

Hydrogen Storage Properties of B- and N-Doped Microporous Carbon

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A B- and N-doped microporous carbon has been synthesized via a substitution reaction. The obtained carbon exhibited much higher surface area than the previously reported B- and N-doped carbon. The hydrogen storage measurements indicated that the B- and N-doped microporous carbon had a 53% higher storage capacity than the carbon materials with similar surface areas. Furthermore, hydrogen storage via spillover was studied on Ru-supported B- and N-doped microporous carbon and a storage capacity of 1.2 wt % at 298 K and 10 MPa was obtained, showing an enhancement factor of 2.2. Ab initio molecular orbital calculations were also performed for the binding energies between the spillover hydrogen atom and various sites on the doped carbon. The theoretical calculations can explain the experimental results well, which also shed light on the most favorable and possible sites with which the spillover hydrogen atoms bind. © 2009 American Institute of Chemical Engineers *AICHE J.*, 55: 1823–1833, 2009

Keywords: microporous carbon, boron- and nitrogen-doped carbon, hydrogen storage, hydrogen spillover, hydrogen binding on carbon

Introduction

In recent years, various carbon materials such as activated carbon (AC), carbon nanotubes (CNTs), microporous carbon and mesoporous carbon have been studied extensively for a variety of applications.^{1–12} Among these studies, there has been growing interest in synthesis of B- and/or N-doped carbon materials due to their unique electronic, mechanical, optical and adsorption properties compared with pure carbon materials.^{13–38} B- and N-doped carbon nanotubes (BCN nanotubes), including single-walled carbon nanotubes (SWNTs), double-walled carbon nanotubes (DWNTs) and multi-walled carbon nanotubes (MWNTs) have been synthesized by various methods including arc-discharge,^{25,26} laser ablation,^{27,28} substitution reaction,^{29–31} and chemical vapor deposition (CVD).^{32–35} B- and N-doped amorphous carbon³⁶ and carbon film³⁷ have been prepared by carbonization of

polymer precursor containing boron and nitrogen. Recently, Vinu et al. reported the synthesis of B- and N-doped mesoporous carbon by using a substitution reaction.³⁸ However, these B- and N-doped carbon materials displayed relatively low-surface areas, and there has been no report on the preparation of B- and N-doped microporous carbon. Developing B- and N-doped carbon materials with different pore structures and a high-surface area is highly desirable.

Herein, we report on the synthesis of a B- and N-doped microporous carbon with a high-surface area. The B- and N-doped microporous carbon was prepared via a substitution reaction involving a plain microporous carbon with a high-surface area and porosity as the template, which facilitated the substitution reaction with the reactant gases of boron oxide and ammonia. More importantly, the obtained B- and N-doped microporous carbon showed a 53% higher hydrogen storage capacity than the corresponding pure carbon materials that had the same surface area. Furthermore, doping of Ru particles on B- and N-doped microporous carbon significantly enhanced its hydrogen uptake by a factor of 2.2 due to hydrogen spillover effect. These results can be explained well by our *ab initio* molecular orbital calculations.

Additional Supporting Information may be found in the online version of this article.

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Experimental Methods

Synthesis

Plain Microporous Carbon. Microporous carbon derived from zeolite EMC-2 was prepared according to a procedure similar to that proposed by Kyotani and Parmentier.^{4,8} EMC-2 was degassed in a flask for 12 h at 473 K, then furfuryl alcohol (FA) was introduced into the flask at a reduced pressure. The mixture of EMC-2 and FA was stirred overnight in a He flow and then filtered. The polymerization of FA in EMC-2 was carried out by heating the composite under a flow of He at 353 K for 24 h, then at 423 K for 8 h. The EMC-2/PFA composite was placed in a vertical quartz tube and heated to 973 K under a N₂ flow. When the temperature reached 973 K, propylene gas (2% in N₂ by volume; flow rate: 150 cm³/min) was passed through the tube for 15 h. After the CVD treatment, the composite was further heated at 1173 K for 2 h under a flow of N₂. The obtained EMC-2/carbon composite was treated in HF solution (40%) for 24 h, and, subsequently, refluxed by concentrated HCl solution for 4 h to dissolve the template of EMC-2. The resulting microporous carbon was collected by filtration and washing with distilled water.

B- and N-doped Microporous Carbon. Typically, equal mass amounts of microporous carbon and B₂O₃ were heated together in an atmosphere of dilute ammonia (1% in He) at 1423 K for 1 h. The heating rate was 60 K/min.

6 wt % Ru Supported on B- and N-doped Microporous Carbon. Typically, 200 mg of well-dried B- and N-doped microporous carbon was dispersed in 20 mL of acetone and was stirred for 0.5 h in a flask at room-temperature. 5 mL of acetone was mixed with 800 mg Ru(NO)(NO₃)_x(OH)_y aqua (1.5 wt % Ru content), which was slowly added to the aforementioned solution under vigorous agitation. The mixture was subjected to ultrasonication (100 W, 42 kHz) for 1 h and then magnetically stirred at room-temperature for 24 h. After being dried in an oven at 333 K overnight, the impregnated sample was transferred to a horizontal quartz tube and treated in a He atmosphere at 1173 K for 1 h.

Mesoporous Carbon (CMK-3). CMK-3 was synthesized according to the reported procedure.¹⁰ In a typical run, 1 g of SBA-15 powder was mixed with an aqueous solution containing 1.25 g of sucrose, 0.14 g of H₂SO₄, and 5.0 g of H₂O. The viscous mixture was placed in a drying oven at 373 K for 6 h, and, subsequently, at 433 K for 6 h. The treated mixture was mixed again with aqueous solution containing 0.75 g of sucrose, 0.09 g of H₂SO₄, and 5.0 g of H₂O. After heating at 433 K for 6 h in the drying oven, the mixture was further carbonized at 1173 K for 4 h under nitrogen atmosphere. The obtained carbon/SBA-15 composite was washed with hydrofluoric acid to remove SBA-15 template.

Characterization

Powder X-ray diffraction (XRD) data were recorded on a Rigaku Miniflex diffractometer at 30 kV, 15 mA for Cu K α ($\lambda = 0.1543$ nm) radiation. High-resolution transmission electron microscopy (HRTEM) images of the materials were obtained on a JEOL 3011 analytical electron microscope. Electron energy-loss spectroscopy (EELS) spectra were

obtained on a JEOL 2010F equipped with a Gatan EELS spectrometer. X-ray photoelectron spectroscopy was recorded on a Kratos Axis ultra XPS spectrometer. Nitrogen adsorption and low-pressure H₂ adsorption isotherms (0-1 atm) were measured with a standard static volumetric technique (Micromeritics ASAP 2020 sorptometer). Hydrogen adsorption at 298 K, and pressures greater than 0.1 MPa and up to 10 MPa was measured using a static volumetric technique with a specially designed Sievert's apparatus. The apparatus was previously tested to prove to be leak-free and accurate through calibration by using LaNi₅, AX-21, zeolites, and MOFs at 298 K. Approximately 200 mg of sample was used for each high-pressure isotherm measurement in this study. Before measurements, the samples were degassed *in vacuo* at 350°C for at least 12 h. The storage data under our experimental conditions refer to excess hydrogen uptake.

