

REACTIONS OF BIS(HEXAMETHYLBENZENE)IRON(0) WITH CARBON MONOXIDE AND WITH UNSATURATED HYDROCARBONS

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Summary

Thermal decomposition of bis(hexamethylbenzene)iron(0) in the presence of carbon monoxide yields a novel carbonyl iron complex, $[\text{C}_6(\text{CH}_3)_6]\text{Fe}(\text{CO})_2$. The cyclohexadiene complex $[\text{C}_6(\text{CH}_3)_6]\text{Fe}(\text{C}_6\text{H}_8)$ is obtained from reaction of bis(hexamethylbenzene)iron(0) with either 1,3-cyclohexadiene or benzene, and the yield is much greater in the presence of hydrogen gas. Interaction of bis-(hexamethylbenzene)iron(0) with 2-butyne induces a catalytic cyclotrimerization to give more hexamethylbenzene. Kinetic and isotope distribution studies indicate that the primary step in these reactions is not a direct loss of one ring ligand, but rather an insertion of the iron center into one of the ligand methyl C–H bonds, leading to a benzyl hydride complex species. Mechanisms for the subsequent reactions of this iron hydride species are proposed.

Introduction

We recently investigated some of the pathways available for stoichiometric and catalytic reactions of the cyclopentadienylcobalt(I) half-sandwich system [1]. This 14-electron core structure, with two coordination sites available for a variety of reactions, shows interesting analogies to as well as characteristic differences from comparable reaction systems based on other 14-electron transition metal sandwich cores. We have now investigated the reactivity patterns of reaction systems based on an isoelectronic arene–iron(0) half-sandwich structure. Hexamethylbenzeneiron(0) requires two ligand electron pairs to complete its 18-electron valence shell and also contains its metal in a low oxidation state susceptible to oxidative addition by relatively inert substrates. It should thus

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exhibit interesting reactions, potentially applicable to homogeneous catalysis.

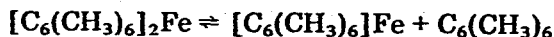
A useful entry into such arene-iron(0) reaction systems appeared to be available via thermal decomposition of bis(hexamethylbenzene)iron(0) (I) a 20-valence-electron species, reported by Fischer and Röhrscheid to be light and temperature sensitive [2]. We describe below studies of the reactions which occur when the decomposition is carried out in the presence of various substrates.

Results

1. Reaction of I with carbon monoxide

Solutions of the bis(arene) complex I in petroleum ether decay to hexamethylbenzene and metallic iron at 20°C with a half life of about 5 days. If carbon monoxide is admitted, however, the color of I disappears fairly rapidly, with a half life of about 8 min. Removal of solvent and subsequent fractional sublimation (first at 20°C to remove hexamethylbenzene, then at 40°C) gives $[\text{C}_6(\text{CH}_3)_6]\text{Fe}(\text{CO})_2$ (II). This compound is characterized by two IR absorptions at 1954 and 1892 cm^{-1} , by the appropriate parent and fragment ions in its mass spectrum, and by an NMR spectrum with a sharp singlet at τ 8.21 ppm, indicating the equivalence of all the methyl groups present*.

The kinetics of the reaction between bis(arene)iron(0) (I) and CO to form II were studied spectroscopically at CO pressures between 4 and 12 atm. at 21°C. The reaction is first order in I, but independent of CO concentrations in the range of 4×10^{-2} to 1.4×10^{-1} M, the rate constant being 8.3×10^{-2} min^{-1} . This result must mean that CO is not involved in the rate-limiting step, which must therefore occur either before or after CO attack. An essentially saturated pre-equilibrium between I and CO to form a monocarbonyl complex, the decomposition of which then governs the rate, would be conceivable but can be ruled out, since the spectrum of I is initially unchanged in the presence of CO. We must thus assume that the rate-determining step for the CO reaction is the formation of some intermediate, which in the absence of CO is in reversible equilibrium with the starting material I. This intermediate cannot be the half-sandwich hexamethylbenzeneiron(0), since neither the decomposition of I, nor its reaction with CO is significantly inhibited by an excess of hexamethylbenzene in solution up to the its saturation concentration. Such an inhibition would be observable if the equilibrium:



determined the rate of subsequent reactions. We thus conclude that this intermediate III still contains all the components, and so is an isomer, of the starting material I. This intermediate, the identity of which is to be discussed in more

