

**The Effect of Mars-relevant Soil Analogs on the Water Uptake of Magnesium Perchlorate and Implications for the Near-Surface of Mars**

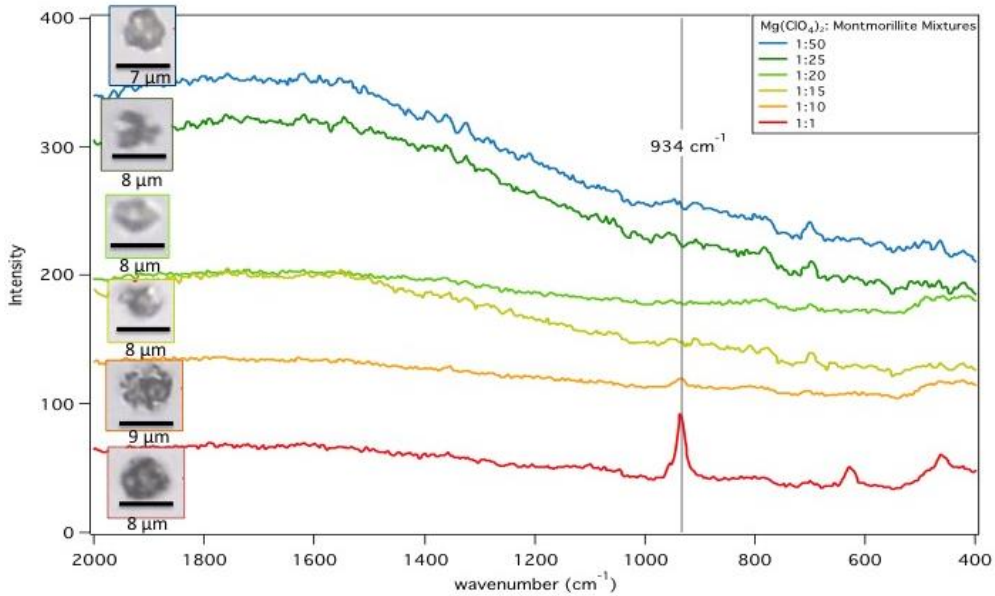
K.M. Primm<sup>1,2</sup>, R.V. Gough<sup>1,2</sup>, J. Wong<sup>2</sup>, E. G. Rivera-Valentin<sup>3</sup>, G. M. Martinez<sup>4</sup>, J. V. Hogancamp<sup>5</sup>, P. D. Archer<sup>5</sup>, D. W. Ming<sup>5</sup>, and M.A. Tolbert<sup>1,2\*</sup>. <sup>1</sup>Cooperative Institute for Research in Environmental Sciences, University of Colorado, Boulder, CO, 80309, USA; <sup>2</sup>Department of Chemistry and Biochemistry, University of Colorado, Boulder, CO, 80309, USA; <sup>3</sup>Lunar and Planetary Institute, Universities Space Research Association, Houston, TX; <sup>4</sup>Department of Climate and Space Sciences and Engineering, University of Michigan, Ann Arbor, MI, USA; <sup>5</sup>Jacobs at NASA Johnson Space Center, Houston, TX 77058

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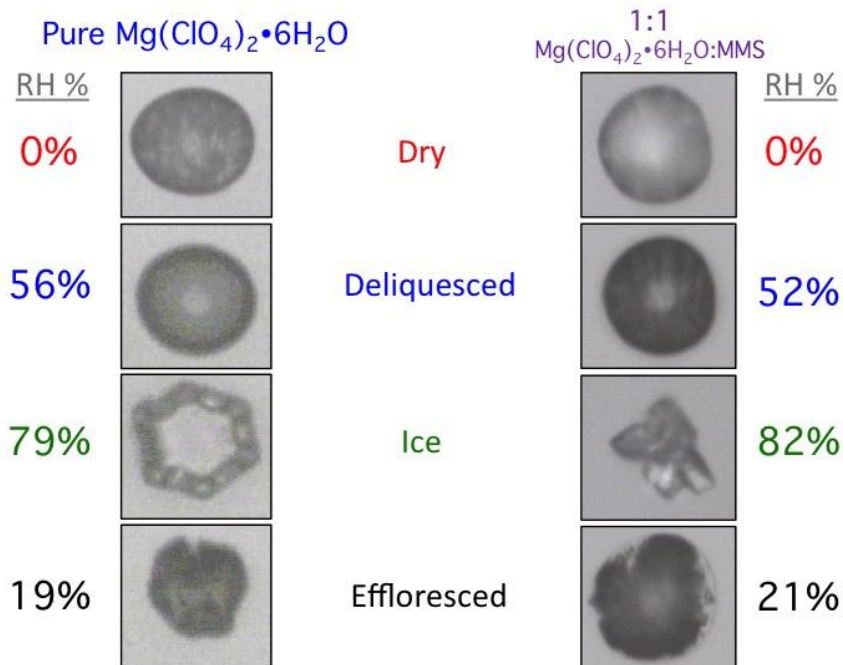
Figures S1 and S2  
Tables S1

**Introduction**

The following 2 Figures and 1 Table are supplementary information that has been mentioned in the main manuscript text.