Molecular Orbital Calculation

To understand the hydrogen adsorption on the B- and N-doped microporous carbon via spillover, *ab initio* molecular orbital calculations were performed for the binding energies between the spillover hydrogen atom and various sites on the doped microporous carbon.

The Gaussian 03 package³⁹ and Cerius2 molecular modeling software⁴⁰ were used for all molecular orbital (MO) calculations. Density functional theory (DFT) calculations were performed with p function added in the basis set, thus, B3LYP/6-31 g (d,p) was used for geometry optimization, self-consistent field energy (SCF), and single-point energy (SPE) calculations. The molecular structure selected for the graphite models used in this study consisted of 14 aromatic rings in one single layer as shown in Figure 1. For boron and/or nitrogen doped graphite models, one or two carbons on the center basal sites were replaced by boron(s) or nitrogen(s), or both boron and nitrogen. The chemisorption of hydrogen atom(s) on unsubstituted graphite, and various doped graphite models were calculated from the following expression

$$E_{\text{ads}} = E_{\text{graphite-hydrogen}} - E_{\text{graphite}} - E_{\text{hydrogen}}$$

where E_{ads} is the energy of chemisorption, $E_{\text{graphite-hydrogen}}$ is the SCF energy of the optimized structure of hydrogen on graphite, E_{graphite} is the SCF energy of the optimized structure of the corresponding graphite, and E_{hydrogen} is the SCF energy of hydrogen atom.

To determine the potential energy curves for the interaction between H atom(s) and graphite models, SPE was calculated on each change of the carbon-hydrogen bond length, the resulting SPEs were converted to electron volts per hydrogen (eV/H) as shown in Figure 2. Basis set superposition error (BSSE) has been neglected.

Results and Discussion

B- and N-doped microporous carbon

XRD. Powder X-ray diffraction patterns of microporous carbon and B- and N-doped microporous carbon samples are shown in Figure 3. The microporous carbon (Figure 3a) showed a strong peak at $2\theta = 6.3^\circ$, indicating that the

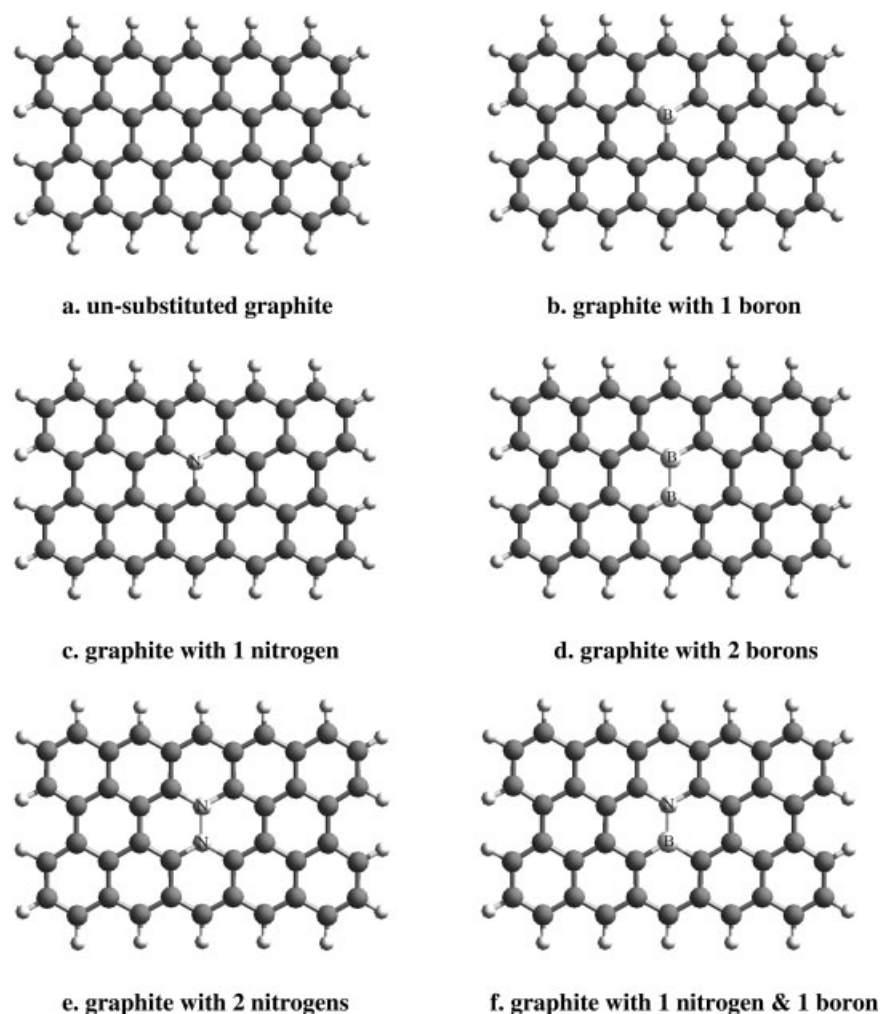


Figure 1. Geometry optimized structures of various graphite models.

ordered structure of zeolite template had been replicated in the carbon. After the substitution treatment, the peak at $2\theta = 6.3^\circ$ disappeared in the low-angle XRD pattern of the B- and N-doped microporous carbon, indicating that the ordered structure of B- and N-doped microporous carbon had been destroyed during the high temperature treatment (Figure 3b). It is worth noticing that the wide-angle XRD pattern of B- and N-doped microporous carbon exhibited two peaks at $2\theta = 26.3^\circ$ and 43.5° , assigned to (002) and (101) diffractions, respectively, of graphite-like structure. This indicates the graphite-like structure was formed in the B- and N-doped microporous carbon during the high-temperature treatment.

