* = 5 Hz, 1H), 7.68-7.61 (multiple peaks, 2H), 7.56 (d, *J* = 6 Hz, 1H), 7.32 (t, *J* = 6 Hz, 1H), 7.22 (t, *J* = 8 Hz, 1H), 7.11-7.08 (multiple peaks, 2H), 4.69 (dd, *J* = 48, 9 Hz, 1H), 4.53 (dd, *J* = 48, 9 Hz, 1F). ¹³C NMR (125 MHz [D₃]chloroform, 25 °C): δ = 156.40, 153.70, 152.45, 152.25, 150.41, 148.29, 148.13, 141.10, 140.82, 139.70, 132.91, 128.34, 128.07, 126.94, 126.86, 126.48 (q, *J* = 322 Hz, 1C), 126.38, 125.27, 124.04, 124.01, 92.21 (d, *J* = 177 Hz, 1C), 40.91, 40.76, 27.60. HRMS-ESI (m/z): [M - C₅H₅N - OTf]⁺ calcd for C₂₀H₂₀FN₂Pd 413.0645; Found, 413.0643.



Compound **7** (20 mg, 0.04 mmol, 1.0 equiv) was dissolved in dry CH₂Cl₂ (6 mL) under N₂. The reaction was stirred for 15 min at 80 °C. The solvent was removed by rotary evaporation, the resulting yellow oil was dissolved in CH₂Cl₂ (1 mL), and the product was precipitated with pentane (15 mL). The precipitate was collected and dried under vacuum to afford **8** as a tacky yellow solid (18 mg, 90% yield). ¹H NMR (700 MHz [D₂]dichloromethane, 25 °C): δ = 9.16 (d, *J* = 5 Hz, 1H), 8.75 (d, *J* = 5 Hz, 1H), 8.15-8.13 (multiple peaks, 2H), 8.06-8.01 (multiple peaks, 2H), 7.61 (t, *J* = 5 Hz, 1H), 7.53 (d, *J* = 6 Hz, 1H), 7.51 (d, *J* = 8 Hz, 1H), 6.99 (t, *J* = 7 Hz, 1H), 6.95 (t, *J* = 8 Hz, 1H), 6.86 (d, *J* = 7 Hz, 1H), 4.54 (dd, *J* = 49, 9 Hz, 1H), 4.51 (dd, *J* = 49, 9 Hz, 1H), 1.39 (s, 3H), 1.38 (s, 3H). ¹⁹F NMR (376 MHz [D₂]dichloromethane, 25 °C): δ = -151.89 (s, 1F), -214.87 (t, *J* = 49 Hz, 1F). ¹³C NMR (175 MHz [D₃]chlorodorm, 25 °C): δ = 163.45, 160.98, 155.59, 155.19, 151.02, 150.11, 138.42, 138.37, 135.29, 126.46, 126.24, 125.22, 123.63, 123.25, 122.27, 122.20, 95.39 (d, *J* = 176 Hz, 1C), 36.62, 36.60, 25.63. HRMS-APCI (m/z): [M – F]⁺ calcd for C₂₀H₂₀FN₂Pd 413.0640; Found, 413.0641.



Compound **3-BF**₄ (30 mg, 0.05 mmol) dissolved combined in CH₂Cl₂ (10 mL). The reaction was stirred for 30 min at 80 °C. The solvent was removed by rotary evaporation, and the resulting yellow oil was dissolved in CH₂Cl₂ (1 mL), and precipitated with pentane (15 mL). The precipitate was collected and dried under vacuum to afford **6-BF**₄ as a tacky yellow solid (26 mg, 87% yield). ¹H NMR (700 MHz [D₂]dichloromethane, 25 °C): δ = 8.85-8.83 (multiple peaks, 2H), 8.48-8.45 (multiple peaks, 2H), 8.16-8.11 (multiple peaks, 2H), 8.01-7.96 (multiple peaks, 2H), 7.75 (d, *J* = 5 Hz, 1H), 7.63 (d, *J* = 7 Hz, 1H), 7.62-7.58 (multiple peaks, 2H), 7.52 (d, *J* = 6 Hz, 1H), 7.75 (d, *J* = 6 Hz, 1H), 7.19 (d, *J* = 8 Hz, 1H), 7.08-7.05 (multiple peaks, 2H), 4.68 (dd, *J* = 48 Hz, *J* = 9 Hz, 1H), 4.50 (dd, *J* = 48 Hz, *J* = 9 Hz, 1H), 1.64 (s, 3H), 1.40 (s, 3H). ¹⁹F NMR (376 MHz [D₃]chloroform, 25 °C): δ = 156.03, 153.23, 151.87 (two overlapping carbon's), 150.16, 147.91, 140.67, 140.49, 139.35, 132.58, 127.94, 127.74, 126.60 (two overlapping carbon's), 126.42, 126.00, 124.87, 123.55, 123.47, 93.31 (d, *J* = 176 Hz, 1C), 40.53, 40.40, 27.19. HRMS-ESI (m/z): [M - C₅H₅N - BF₄]⁺ calcd for C₂₀H₂₀FN₂Pd 413.0645; Found, 413.0655.

