

Di- and Triheteronuclear Metal Cluster Anions Containing Transition Metals

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Introduction

None of the naked metal cluster polyanions reported so far, such as Sn_9^{4-} [1], Sb_7^{3-} [2], Sn_4^{2-} [3], Sn_5^{2-} [4], $\text{Sn}_8\text{Ti}^{5-}$ [3], $\text{Sn}_{9-x}\text{Ge}_x^{4-}$ ($x = 0-9$) [3] contain transition metal atoms. These species have been either isolated in solid state through the use of cryptate ligands in amine solvents or very recently for Sn_9^{4-} as $[\text{K}(\text{HMPA})_2]_4\text{Sn}_9$ in en/HMPA [5], or studied in solution, e.g. $\text{Sn}_{9-x}\text{Ge}_x^{4-}$. However, we have shown recently by the syntheses of $\text{K}_4[(\text{PPh}_3)_2\text{PtSn}_9]$ and $\text{K}_4[(\text{PPh}_3)_2\text{PtPb}_9]$ [6] that a transition metal can be incorporated into a metal cluster polyanion if it is bonded to exocenter ligands, in close parallelism with heteroborane chemistry [7]. In this communication we wish to report the new triheteronuclear metal cluster polyanions, $\text{Na}_5[\text{L}_2\text{PtTiSn}_8]$ and $\text{Na}_5[\text{L}_2\text{PdTiSn}_8]$. These react quickly to pick up another Sn atom, probably from a solid phase, liberating the formally Ti^- to yield the species $\text{Na}_4[\text{L}_2\text{PtSn}_9]$ and $\text{Na}_4[\text{L}_2\text{PdSn}_9]$, this latter species being previously unreported in the literature.

Experimental

Solutions of $\text{K}_4[\text{Sn}_9]$, $\text{Na}_4[\text{Sn}_9]$ and $\text{Na}_5[\text{TiSn}_8]$ were obtained by extracting alloys of compositions KSn_2 , $\text{NaSn}_{2.25}$ and $\text{NaSnTi}_{1.5}$ as previously described [3, 6, 8]. The complexes $\text{Pt}(\text{PPh}_3)_4$ and $\text{Pd}(\text{PPh}_3)_4$ were synthesized following literature procedures. Ethylenediamine(en) had been dried with Na/benzophenone and stored under vacuum.

All syntheses were conducted in effectively the same way. The reaction H cell is schematically indi-

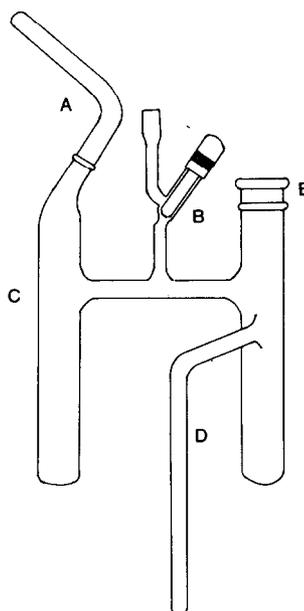


Fig. 1. Schematic representation of the H cell. A, tiptube; B, stopcock, teflon valve; C, H cell; D, NMR tube; E, stopper.

cated in Fig. 1. The synthesis of $\text{Na}_4[\text{L}_2\text{PdSn}_9]$ was as follows:

In a glove box filled with pre-purified dinitrogen, 0.3 g of $\text{NaSn}_{2.25}$ (0.26 mmols of Sn_9^{4-}) were placed in the left arm of the H cell and 0.3 g of $\text{Pd}(\text{PPh}_3)_4$ (0.26 mmols) in the tip tube. Thoroughly dried en (10 ml) were added to the left arm, and after assembling, the H cell was taken out of the glove box. Vacuum (0.01 torr) was applied through the stopcock and the cell stored for 10 days at room temperature to allow the alloy to dissolve. The contents of the tip tube were added to the deep orange-red solution of $\text{Na}_4[\text{Sn}_9]$. After a few hours a progressive darkening of the solution was observed. When no more color change was apparent the dark brown solution was carefully transferred to the other arm of the cell, and an appropriate fraction of it to the NMR tube which was later torch-sealed.

Discussion

The conversion of $\text{Na}_5[\text{L}_2\text{PtTiSn}_8]$ (I) and $\text{Na}_5[\text{L}_2\text{PdTiSn}_8]$ (IV) into $\text{Na}_4[\text{L}_2\text{PtSn}_9]$ (II) and $\text{Na}_4[\text{L}_2\text{PdSn}_9]$ (V), as will be discussed later, poses serious difficulty to a solid state study of the former species. However, the nuclear magnetic properties of the participating elements provide information on the composition and behaviour of the clusters. We have used ^{119}Sn -NMR spectroscopy to study these

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TABLE I. NMR Spectral Data for the Metal Clusters.^a

Compound	δ (ppm)	$J_{119\text{Sn}-205\text{Tl}}$	$J_{119\text{Sn}-175\text{Pt}}$	$J_{119\text{Sn}-117\text{Sn}}$	Ref.
$\text{Na}_5[\text{L}_2\text{PtTlSn}_8]$ (I)	-744	1451 (Hz)	1500 (Hz)	36.3 (Hz)	This work
$\text{Na}_4[\text{L}_2\text{PtSn}_9]$ (II)	-753		1554	95.2	This work
$\text{K}_4[\text{L}_2\text{PtSn}_9]$ (III)	-736		1544	79.0	6
$\text{Na}_5[\text{L}_2\text{PdTlSn}_8]$ (IV)	-745	1418		51.0	This work
$\text{Na}_4[\text{L}_2\text{PdSn}_9]$ (V)	-755			39.0	This work
$\text{Na}_4[\text{Sn}_9]$	-1230			254.0	8
$\text{Na}_5[\text{Sn}_8\text{Tl}]$	-1167	793		410	3

^aChemical shifts are referred to tetramethyl tin (TMT). Upfield chemical shift values from TMT are designated negative.