Nitrogen Isotherms. Nitrogen isotherms are shown in Figure 4. The isotherms for microporous carbon and B- and N-doped microporous carbon samples both exhibited the typical Type I isotherm, indicating the presence of microporosity in these two samples. The BET surface area and pore volume of the microporous carbon were $3771 \text{ m}^2/\text{g}$ and $1.83 \text{ cm}^3/\text{g}$, respectively. These textural properties were comparable to those of reported zeolite templated carbon with high-surface areas.^{4,8} After the substitution reaction, the BET surface area and pore volume of B- and N-doped microporous carbon were $1215 \text{ m}^2/\text{g}$ and $0.62 \text{ cm}^3/\text{g}$, respectively, which were much lower compared to the microporous carbon tem-

plate. This is mainly attributed to the partial porosity collapse during the substitution reaction at high-temperature, in agreement with the XRD results. However, it is encouraging

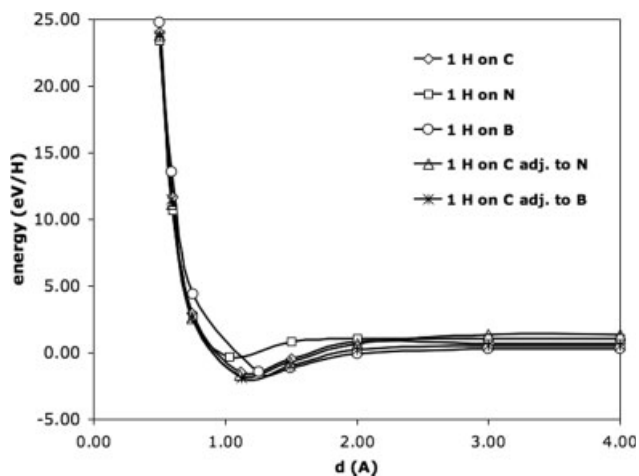


Figure 2. Potential energy diagram for an atomic hydrogen on various graphite models ($1 \text{ eV} = 1.602 \times 10^{-19} \text{ J}$).

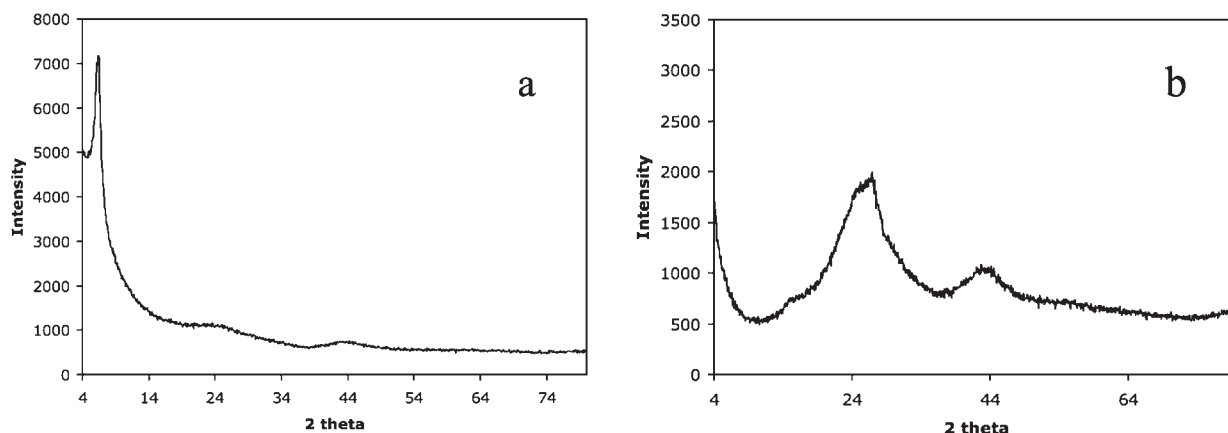


Figure 3. (a) X-ray diffraction patterns of microporous carbon (a), B- and N-doped microporous carbon (b).

that the surface area of the B- and N-doped microporous carbon was much higher than all previously reported high surface area mesoporous BCN (740 m²/g),³⁸ B- and N-doped amorphous carbon (188 m²/g),³⁶ BCN nanotubes (356 m²/g),³¹ and activated BN (168 m²/g),⁴¹ indicating that the use of microporous carbon is helpful for synthesizing high-surface area B- and N-doped microporous carbon.

TEM. High-resolution TEM images of the microporous carbon and the B- and N-doped microporous carbon samples are shown in Figure 5. As shown in Figure 5a, microporous carbon exhibited highly ordered microporous channels, confirming the successful structural transfer from the zeolite template. Although for the B- and N-doped microporous carbon (Figure 5b), it is difficult to observe ordered microporous channels. It is noted that the graphite-like structures could be clearly observed at the edges of the B- and N-doped microporous carbon sample. This indicates that the ordered microstructure collapsed and graphitized during the substitution reaction at high temperature. These results are in agreement with the XRD and nitrogen observations.

EELS. Electron energy-loss spectrum (EELS) of the B- and N-doped microporous carbon clearly showed the presence of the K-shell excitation of carbon, boron, nitrogen and oxygen (Figure 6). Sharp π^* -peaks (lefthand sides of the B and N K-shell excitation edges), and the shape of σ^* -bands

(righthand sides of the edges) are characteristics of B and N atoms in the graphite-like sp² bonded network, indicating that B and N atoms have been substituted into the graphitic

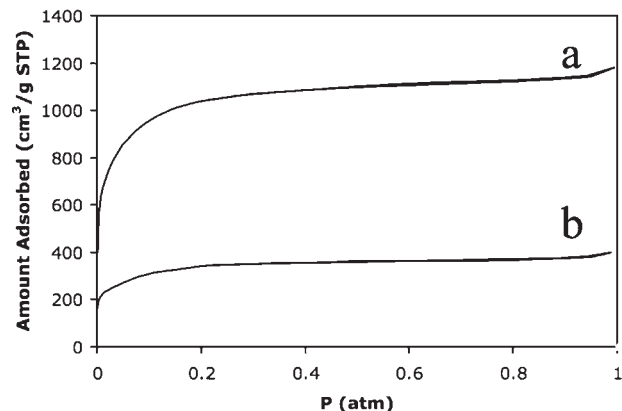


Figure 4. Nitrogen isotherms of microporous carbon (a), B- and N-doped microporous carbon (b).