* Synthesis of a dimer of this species, $[\{\text{C}_6(\text{CH}_3)_6\}\text{Fe}(\text{CO})_2]_2$, was reported by Fischer et al. [3]. No equilibrium was observed between this dimer and the monomer reported here. Preliminary mass spectral and NMR data obtained on the dimeric carbonyl indicate that this species might have a composition $[\{\text{C}_6(\text{CH}_3)_6\text{H}\}\text{Fe}(\text{CO})_2]_2$; such a cyclohexadienyl species, obtained by single hydrogenation of each arene ring, would also allow for a metal-metal bond, usually associated with bridging carbonyl ligands [4].

detail below, then appears to undergo a subsequent ligand exchange, which is fast at all CO pressures studied, to form the dicarbonyl II.

2. Reaction of I with 1,3-cyclohexadiene

If the decomposition of I is carried out at 40°C in neat 1,3-cyclohexadiene, the diene complex $[\text{C}_6(\text{CH}_3)_6]\text{Fe}(\text{C}_6\text{H}_8)$ (IV) is obtained in about 67% yield after removal of excess diene and of hexamethylbenzene. This complex is characterized by a mass spectrum with a parent peak at m/e 298 and fragment peaks at m/e 218 ($[\text{C}_6(\text{CH}_3)_6]\text{Fe}^+$), 162 ($[\text{C}_6(\text{CH}_3)_6]^+$) and 147 ($[\text{C}_6(\text{CH}_3)_5]^+$). The NMR spectrum exhibits a sharp singlet at τ 8.01 ppm due to 18 equivalent methyl protons, and a double doublet at τ 6.00 ppm due to the "central" diene protons, as well as unresolved multiplets at τ 8.30 and 8.60 ppm due, respectively, to the "terminal" diene protons and the four methylene protons*. The formation of IV at 20°C from 1 in pure cyclohexadiene occurs with a first-order rate constant of $6.8 \times 10^{-3} \text{ min}^{-1}$. This reaction is slower by an order of magnitude than the decomposition of I in the presence of CO. We conclude, that cyclohexadiene is less efficient than CO in facilitating loss of hexamethylbenzene from the intermediate III.

3. Reaction of I with 2-butyne.

Decomposition of bis(arene)iron(0) (I), in petroleum ether at 45°C in the presence of a ninetyfold excess of 2-butyne leads to the formation of additional hexamethylbenzene. After a reaction time of 42 h 5 mmol of $\text{C}_6(\text{CH}_3)_6$ are obtained per mmol of starting material I. Even after subtraction of 1–2 mmoles of $\text{C}_6(\text{CH}_3)_6$ formed by decomposition of I, a significant amount of the 2-butyne used must have undergone catalytic trimerization to $\text{C}_6(\text{CH}_3)_6$.

A fraction containing a yellow-brown material can be sublimed from the reaction mixture. This material exhibits a parent ion at m/e 326 and a fragment ion at m/e 218 in its mass spectrum corresponding to a composition $[\text{C}_6(\text{CH}_3)_6]\text{Fe}[\text{C}_4(\text{CH}_3)_4]$ (V). It could, thus be an iron-hexamethylbenzenecyclobutadiene complex. We could not isolate and further characterize this material however, because of the small yields and the difficulties encountered in separating it from the large excess of hexamethylbenzene present.

4. Reaction of I with arene solvents and hydrogen

Decomposition of I in benzene always leads to formation of small amounts (ca. 3%) of the cyclohexadiene complex IV, while decomposition in toluene leads to the methyl homologue $[\text{C}_6(\text{CH}_3)_6]\text{Fe}(\text{C}_6\text{H}_7\text{CH}_3)$. These products were identified by their mass spectra.

That the two extra hydrogen atoms gained in the conversion of benzene to a cyclohexadiene ligand are largely derived from the hexamethylbenzene rings of I is established by the isotopic distribution of this product formed in benzene- d_6 ; more than half of the newly acquired hydrogen atoms being found to consist of ^1H .

Admission of hydrogen to the reaction mixture of I in benzene induces a rather dramatic increase in the yields of cyclohexadiene complex IV; about

* Largely similar NMR data have been reported for $(\text{C}_6\text{H}_8)\text{Fe}(\text{CO})_3$ [5].