**Figure S1. Raman spectra of the perchlorate region ( $934\text{ cm}^{-1}$ ) of magnesium perchlorate and montmorillonite mixtures. The range is from 1:1 to 1:50 magnesium perchlorate: montmorillonite. The spectra show that the perchlorate peak is the strongest in the 1:1 mixture and completely disappears at 1:10, thus the reason for choosing a 1:1 mixture of salt to mineral.**



**Figure S2.** Optical images of pure  $\text{Mg}(\text{ClO}_4)_2 \cdot 6\text{H}_2\text{O}$  undergoing phase transitions (left) compared to images of the same phase transitions for the mixture of 1:1  $\text{Mg}(\text{ClO}_4)_2 \cdot 6\text{H}_2\text{O}:\text{MMS}$  (right). Although Raman spectra was unable to be obtained due to laser absorption of MMS, the following images show the visual changes while the mixture particle took up and released water. These images were taken where  $25^\circ\text{C} \geq T \geq -53^\circ\text{C}$ .

**Table S1. Complete data set of DRH, ERH, and ice formation RH of pure  $\text{Mg}(\text{ClO}_4)_2 \cdot 6\text{H}_2\text{O}$ , 1:1  $\text{Mg}(\text{ClO}_4)_2 \cdot 6\text{H}_2\text{O}$ : montmorillonite, and 1:1  $\text{Mg}(\text{ClO}_4)_2 \cdot 6\text{H}_2\text{O}$ :MMS.**

<b>1:1 <math>\text{Mg}(\text{ClO}_4)_2</math>:Mont.</b>			
	T(K)	% RH of Ice formation	$S_{\text{ice}}$
	$255.6 \pm 0.1$	$101.4 \pm 0.5$	$1.20 \pm 0.01$
	$246.5 \pm 1.7$	$99.9 \pm 1.1$	$1.30 \pm 0.01$
	$233.4 \pm 1.7$	$89.7 \pm 4.5$	$1.32 \pm 0.05$
	$224.9 \pm 0.6$	$81.2 \pm 2.9$	$1.28 \pm 0.03$
	$214.1 \pm 1.5$	$89.3 \pm 1.2$	$1.53 \pm 0.01$
	T(K)	DRH	
	$265.0 \pm 0.5$	$46.7 \pm 0.5$	
	$252.3 \pm 0.6$	$49.4 \pm 0.6$	
	$245.0 \pm 0.4$	$49.9 \pm 0.9$	
	$232.1 \pm 0.3$	$55.8 \pm 1.1$	
	$222.4 \pm 1.0$	$61.5 \pm 5.1$	
	T(K)	ERH	
	$273.7 \pm 0.8$	$24.4 \pm 0.7$	
	$262.5 \pm 0.6$	$21.0 \pm 1.0$	
	$254.3 \pm 1.5$	$22.9 \pm 1.8$	
	$243.2 \pm 1.0$	$21.0 \pm 0.6$	
<b>1:1 <math>\text{Mg}(\text{ClO}_4)_2</math>:MMS</b>			
	T(K)	% RH of Ice formation	$S_{\text{ice}}$
	$256.6 \pm 1.4$	$96.1 \pm 3.3$	$1.13 \pm 0.03$
	$241.2 \pm 1.6$	$95.3 \pm 1.6$	$1.30 \pm 0.02$
	$223.8 \pm 2.2$	$82.4 \pm 1.9$	$1.31 \pm 0.02$
	T(K)	DRH	
	$266.8 \pm 0.7$	$42.0 \pm 0.4$	
	$249.1 \pm 2.0$	$46.2 \pm 0.8$	
	$227.0 \pm 2.8$	$54.4 \pm 2.5$	
	T(K)	ERH	
	$277.0 \pm 0.2$	$20.0 \pm 1.1$	
	$258.8 \pm 2.5$	$21.4 \pm 1.6$	
	$237.2 \pm 2.2$	$22.2 \pm 0.7$	