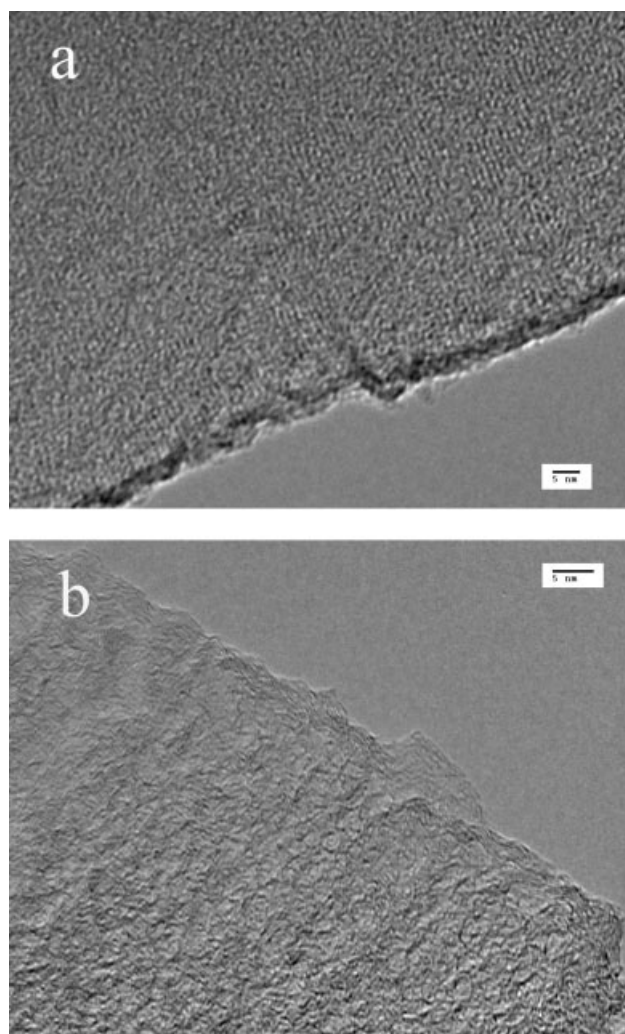


Figure 5. TEM images of microporous carbon (a), B- and N-doped microporous carbon (b), scale bar: 5 nm.

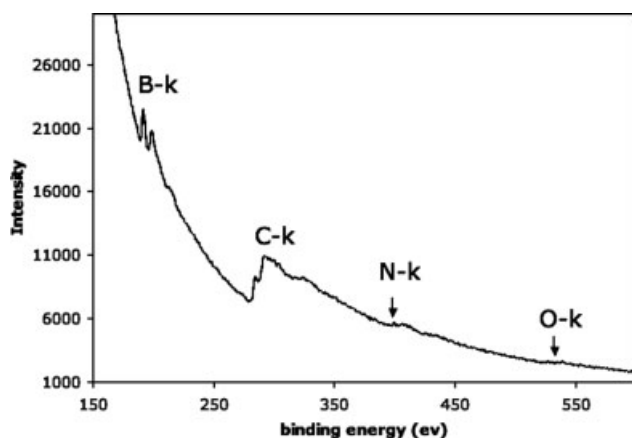


Figure 6. EELS spectrum of B- and N-doped microporous carbon.

rings in the framework of B- and N-doped microporous carbon. This result is in agreement with the observed graphite-like structure of the B- and N-doped microporous carbon from the XRD and TEM results.

XPS. The X-ray Photoelectron survey spectrum (Figure 7) of B- and N-doped microporous carbon showed four sharp signals for B, N, C and O elements, in agreement with the EELS results. The elemental composition of the B- and N-doped microporous carbon was approximately $B_9C_{64}N_{10}O_{17}$, indicating that using the microporous carbon as a template was efficient for the substitution reaction.

SEM. SEM images of the microporous carbon and the B- and N-doped microporous carbon samples are shown in Figure 8. The microporous carbon consisted of pure hexagonal platelets, in good agreement with the typical morphology of EMC-2 zeolite. This also confirms that the structure of the original zeolite has been successfully transferred into the microporous carbon. The B- and N-doped microporous carbon showed the same morphology as that of the microporous carbon. This indicates that the morphology of microporous carbon has been retained during the substitution treatment. No other morphology than hexagonal platelets could be

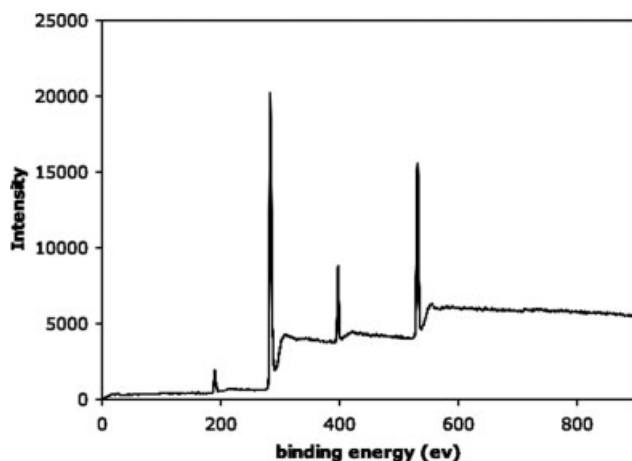


Figure 7. XPS spectrum of B- and N-doped microporous carbon.

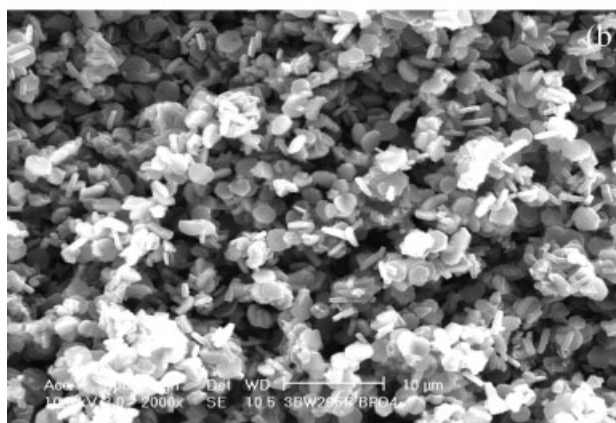
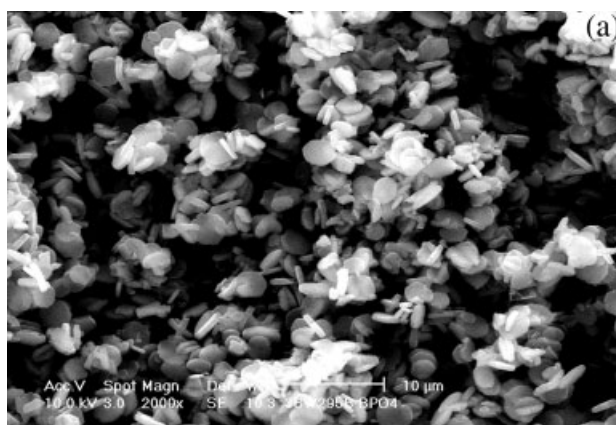


Figure 8. SEM images of microporous carbon (a), B- and N-doped microporous carbon (b).

observed on the B- and N-doped microporous carbon, indicating the purity in the obtained samples.