50–60% of I is now converted to IV, compared with 3–5% in the absence of hydrogen. At the same time, the fraction of ^1H in the two extra hydrogens in the product obtained from benzene- d_6 rises to over 90%, indicating an efficient utilization of H_2 gas for the hydrogenation of the benzene substrate. Similar results were obtained in a reaction with D_2 gas and undeuterated benzene- h_6 .

Undoubtedly, in the decomposition of I, some iron–hydride intermediate is formed which can then transfer its hydrogen ligands to an arene. In the presence of both benzene and hydrogen gas, such a hydride intermediate leads to a preponderant conversion of compound I to the cyclohexadiene compound IV.

Further information on the course of the hydrogen transfer reaction is obtained from an analysis of the stereochemistry of the resulting cyclohexadiene complex. Studies on NMR spectra of related cyclo-diene complexes have been described [5,6] and it appears certain that the *exo*- and *endo* positions of such a complex differ significantly in their respective chemical shifts. The methylene hydrogen signals in $[\text{C}_6(\text{CH}_3)_6]\text{Fe}(\text{C}_6\text{D}_6\text{H}_2)$ (prepared with H_2 gas) appear exclusively on the high-field side of the quartet of methylene proton peaks, and partial deuteration leads to partial disappearance of this high-field part of the complex methylene spectrum. We conclude that the two incoming hydrogen atoms are transferred to the same side of the newly formed cyclohexadiene ligand, although we do not know whether they are both in *exo*, or both in *endo* position. A *cis*-hydrogen transfer to unsaturated substrates is common in homogeneous hydrogenations; in our case it is readily explained by successive transfer of two hydrogen ligands from some hydride-containing intermediate to a coordinated benzene molecule.

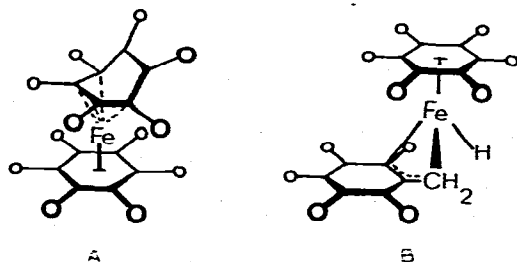
Decay of I in benzene, under vacuum or 1 atm of N_2 , occurs with a half-life of 10 h at 20°C . The reaction rate ($1.15 \times 10^{-3} \text{ min}^{-1}$) is intermediate between that for decay of I in alkanes ($1 \times 10^{-4} \text{ min}^{-1}$) and in cyclohexadiene ($6.8 \times 10^{-3} \text{ min}^{-1}$). In the presence of 1 atm of H_2 , the rate of decay of I in benzene is the same as that under vacuum or N_2 .

Discussion

An essential point of our study is the identity of the intermediate III, which appears to exist in a reversible isomerisation equilibrium with I, and apparently undergoes rapid conversion to the dicarbonyl II in the presence of CO. For the rate determining, primary isomerisation step, two possibilities appear particularly plausible. First, one of the aromatic rings could rearrange to a geometry with only one or two double bonds coordinated to the iron atom (IA), such as that observed in the stable form of the analogous bis(arene) ruthenium(0) complex $[\text{C}_6(\text{CH}_3)_6]_2\text{Ru}$ [7,8]. Alternatively, the intermediate in could be a benzyl hydride species $[\text{C}_6(\text{CH}_3)_6]\text{Fe}(\text{H})[\text{CH}_2\text{C}_6(\text{CH}_3)_5]$, derived from I by the insertion of the iron atom into one of the ligand C–H bonds (IB), analogous to the insertions observed e.g. with decamethyl titanocene derivatives [9]. Such a species could attain an 18-electron configuration by the coordination of an allylic ring segment occasionally observed with benzyl complexes [10]. The rate data obtained for the reaction of I with CO indicate that the rate-limiting step in the reaction with CO, which eventually leads to the dicarbonyl complex II, is probably an isomerisation of I to the benzyl hydride species IB.