Determining Order in Pyridine with 3-BF₄ at 45 °C in CD₂Cl₂

Complex **3-BF**₄ (4.4 mg, 0.00758 mmol, 1.0 equiv) and C₅D₅N (0.001895 to 0.0114 mmol, 3.8 mM to 20.0 mM) were combined in a screw cap NMR tube and dissolved in CD₂Cl₂ (0.5 mL). An internal standard (2-nitrobenzotrifluoride) was added (20 μ l of a stock solution in CD₂Cl₂, 0.00758, 1 equiv) and the tube was sealed with a Teflon®-lined cap. The tube was immediately placed in an NMR spectrometer with the temperature pre-equilbrated to 45 °C, and the reaction was allowed to equilibrate for 2 min. The rate of reductive elimination was monitored by ¹⁹F NMR spectroscopy by monitoring the disappearance of the starting material. The reaction was followed to between 2-3 half lives, and the data was plotted as –ln[**3-BF**₄/**3-BF**₄°] versus time. A representative kinetics run is shown in **Figure S1**. The values of *k*_{obs} for each [pyridine] are reported in **Table S1**.

Table S1. Rate as a Function of $[C_5D_5N]$ at 45 °C



equiv C₅D₅N	[C₅D₅N]	[1/C₅D₅N]	k obs
0.25	0.001895	528	1.60 x 10 ⁻⁴
0.5	0.00379	264	2.75 x 10 ⁻⁴
0.75	0.00568	176	5.50 x 10 ⁻⁵
1	0.00758	132	8.02 x 10 ⁻⁵
1.5	0.0114	88	2.81 x 10 ⁻⁵

Figure S1. Representative Rate Data (Reductive Elimination from $3-BF_4$ in the Presence of 5.68 mM C_5D_5N)



Study of the Appearance of 7 at 40 °C in CD₂CI₂

Complex **5** (4.9 mg, 0.0113 mmol, 1.0 equiv) and NMe₄NF (2.5 mg, 0.0269 mmol, 2.4 equiv) or NBu₄PF₆ (10.4 mg, 0.0269 mmol, 2.4 equiv) were combined in a screw cap NMR tube and dissolved in CD_2CI_2 (0.5 mL) from a stock solution that contained the internal standard dichloroethane (0.0057 mmol, 0.5 equiv). The tube was sealed with a Teflon®-lined cap. The tube was immediately placed in an NMR spectrometer with the temperature pre-equilbrated to 40 °C and the reaction was allowed to equilibrate for 2 minutes. The rate of reductive elimination was monitored by ¹H NMR spectroscopy by monitoring the appearance of compound **7**.

Additive	k _{obs}	
none	5.75 x 10 ⁻⁴	
NMe ₄ F	3.26 x10 ⁻⁴	
NBu ₄ PF ₆	3.30 x10 ⁻³	

X-ray Crystallography Details

Yellow cubes of **4** were grown by diffusion of pentane into a solution of **3** in wet acetone at -35 °C. A crystal of dimensions 0.10 x 0.10 x 0.10 mm was mounted on a Bruker SMART APEX CCD-based X-ray diffractometer equipped with a low temperature device and fine focus Mo-target X-ray tube (I = 0.71073 A) operated at 1500 W power (50 kV, 30 mA). The X-ray intensities were measured at 85(1) K; the detector was placed at a distance 5.055 cm from the crystal. A total of 4717 frames were collected with a scan width of 0.5° in ω and 0.45° in phi with an exposure time of 30 s/frame. The integration of the data yielded a total of 119486 reflections to a maximum 20 value of 56.64° of which 7429 were independent and 7186 were greater than $2\sigma(I)$. The final cell constants (Table 1) were based on the xyz centroids of 9866 reflections above $10\sigma(I)$. Analysis of the data showed negligible decay during data collection; the data were processed with SADABS and corrected for absorption. The structure was solved and refined with the Bruker SHELXTL (version 2008/4) software package, using the space group Pca2(1) with Z = 4 for the formula $C_{33}H_{26}BN_3F_4Cl_2Pd$. All non-hydrogen atoms were refined anisotropically with the hydrogen atoms placed in idealized positions. Full matrix least-squares refinement based on Φ^2 converged at R1 = 0.0219 and wR2 = 0.0533 [based on I > 2sigma(I)], R1 = 0.0232 and wR2 = 0.0542 for all data. Additional details are presented in **Table S2** and are available in the corresponding CIF file (deposited in the Cambridge Structural Database: CCDC 852596).

Sheldrick, G.M. SHELXTL, v. 2008/4; Bruker Analytical X-ray, Madison, WI, 2008.

Saint Plus, v. 7.60A, Bruker Analytical X-ray, Madison, WI, 2009.