Hydrogen Isotherms. High-pressure hydrogen isotherms at 298 K for the B- and N-doped microporous carbon are presented in Figure 9. As shown in Figure 9, the B- and N-doped microporous carbon had a hydrogen storage capacity of 0.55 wt % at 298 K and 10 MPa. In comparison, a porous

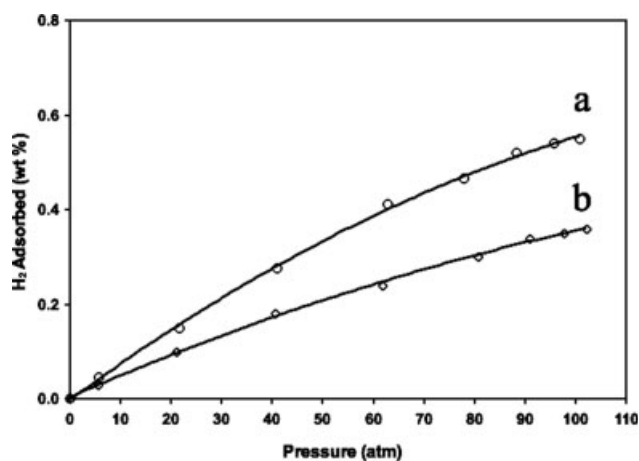


Figure 9. High-pressure hydrogen isotherms at 298 K for B- and N-doped microporous carbon (a: \circ) and porous carbon CMK-3 (b: \diamond).

Table 1. Stability Comparison Between B and N Substituted Graphite Models

Graphite models	SCF	
	Energy (Kcal/mol)	% difference in stability
A-unsubstituted	-1059064.23	
B-substituted with 1 boron	-1050742.55	-0.79
C-substituted with 1 nitrogen	-1069494.67	+0.98
D-substituted with 2 adjacent borons	-1042378.95	-1.63
E-substituted with 2 adjacent nitrogens	-1079885.15	+1.93
F-substituted with 1 boron & 1 nitrogen	-1061201.63	+0.20

pure carbon CMK-3 with a BET surface area of 1260 m²/g had a storage capacity of 0.36 wt % under the same conditions (Supporting Table 1). Obviously, the hydrogen uptake on the B- and N-doped microporous carbon was 53% higher than the carbonaceous materials with similar surface area. Recently, Chung et al. also observed an enhanced hydrogen uptake of 0.4 wt % on a boron substituted carbon with a BET surface area of 780 m²/g at ambient temperature and >60 bar.⁴² Compared with their carbon, the B- and N-doped microporous carbon in this work showed a larger BET surface area (1215 m²/g), and a higher storage capacity for hydrogen.

The heats of adsorption of H₂ on the B- and N-doped microporous carbon were calculated from the H₂ adsorption isotherms at 273 K and 298 K by using the Clausius-Clapeyron equation, as shown in Figure 10. The isosteric heats of adsorption were determined by evaluating the slope of the plot of ln(P) vs. (1/T) at the same adsorption amount. It can be seen that the H₂ adsorption amounts at all pressures up to 1 atm decreased with an increase in temperature. Figure 10 inset showed that the absolute values of the heat of adsorption decreased with adsorption amount for the B- and N-doped microporous carbon. The heats of adsorption was ~14 kJ/mol at low-surface coverage, and leveled off to ~10 kJ/mol at relatively high-surface coverage. The high values of heats of adsorption at low-surface coverage can be attributed to the adsorption of H₂ on the more energetic sites on carbon. Defect sites, edge sites (i.e., armchair and zigzag edge

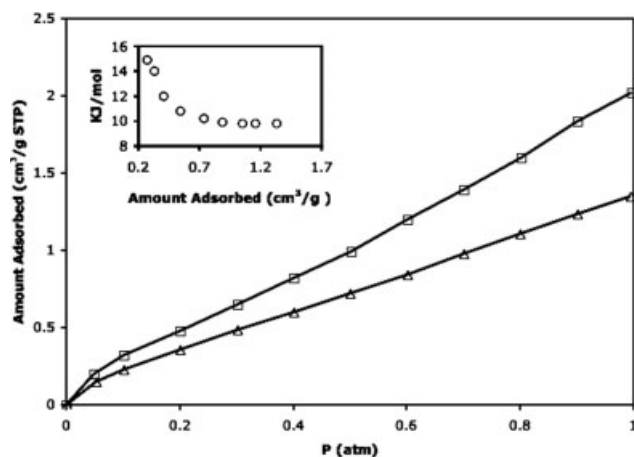


Figure 10. Low-pressure H₂ adsorption isotherms for B- and N-doped microporous carbon at 273 K (□), and 298 K (△); Inset: Calculated isosteric heats of adsorption: -ΔH (○).

sites of graphite crystallites), and surface oxygen complexes are strong sites for adsorption. The heats of adsorption (~10 kJ/mol) at high H₂ adsorption amount on the B- and N-doped microporous carbon were still higher than that on pure carbon (~6 kJ/mol).⁴³ The relative high-heat of adsorption on B- and N-doped microporous carbon suggests more H₂ favorably bonded to B- and N-doped microporous carbon. The higher heats of adsorption on heteroatom substituted carbon than pure carbon have been also reported by other authors.^{42,44}

Ru/B- and N-doped microporous carbon

Hydrogen storage is the crucially missing link to a future “hydrogen economy.”⁴⁵ Our recent studies showed that the hydrogen storage capacities at 298 K in nanostructured and porous materials including carbon, zeolites and metal-organic frameworks could be enhanced significantly by exploiting the hydrogen spillover phenomenon.^{46–52} Hydrogen spillover is defined as the dissociative chemisorption of hydrogen on metal nanoparticles, and subsequent migration of hydrogen atoms onto adjacent surfaces of a receptor via spillover or surface diffusion.^{53–55} Hydrogen storage on various carbon materials (CNTs, active carbon, carbon nanofibers, etc) doped with metals have been studied.^{56–62} However, there is no study on hydrogen spillover-storage on boron- and nitrogen-doped carbon. Enhanced hydrogen storage capacity has been obtained on B- and N-doped microporous carbon in this work, and it is expected that doping of Ru metal on the B- and N-doped microporous carbon would further enhance the storage capacity due to hydrogen spillover. Thus, we synthesized Ru/B- and N-doped microporous carbon and investigated its hydrogen storage properties.