The constancy of the rate of formation of II at different CO pressures indicates that at these CO pressures the reconversion of the crucial isomerisation product to the original structure I does not compete with its conversion to the carbonyl complex II. It follows that under these conditions the observed rate of formation of II ($8.3 \times 10^{-2} \text{ min}^{-1}$) is equal to the rate of isomerisation of I to III. A mere distortion of the axially symmetric structure I to an unsymmetrically π -coordinated benzene complex is unlikely to yield a forward rate constant of only about $8 \times 10^{-2} \text{ min}^{-1}$. That the regeneration of I from IA is slow in comparison to carbonyl complex formation is even less likely. For the analogous bis(hexamethylbenzene)ruthenium(0) complex, for which a distorted *tetrahapto* structure of one of the benzene rings has been established from crystal structure data [8] and from NMR spectra at lower temperatures, the NMR data at room temperature are consistent with an interconversion of coordinated and uncoordinated ring positions which is rapid on the NMR time scale [7]. Conversion of I to IB by insertion of the iron atom into a C—H bond, on the other hand, is a reasonable possibility for an isomerisation with a rate similar to the observed rate of formation of II, and with a reverse reaction too slow to compete against the bimolecular reaction with CO.

SCHEME 1. Possible structures for intermediate III: $-\text{O} = -\text{CH}_3$.

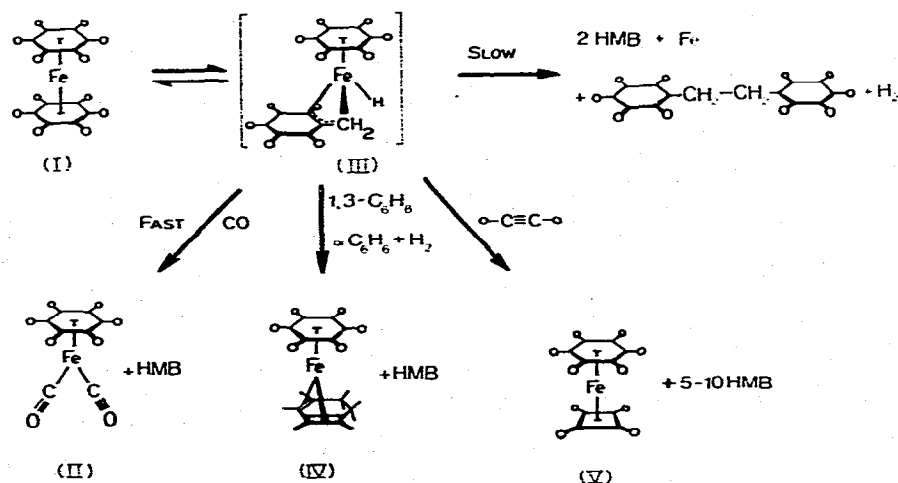


Additional evidence for the occurrence of a benzylhydride intermediate in the thermal decomposition of I is derived from the observation that small but significant amounts of hydrogen gas (ca. 0.05 mol/mol of I) are reproducibly formed during the decomposition of I, and that the organic decomposition product, mainly hexamethylbenzene, always contains some bis(pentamethylphenyl) ethane, $(\text{CH}_3)_5\text{C}_6\text{H}_2\text{CH}_2\text{CH}_2\text{C}_6(\text{CH}_3)_5$ as shown by mass spectrometry. These two products are to be expected if a ligand exchange between two units of IB occasionally precedes the reductive elimination of hydride and benzyl ligand groups.

The occurrence of a benzyl hydride species as an intermediate in thermal reactions of I can be rationalized as follows. Direct elimination of hexamethylbenzene from the 20-electron species I would generate the highly unstable, monomeric 14-electron species $[\text{C}_6(\text{CH}_3)_6]\text{Fe}(0)$. Through isomerisation of I to an intermediate with structure IB, on the other hand, reaction paths which involve only 16- and 18-electron species are available. Conversion of structure IB to a *monohapto*-benzyl structure [11], uptake of a CO ligand, and subsequent reductive elimination of hexamethylbenzene from the 18-electron intermediate $[\text{C}_6(\text{CH}_3)_6]\text{Fe}(\text{CO})\text{H}[\text{h}^1\text{-CH}_2\text{C}_6(\text{CH}_3)_5]$ can generate the 16-electron species $[\text{C}_6(\text{CH}_3)_6]\text{Fe}(\text{CO})$. Other ligands, such as 2-butyne or 1,3-cyclohexadiene could

react in the same way with IB and thus convert it to an 18-electron *monohapto*-benzyl hydride, again facilitating the elimination of hexamethylbenzene by circumvention of an otherwise required 14-electron intermediate. The observed acceleration of the decay of I by these substrates is therefore clearly comparable with the intervention of intermediate III (structure IB) in the formation of products IV and V as well as in that of II, as summarized in Scheme 2. In the reaction of 2-butyne, a common intermediate for the formation of both V and the extra hexamethylbenzene might be a metallocyclopentadiene species similar to those observed in related $(C_5H_5)Co$ half-sandwich system [1]. From such an intermediate, formation of a cyclobutadiene compound V or coordination of an additional 2-butyne molecule, with subsequent regeneration of I could occur as competing reactions. Repeated regeneration of I by coordination of acetylene ligands, followed by decay to III and acetylene-assisted hexamethylbenzene elimination, would provide a satisfactory explanation for the catalytic acetylene trimerization.