Sheldrick, G.M. SADABS, v. 2008/1. Program for Empirical Absorption Correction of Area Detector Data, University of Gottingen: Gottingen, Germany, 2008.

 Table S2. Crystal data and structure refinement for 5.

Empirical formula	C21 H22 F4 N2 O4 Pd S
Formula weight	580.87
Temperature	85(2) K
Wavelength	1.54178 A
Crystal system, space group	Triclinic, P-1
Unit cell dimensions	a = 9.3652(2) A alpha = 94.404(7) deg.
	b = 11.5602(3) A beta = 108.848(8) deg.
	c = 11.9928(8) A gamma = 111.654(8) deg.
Volume	1113.33(8) A^3
Z, Calculated density	2, 1.733 Mg/m^3
Absorption coefficient	8.196 mm^-1
F(000)	584
Crystal size	0.18 x 0.12 x 0.09 mm
Theta range for data collection	4.00 to 68.17 deg.
Limiting indices	-11<=h<=11, -13<=k<=13, -11<=l<=12
Reflections collected / unique	16877 / 3453 [R(int) = 0.0504]
Completeness to theta = 68.17	84.7 % (96% complete to 0.86 A).
Absorption correction	Semi-empirical from equivalents
Max. and min. transmission	0.686 and 0.557
Refinement method	Full-matrix least-squares on F^2
Data / restraints / parameters	3453 / 0 / 375
Goodness-of-fit on F^2	1.261
Final R indices [I>2sigma(I)]	R1 = 0.0688, wR2 = 0.2137
R indices (all data)	R1 = 0.0699, wR2 = 0.2141
Largest diff. peak and hole	2.236 and -1.268 e.A^-3

Computational Methods

Using the Gaussian 09 suite of programs,³ all density functional theory (DFT) calculations were performed with the M06 functional⁴ along with the Stevens (CEP-31G) valence basis sets with effective core potentials.^{5,6} The CEP-31G basis sets are triple- ξ for Pd and double- ξ for all main group elements. A d-polarization function (see 6-31G^{*})^{7,8} was added to all non-hydrogen main group elements: $\xi_d = 0.8$ for carbon, nitrogen, oxygen, and fluorine (referred to as CEP-31G(d) level of theory). All geometries were optimized using CEP-31G(d)/M06 without symmetry constraints using the restricted Kohn-Sham formalism for all complexes.. All minima were confirmed by the absence of imaginary frequencies and all transition states were verified by visual inspection of the single imaginary frequency vibration and optimized along the reaction coordinate. Thermochemical data was calculated using unscaled vibrational frequencies and default parameters at 298.15 K and 1 atm. Solvent corrections (CH₂Cl₂) were performed as single point calculations using the SMD model⁹ with default settings at the optimized geometries from gas phase calculations.

Energy Comparisons

Calculated energies of the intermediates and transition states calculated are shown below. All energies have been solvent corrected using dichloromethane. All energies are listed in kcal/mol and complex **8** has been set to 0.0 for reference.



(complex rearranges such that this is the same as 8 sp³ ts)

Complex	Electronic (SCF)	Enthalpy	Free Energy
8	0.0	0.0	0.0
8-I	3.2	3.2	2.3
8 sp ³ ts	17.0	15.2	16.5
8 sp ² ts	29.5	28.0	27.0
8-l sp ² ts	22.7	21.1	23.3

XYZ coodinates – Calculated Structures

Complex 8



Dd	0.01303000	0 40412300	0 47371000
C C	2 00262500	0 11813800	0.04503100
C	2 68542400	2 16451800	0.04030100
C	-2.00342400	0.00718600	0.18073/00
C	4.07856200	2 27707200	0.10973400
	-4.07030300	-2.37707200	-0.30317000
Н	-1.97891300	-2.96082900	-0.75867300
	-4.89326000	-1.27589400	-0.01894600
н	-4.93064500	0.85006400	0.45408300
н	-4.50160500	-3.37087300	-0.53212800
Н	-5.97613700	-1.39/5/100	0.08550800
C	-0.06450100	2.49880100	0.19875600
C	-2.15694600	1.40408000	0.23440300
С	-0.64427900	3.73487800	0.55156100
Н	1.01508300	2.39558800	0.02845500
С	-2.80199000	2.61109900	0.59171700
С	-2.04091500	3.78837800	0.75152600
Н	-0.01041100	4.61843700	0.66115500
Н	-3.88485200	2.63308300	0.74369500
Н	-2.53170000	4.72707600	1.02705600
С	0.76568700	-0.83288500	1.45904800
Н	0.16451700	-1.71578200	1.74141000
Н	0.44066200	0.09349300	1.96541700
С	2.28260500	-1.03299500	1.42474200
С	2.65778700	-2.51262300	1.20026800
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¹ For further information on using this method to reference NMR spectra, refer to the following website: <u>www.iupac.org/publications/pac/2001/7311/7311x1795.html</u>.

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NMR Spectra of New Compounds









