XRD. After doping Ru on the B- and N-doped microporous carbon, the wide-angle XRD pattern (Figure 11) of Ru/B- and N-doped microporous carbon showed typical peaks at 38.6° (100), 42.3° (002), 44.2° (101), 58.5° (102), 69.7° (110) and 78.6° (103) characteristic of the metallic ruthenium (ICDD-JCPDS Card No. 06-0663). The particle sizes of Ru calculated from Scherrer equation were approximately 10 nm. These results confirm that nanosized Ru metals have

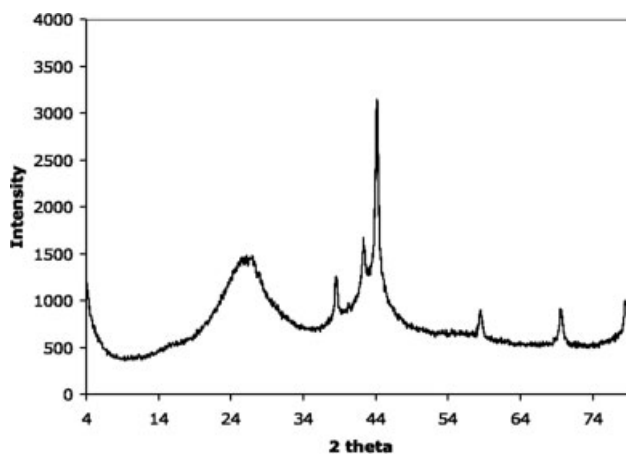


Figure 11. X-ray diffraction pattern of Ru/B- and N-doped microporous carbon.

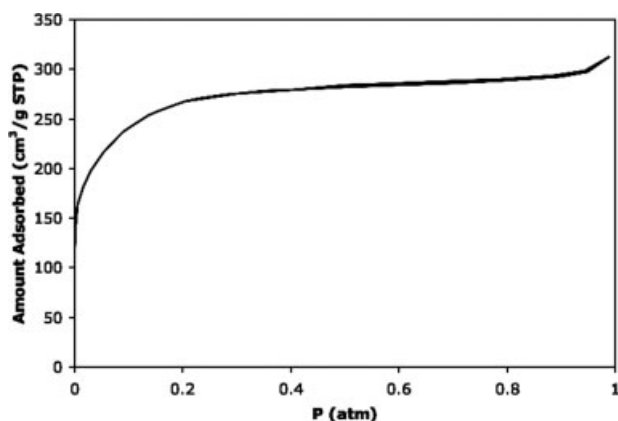


Figure 12. Nitrogen isotherm of Ru/B- and N-doped microporous carbon.

been doped on the B- and N-doped microporous carbon. Ru/B- and N-doped microporous carbon showed a broad peak around 26.3° , the same as that of B- and N-doped microporous carbon, which corresponds to the (101) diffraction of graphite-like structure.

Nitrogen Isotherm. The nitrogen isotherm (Figure 12) of Ru/B- and N-doped microporous carbon showed typical Type I isotherm, indicating the presence of microporosity. The BET surface area and pore volume of the Ru/B- and N-doped microporous carbon were $953 \text{ m}^2/\text{g}$ and $0.48 \text{ cm}^3/\text{g}$, respectively. The BET surface area and pore volume of the Ru/B- and N-doped microporous carbon were lower than that of the B- and N-doped microporous carbon (Supporting Table 1). This was due to the increased weight and micropore blocking caused by Ru metal particles.

TEM. High-resolution TEM images of Ru/B- and N-doped microporous carbon are shown in Figure 13. The nanosized black spots of Ru were well dispersed on the surface of the Ru/B- and N-doped microporous carbon. Particles of 5–10 nm can be clearly observed. It is noted that XRD pattern is more sensitive to large particles, thus, the

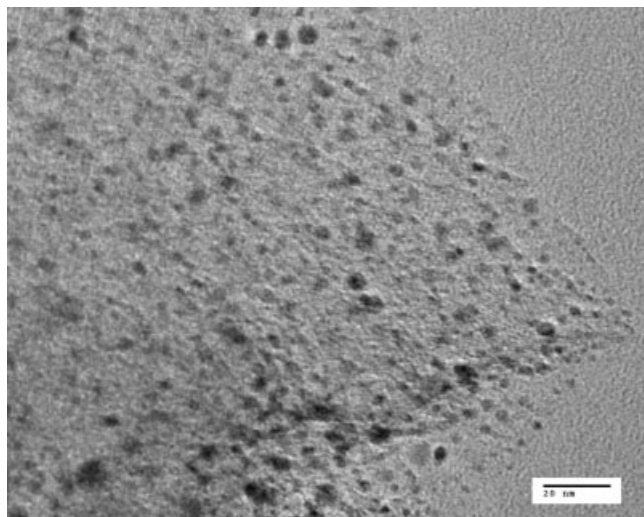


Figure 13. TEM image of Ru/B- and N-doped microporous carbon, scale bar: 20 nm.

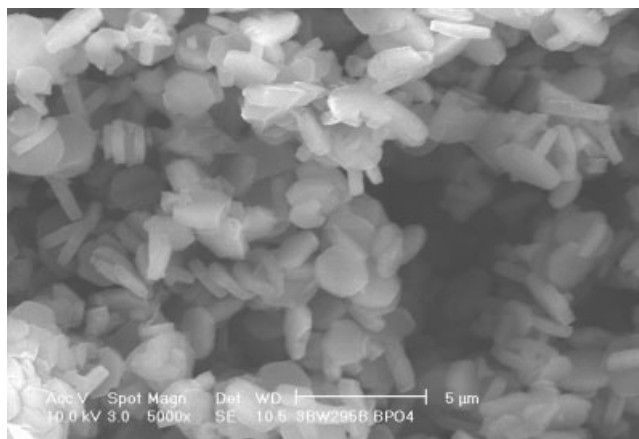


Figure 14. SEM image of Ru/B- and N-doped microporous carbon.

TEM observations were in agreement with the XRD results. These results further confirm that Ru metals have been successfully doped on the B- and N-doped microporous carbon support.

SEM. SEM image (Figure 14) showed that the Ru/B- and N-doped microporous carbon consisted of the pure hexagonal platelets, the same morphology as that of the microporous carbon and B- and N-doped microporous carbon. This indicates that the morphology of the B- and N-doped microporous carbon has been retained during the doping treatment. No other morphology than hexagonal platelets can be observed on the Ru/B- and N-doped microporous carbon, indicating the purity in the obtained sample.