SCHEME 2. Proposed pathways for decomposition of I and for reactions of intermediate III with various substrate molecules in the formation of dicarbonyl complex II, diene complex IV, and of a complex V which probably contains a cyclobutadiene ligand. Intermediate III was not observed directly and is postulated primarily on the basis of kinetic data (see text); $-O = -CH_3$; HMB = $C_6(CH_3)_6$.



As for the incorporation of benzene into the 1,3-hexadiene compound IV which occurs even in the absence of H_2 , a hydride intermediate such as III would provide a reasonable pathway for the observed transfer of $C_6(CH_3)_6$ protons to an arene molecule. The question whether the arene substrates interact with intermediate III directly or rather with a secondary hydride complex derived from it remains unanswered, as does that concerning the kind of hydride species involved in the incorporation of molecular H_2 into the 1,3- C_6H_8 complex. In this respect, it is significant that in the presence of excess D_2 gas, 66% of the two extra hydrogen atoms are still found to be 1H . Since the solvent provides virtually no hydrogen in this transfer, it appears that even in the presence of molecular hydrogen the hexamethylbenzene ligand is still the major source of hy-

drogen atoms transferred. This indicates that a ligand-derived hydride such as III must intervene even in the hydrogen-assisted conversion of benzene into a C_6H_8 ligand; excess H_2 or D_2 appears to favour some sort of hydride structure derived from III, so as to make its reaction with benzene near-stoichiometric.

Interestingly, a stable tetrameric hydride $(C_5H_5)_4Co_4H_4$ has been isolated for the related isoelectronic cobalt(I) half-sandwich system [12]. The non-accumulation of hexamethylbenzeneiron(0) hydrides in the absence of a reducible substrate makes their identification difficult, but this very instability is, of course, an essential feature of their participation in efficient hydride transfer reactions.

Experimental

General techniques used

All manipulations were carried out under nitrogen or argon, using Schlenk or high-vacuum techniques.

Aluminum bromide and sodium (reagent grade, Merck), naphthalene (scintillation grade, Merck), hexamethylbenzene (pure grade, Roth) and iron(II) bromide (Alfa Inorganics) were used without further purification. 1,3-Cyclohexadiene (EGA-Chemie), 2-butyne (Chemical Procurement Laboratories) and all solvents (reagent grade, Merck) were doubly distilled, and stored under vacuum over a suitable drying and deoxygenating agent.

Visible, IR, NMR and mass spectra were recorded on normal commercial instruments. Mass spectra were recorded at several different ionization voltages, usually down to a nominal voltage of 30 eV, to minimize fragmentation.

*(1) Preparation of $[C_6(CH_3)_6]_2Fe(0)$ **

A suspension of 5.0 g (7.5 mmol) of dry $[C_6(CH_3)_6]_2Fe$ (II) $(PF_6)_2$ (prepared according to [13,14]), 0.50 g naphthalene, and 0.52 g (23 mmol) of sodium in 100 ml of tetrahydrofuran was stirred at room temperature. The solid phase slowly turned purple. After a few hours a dark brown solution formed, but the solid phase remained. After about 8 h, the solid phase was also black. (The Teflon stir bar becomes black owing to residual sodium naphthalene when the reaction is complete.)

The tetrahydrofuran was removed in vacuo. The naphthalene was sublimed out at 40°C during 12 h, and the residue transferred to a Soxhlet extractor (adapted for work under grease-free vacuum conditions by replacing ground joints and stopcocks with O-ring joints and Teflon valves). Extraction with pentane was carried out at approximately 20°C in a partial vacuum until the extracting solvent was colourless.

The black solid was removed from the extract by filtration under argon, washed three times with 10 ml portions of pentane, and dried thoroughly in vacuo, to give a yield of 1.53 g (54%) of $(C_6(CH_3)_6)_2Fe(0)$. The IR spectrum

* Preparation of bis(hexamethylbenzene)iron(0) as described in ref. 2 yields pure material, but in small yields. Alternative methods of reduction of the di-cation lead to the formation of variable amounts of a dihydrogenated side product, $[C_6(CH_3)_6][C_6(CH_3)_6H_2]Fe(0)$. In the sodium naphthalide reduction used here, production of this impurity is low enough to allow its complete removal (as judged by its absence from the mass spectrum of the remaining material) by several washings of the product with petroleum ether.

agreed with that previously reported [2]. The mass spectrum showed a parent ion at m/e 380 and fragment ions at m/e 218 $[(C_6(CH_3)_6)Fe]^+$ and at m/e 162 $[(C_6(CH_3)_6)]^+$.

(2) Reaction of $[C_6(CH_3)_6]_2Fe(0)$ with carbon monoxide

A solution of 294 mg (0.775 mmol) of $(C_6(CH_3)_6)_2Fe(0)$ in 40 ml of petroleum ether was exposed to 3.44 mmol of carbon monoxide on a vacuum manifold and stirred at 40°C for 16 h, during which the solution turned pale brown and a pale brown precipitate formed. Unreacted carbon monoxide (1.69 mmol) was removed with a Toepler pump; CO uptake: 1.75 mmol CO/mmol $[C_6(CH_3)_6]_2Fe(0)$. Removal of the petroleum ether in vacuo left a purple-brown residue. Hexamethylbenzene and the orange dicarbonyl product were sublimed out at 60°C. Resublimation at 20°C removed the hexamethylbenzene, and the hexamethylbenzeneiron dicarbonyl was then resublimed at 40°C. A yield of 126 mg (60%) of $[C_6(CH_3)_6]Fe(CO)_2$ is obtained. Its mass spectrum shows a parent peak at m/e 274 with major fragments at m/e 246, 218 and 162. Other spectral properties of this compound are discussed above.

3. Reaction of $(C_6(CH_3)_6)_2Fe(0)$ with 1,3-cyclohexadiene

Degassed and dried cyclohexadiene (10 ml) was distilled in vacuo on to 357 mg (0.94 mmol) of $[C_6(CH_3)_6]_2Fe(0)$. The solution was stirred at 40°C for 48 h, during which the solution remained dark brown. After removal of the cyclohexadiene in vacuo, the hexamethylbenzene and the orange product were sublimed out at 40°C. In a slow resublimation, the hexamethylbenzene was removed at 17°C, followed by sublimation of the $[C_6(CH_3)_6](C_6H_8)Fe(0)$ complex at 25°C. The yield of hexamethylbenzene(cyclohexadiene)iron(0) was 193 mg (67%). The Nujol-mull infrared spectrum of the product agrees well with that of $(C_6H_6)(C_6H_8)Fe(0)$ [6] when account is taken of the differing aromatic rings: The main absorptions of IV occur at 1310 m, 1250 m, 1160 s, 1058 m, 990 m, 844 m and 822 m cm^{-1} . Other spectral properties of this complex are discussed above.

4. Reaction of $[C_6(CH_3)_6]_2Fe(0)$ with 2-butyne

Approximately 1 mmol of $(C_6(CH_3)_6)_2Fe(0)$ was dissolved in 50 ml of petroleum ether. About 90 mmol of 2-butyne was distilled in, and the solution was stirred at 45°C for 42 h, after which a large amount of white precipitate was present with the dark brown solution. The IR and NMR spectra of the precipitate, after filtration and recrystallization, showed it to be pure hexamethylbenzene. A total of 790 mg (4.9 mmol) of hexamethylbenzene was recovered.

The filtrate from this reaction gave, in addition to hexamethylbenzene, a very small amount of brown-gold sublimate at 40°C, which shows two peaks in the mass spectrum at m/e 326 and 218. The m/e 326 peak is consistent with a formulation $[C_6(CH_3)_6][C_4(CH_3)_4]Fe(0)$.

5. Thermal decomposition of $[C_6(CH_3)_6]_2Fe(0)$ in alkanes

A solution of 320 mg (0.60 mmol) of $[C_6(CH_3)_6]_2Fe(0)$ in 50 ml of cyclohexane was stirred at 80°C for 40 h, during which time a gray-black precipitate formed and the solution becomes pale gold. Collection of evolved gas with a

Toepler pump indicated that 0.037 mmol was generated, and the gas was identified as hydrogen by conversion into H_2O in a CuO tube at 310°C . After removal of cyclohexane and sublimation of hexamethylbenzene a dark gray material was left, and was metallic iron, judged by its magnetism.