Hydrogen Isotherms. High-pressure hydrogen isotherms at 298 K for the B- and N-doped microporous carbon and the Ru/B- and N-doped microporous carbon samples are compared in Figure 15. The B- and N-doped microporous carbon had a hydrogen storage capacity of 0.55 wt % at 298 K and 10 MPa. By doping 6.0 wt % Ru metal on the B- and N-doped microporous carbon, the hydrogen uptake at 10

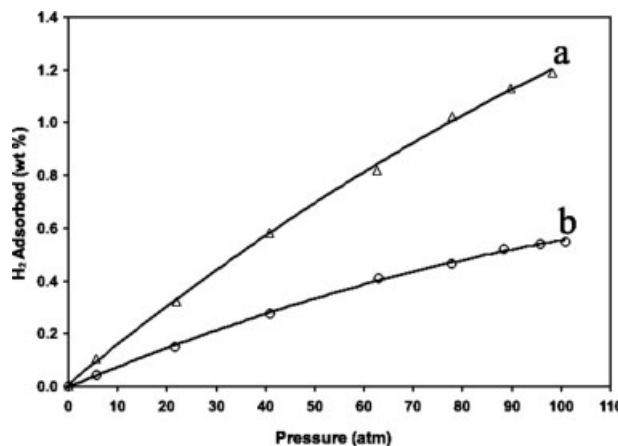


Figure 15. High-pressure hydrogen isotherms at 298 K on Ru/B- and N-doped microporous carbon (a: Δ) and B- and N-doped microporous carbon (b: \circ).

MPa was increased to 1.2 wt %, i.e., more than doubled. The enhanced hydrogen storage capacity should not be attributed to any differences in surface area because the Ru/B- and N-doped microporous carbon had a lower surface area than that of the B- and N-doped microporous carbon. The enhancement of hydrogen storage was due to the spillover of atomic hydrogen from the Ru metal particles to the B- and N-doped microporous carbon. Ru, Pt, Ni and Pd metals are all effective as hydrogen dissociation sources, and the enhanced hydrogen storage by metal doped on pure carbon materials (CNTs, active carbon, carbon nanofiber, etc.) have been reported.⁵⁶⁻⁶² In our case, in comparison with the B- and N-doped microporous carbon, it is remarkable that the hydrogen adsorption amount of the Ru/B- and N-doped microporous carbon has been enhanced by a factor of 2.2. Furthermore, reversibility on Ru/B- and N-doped microporous carbon was evaluated by measuring the desorption branch down to 1 atm. It can be seen from supporting Figure 1 that the desorption branch nearly followed the adsorption branch, although there appeared to be a slight hysteresis. The sample was then evacuated to a pressure of 1 Pa for 12 h at 298 K and total desorption occurred. The second adsorption isotherm was in agreement with the first adsorption isotherm, indicating reversible adsorption.

The heats of adsorption of H₂ on the Ru/B- and N-doped microporous carbon were calculated from the H₂ adsorption isotherms at 273 K and 298 K by using the Clausius-Clapeyron equation. As shown in Figure 16, the heats of adsorption were > 20 kJ/mol at low-surface coverage and leveled off to ~15 kJ/mol at relatively high-surface coverage. The high values of heats of adsorption at low-surface coverage can be attributed to the strong adsorption of H atoms on the Ru particles, as well as the H atoms on the strongest sites on carbon. It has been reported that atomic hydrogen can be strongly adsorbed at defect sites on carbon materials and some defects can be generated during the catalyst preparation.⁶³⁻⁶⁵

It is known that the Clausius-Clapeyron equation yields the overall heats of adsorption. In this case, the heats of adsorption are approximately the overall values of the bonding energies of H₂ on carbon, H atoms on metal and various carbon sites. As a first-order analysis, we take the heat of adsorption at high H₂ adsorption amount as an indicator of the adsorption strength of hydrogen atoms on the surface of the Ru/B- and N-doped microporous carbon. The heat of adsorption at high H₂ adsorption amount on the Ru/B- and N-doped microporous carbon sample is about 15 kJ/mol. The relative higher heat of adsorption on Ru/B- and N-doped microporous carbon than on B- and N-doped microporous carbon suggests more H atoms were favorably bonded to Ru/B- and N-doped microporous carbon, in agreement with the high-pressure results. It is noted that the heat of adsorption on Ru/B- and N-doped microporous carbon is also higher than Ru on pure carbons (11 kJ/mol).⁶² This indicates that B and N doping is more favorable for hydrogen adsorption via spillover than that on pure carbons. From our experimental results, hydrogen storage capacity on B- and N-doped carbon can be significantly enhanced via spillover. However, it is still not clear that which sites on the doped carbon are more favorable for adsorption of spillover H atoms. Understanding the contributions of various sites on

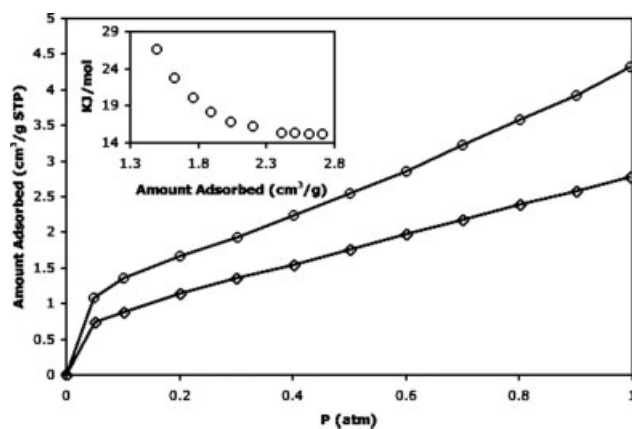


Figure 16. Low-pressure H₂ adsorption isotherms for Ru/B- and N-doped microporous carbon at 273 K (○), and 298 K (◇); Inset: Calculated isosteric heats of adsorption: $-\Delta H$ (○).

doped carbon will be helpful for further design of promising adsorbents for hydrogen storage. Thus, the B and N doping effect on hydrogen adsorption is further discussed in the following molecular orbital calculation results.

Molecular orbital calculations

It is known that hydrogen spillover involves a number of complex steps: dissociation on the metal, migration of H from metal to carbon, migration on carbon, recombination on carbon, and binding on carbon sites. However, complete simulation of the spillover process has not been done; simulations of some of the aforementioned steps have provided insights into the spillover process.⁶⁶ Meantime, much evidence has been reported in the literature on the existence of H atoms that are involved in the spillover process.⁵⁵ Recent examples include direct observation using inelastic neutron scattering studies^{67,68} which have provided evidence of atomic hydrogen spillover from Pt to carbon at room-temperature. The calculation results for H atoms binding on different graphene sites only represent the last step of the spillover process. Without a complete simulation of the spillover process, it is not possible to directly relate the binding energy from the calculation and the overall heats of adsorption that are obtained from the temperature dependence of the isotherms. However, it is clear that they are related; i.e., a stronger binding energy leads to a higher overall heat of adsorption. Thus, the calculation results should provide an indication whether boron and nitrogen doping will lead to higher heats of adsorption, or more favorable adsorption.

Theoretical calculations of hydrogen atoms on boron and nitrogen doped carbon materials have been reported by several authors,^{44,69-72} some of these studies also dealt with molecular hydrogen interaction on carbon materials. In case of chemisorption of hydrogen atom on boron- and nitrogen-doped carbon materials, the interaction of hydrogen atom on carbon materials substituted with both boron and nitrogen on the same graphite structure has not been reported.

Because it has been reported that graphite models containing less than 13 aromatic rings may result in incorrect result from their MO calculations,⁴⁴ in this study, we selected a

Table 2. Bond Length (Å) in Geometry Optimized Graphite Models

Graphite models	Bond length (Å)	Reported ⁴⁴ for graphene
A-unsubstituted	C-C : 1.421	1.422
B-substituted with 1 boron	B-C : 1.492	1.487
C-substituted with 1 nitrogen	N-C : 1.415	1.419
D-substituted with 2 adjacent borons	B-C : 1.499	
E-substituted with 2 adjacent nitrogens	N-C : 1.410	
F-substituted with 1 boron & 1 nitrogen	N-C : 1.410	

graphite model with 14 aromatic rings in the molecular structure. The central one or two carbon atoms were substituted for a boron or/and a nitrogen atoms, thus, resulting in various boron and nitrogen substituted graphite models, their geometry optimized structures are shown in Figure 1. The optimized structure of the graphite remained as a flat sheet after substitutions, but became slightly curved after hydrogen adsorption in all graphite models.

The result of our MO calculations indicated that boron substitution destabilized the graphite structure, which was also observed by Zhu et al.⁴⁴ However, the opposite was true for nitrogen substitution, which stabilized the graphite structure as shown in Table 1. This may be explained by the difference in valence electrons. Boron has one less valence electron than carbon, it does not form resonance structure with its three neighboring carbons, therefore, boron-carbon bond length is slightly longer than the unsubstituted carbon-carbon bond, whereas nitrogen has one more valence electron than carbon, it can still form resonance structure with its neighboring carbons, so nitrogen-carbon bond length is slightly shorter, as shown in Table 2. When the graphite structure was substituted with both 1 boron and 1 nitrogen at the same time, then the structure was found to be still slightly stabilized.

When hydrogen is chemisorbed on the center carbon of the unsubstituted graphite, the C-C bond length of the active carbon to the neighboring three carbons is increased to 1.51 Å, this is due a change of C-C resonance bond to C-C single bond. In case of substituted graphite models, B-C bond is

Table 3. Bond Length (Å) in Geometry Optimized Graphite Models with Chemisorption of Hydrogen(s)

Graphite models	Bond length (Å)	Reported ⁷¹ for (8,0) carbon nanotubes
A-unsubstituted with 1 hydrogen on carbon	C-C : 1.510	1.509
	C-H : 1.116	
B-with 1 hydrogen on boron	B-C : 1.533	1.573
	B-H : 1.250	
C-with 1 hydrogen on nitrogen	N-C : 1.511	1.506
	N-H : 1.036	
D-with 2 hydrogens on 2 borons	B-C : 1.545	
	B-H : 1.228	
E-with 2 hydrogens on 2 nitrogens	N-C : 1.514	
	N-H : 1.045	
F-with 1 hydrogen on 1 boron, and 1 hydrogen on 1 nitrogen	B-C : 1.575	
	B-H : 1.224	
	N-C : 1.496	
	N-H : 1.023	

Table 4. Calculated E_{ads} for Chemisorption of 1 Hydrogen (Kcal = 4.184 KJ)

Graphite models	E_{ads} (Kcal/mol)	Reported ⁷¹ for (8,0) carbon nanotubes
A-unsubstituted with 1 hydrogen on carbon	-33.77	-46.7
B-with 1 hydrogen on boron	-34.01	-49.8
B-with 1 hydrogen on carbon adj. to boron	-43.73	
C-with 1 hydrogen on nitrogen	-6.32	-28.4
C-with 1 hydrogen on carbon adj. nitrogen	-38.34	

longer than C-C bond, and N-C bond is close to the C-C bond for a single nitrogen substitution. In case of substitution with one nitrogen and one boron, the N-C bond is the shortest, i.e., 1.496 Å and the B-C bond is the longest, i.e., 1.575 Å, thus, B-C > C-C > N-C, this same trend is observed for the corresponding hydrogen bonds, i.e., B-H > C-H > N-H as shown in Table 3. Some of our observed bond lengths are quite comparable to those reported by Zhu⁴⁴ and Zhou⁷¹ on their carbon model system.

In Tables 4 and 5, the calculated energy of chemisorption E_{ads} in Kcal/mol/H is tabulated for one hydrogen atom, and two hydrogen atoms, respectively. When hydrogen is chemisorbed directly on boron, the E_{ads} is only slightly increased compare to the unsubstituted graphite model, and when directly on nitrogen, it is substantially decreased. However, when hydrogen is adsorbed on a carbon adjacent to boron or nitrogen, E_{ads} is increased and is greater than the unsubstituted value.

The potential energy diagram in Figure 2 can further illustrate this trend. The same trend is also observed for two hydrogen atoms chemisorbed on the graphite models. Thus, E_{ads} is much higher when hydrogen is on adjacent carbon to boron or nitrogen compare to directly on boron or nitrogen, and it is highest when 2 hydrogens are on carbons adjacent to a boron and a nitrogen. The E_{ads} calculation indicates the doped carbon is favorable for hydrogen adsorption, which is in agreement with the experimental observations. The calculation results further shed light on the most favorable and possible sites with which the spillover hydrogen atoms bind.

Conclusions

In this study, we have demonstrated the synthesis of B- and N-doped microporous carbon via a substitution reaction. The obtained B- and N-doped microporous carbon exhibited much higher surface area than previously reported high-surface area B- and N-doped mesoporous carbon, carbon

Table 5. Calculated E_{ads} for Chemisorption of 2 Hydrogens

Graphite models	E_{ads} (Kcal/mol/H)
B-with 1 hydrogen on boron & 1 hydrogen on adjacent carbon	-35.23
C-with 1 hydrogen on nitrogen & 1 hydrogen on adjacent carbon	-25.67
F-with 1 hydrogen on boron & 1 hydrogen on nitrogen	-29.69
F-with 1 hydrogen on carbon adj. to boron & 1 hydrogen on carbon adj. to nitrogen	-47.16

nanotubes, amorphous carbon, and activated BN due to the application of a microporous carbon template with high-surface area and porosity. The hydrogen storage tests showed B- and N-doped microporous carbon had a 53% higher storage capacity than the pure carbon with a similar surface area. Doping of Ru on B- and N-doped microporous carbon further significantly enhanced its hydrogen uptake by a factor of 2.2 due to the spillover effect. Hydrogen adsorption on B- and N-doped microporous carbon via spillover has been further studied by molecular orbital calculations, and the experimental results are in agreement with the theoretical predictions on the effects of B- and N-substitution.

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