6. Thermal decomposition of $[\text{C}_6(\text{CH}_3)_6]_2\text{Fe}(0)$ in benzene

A solution of 307 mg (0.81 mmol) of $[\text{C}_6(\text{CH}_3)_6]_2\text{Fe}(0)$ in 30 ml of benzene is stirred at 60°C for 5 h, after which it was pale in colour. 0.05 mmol of hydrogen was evolved (0.06 mmol H_2 /mmol $[\text{C}_6(\text{CH}_3)_6]_2\text{Fe}(0)$) and metallic iron produced. Filtration of the solution and removal of the benzene in vacuo gave a pale orange product, which sublimed at 25°C . The mass and NMR spectra indicated that this sublimate is a mixture of hexamethylbenzene and $[\text{C}_6(\text{CH}_3)_6](\text{C}_6\text{H}_8)\text{Fe}(0)$ (IV). The yield of the latter product after further purification was 3–5%.

7. Thermal decomposition of $[\text{C}_6(\text{CH}_3)_6]_2\text{Fe}(0)$ in deuterated benzene

When the above reaction is carried out in deuterated benzene, a deuterated analog of the cyclohexadiene complex is obtained. Mass spectral evidence indicates that 66.2% of the added hydrogen atoms are ^1H .

8. Thermal decomposition of $[\text{C}_6(\text{CH}_3)_6]_2\text{Fe}(0)$ in deuterated benzene in the presence of H_2

A solution of 188.0 mg (0.495 mmol) of $[\text{C}_6(\text{CH}_3)_6]_2\text{Fe}(0)$ in 15 ml of deuterated benzene was stirred under 1.97 mmol of H_2 at 45°C for 12 h. The solution became orange and a dark precipitate formed, some of which was metallic iron. Removal of residual H_2 with a Toepler pump showed that 0.30 mmol of hydrogen were taken up during the experiment (0.61 mmol H_2 /mmol $[\text{C}_6(\text{CH}_3)_6]_2\text{Fe}(0)$). A yield of 81 mg (54%) of $[\text{C}_6(\text{CH}_3)_6](\text{C}_6\text{D}_{6-n}\text{H}_{2-n})\text{Fe}(0)$ was obtained. The mass spectrum showed that only 6.9% of the hydrogen atoms taken up by the benzene substrate were deuterium.

9. Thermal decomposition of $[\text{C}_6(\text{CH}_3)_6]_2\text{Fe}(0)$ in benzene in the presence of D_2

When the above experiment was carried out in C_6H_6 under deuterium, the corresponding products were obtained in similar yields but 34.1% of the added hydrogens were deuterium.

10. Kinetics of the reaction of $[\text{C}_6(\text{CH}_3)_6]_2\text{Fe}(0)$ with carbon monoxide

The kinetics were observed at pressures of from 4 to 12 atm CO, providing solutions with a CO concentration between $4.64 \times 10^{-2} \text{ M}$ and $1.39 \times 10^{-1} \text{ M}$, at a constant temperature of 21°C .

Typically, 15 ml of a $5 \times 10^{-4} \text{ M}$ solution of $[\text{C}_6(\text{CH}_3)_6]_2\text{Fe}(0)$ was syringed into a high-pressure cell with 5 cm pathlength. An initial absorption measurement was taken at 650 nm. Carbon monoxide was added to the cell at the appropriate pressure and the disappearance of the 650 nm absorption was measured as a function of time. Plots of $\ln(c - c_\infty)$ vs. time for each experiment yielded first order rate constants in the range of 7×10^{-2} to $10.7 \times 10^{-2} \text{ min}^{-1}$; the reproducibility of the rate constants was within $\pm 25\%$.

11. Kinetics of reactions of $[\text{C}_6(\text{CH}_3)_6]_2\text{Fe}(0)$ without gaseous reactants

A 1 cm cell was used for the decomposition of $[\text{C}_6(\text{CH}_3)_6]_2\text{Fe}(0)$ in petroleum

ether and the reactions with benzene and cyclohexadiene. The rate of decay of the peak at 465 nm was recorded.

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