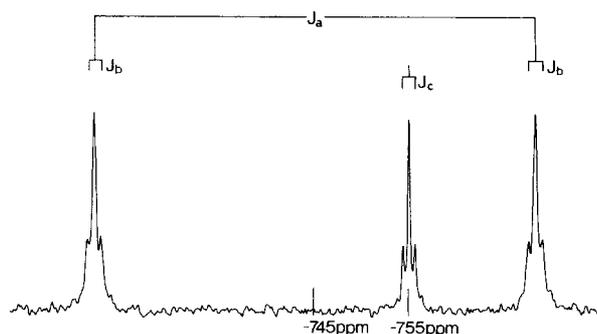


Fig. 2. ^{119}Sn -NMR spectrum of the mixture $\text{Na}_5[\text{L}_2\text{PdTlSn}_8]$ (IV) and $\text{Na}_4[\text{L}_2\text{PdSn}_9]$ (V); $J_a \equiv J_{119\text{Sn}, 205, 203\text{Tl}}$; $J_b \equiv J_{119\text{Sn}-117\text{Sn}}$ (IV); $J_c \equiv J_{119\text{Sn}-117\text{Sn}}$ (V); after one month of mixing the reactants.

clusters, as it has been proven very useful in related [3, 6, 8] and non-related compounds [9–11].

Addition of $\text{Pt}(\text{PPh}_3)_4$ to a solution of $\text{Na}_5[\text{Sn}_8\text{Tl}]$ in en causes a slow but gradual change of the solution from deep orange-red to dark brown. The ^{119}Sn -NMR spectrum of the latter solution displays at -744 ppm a doublet of triplets (1:4:1) with each component of the triplet showing fine structure. This kind of spectrum is in accordance with a fluxional cluster of formulation $\text{L}_2\text{PtTlSn}_8^{5-}$ (I).

In addition another triplet (1:4:1) of quintuplets (0.06:0.325:1.0:0.325:0.06)* is observed in the spectrum at -753 ppm assigned by comparison, to the previously described $\text{L}_2\text{PtSn}_9^{4-}$ (II–III) cluster [6].

In similar fashion, addition of $\text{Pd}(\text{PPh}_3)_4$ to a solution of $\text{Na}_4[\text{Sn}_8\text{Tl}]$ also produces a darkening of the solution. The ^{119}Sn -NMR spectrum of this solution (Fig. 2) displays a doublet of quintuplets at -745 ppm and a quintuplet at -755 ppm which are assigned to $\text{Na}_5[\text{L}_2\text{PdTlSn}_8]$ (IV) and $\text{Na}_4[\text{L}_2\text{PdSn}_9]$ (V), respectively.

Further evidence that supports the assignment of compound (V) was obtained by reacting $\text{Pd}(\text{PPh}_3)_4$

with a solution of $\text{Na}_4[\text{Sn}_9]$ in en, paralleling the reaction of $\text{Pt}(\text{PPh}_3)_4$ with $\text{K}_4[\text{Sn}_9]$ to give III [6]. The ^{119}Sn -NMR of the solution displayed a quintuplet, with $\delta(119\text{-Sn})$ and $J_{119\text{Sn}-117\text{Sn}}$ being the same as these for compound V.

It is interesting to note that in the reactions of both $\text{Pd}(\text{PPh}_3)_4$ and $\text{Pt}(\text{PPh}_3)_3$ with $\text{Na}_5[\text{Sn}_8\text{Tl}]$, the relative concentrations of the resulting products vary as a function of time. Within a day after initiation of the reaction, the concentration of the diheteronuclear species (Pt case) rose until it exceeded that of the triheteronuclear. The conversion process was faster for the platinum system, for which one week after mixing no triheteronuclear species could be detected by ^{119}Sn -NMR, whereas in the Pd case both species still existed in comparable concentrations at the end of a month. We believe that the triheteronuclear species form quickly and then slowly react with a Na–Sn phase, freeing the Tl and forming the diheteronuclear species.

Due to a partial overlap of the ^{119}Sn -NMR absorptions of compounds I and IV (see Fig. 2) the number of Sn atoms in the cluster cannot be calculated exactly from the intensity data and must be inferred by analogy to the better established $\text{L}_2\text{PtSn}_9^{4-}$ and $\text{L}_2\text{-PdSn}_9^{4-}$ clusters. It is known that ^{119}Sn chemical shifts usually parallel electronegativities and good

*This pattern arises because of $J_{119\text{Sn}-117\text{Sn}}$ coupling. Calculations based on abundances of 8.58 and 7.61% for ^{119}Sn and ^{117}Sn respectively have been carried out. See R. W. Rudolph, R. C. Taylor and D. C. Young, 'Fundamental research in Homogeneous Catalysis'. M. Tutsui, Ed., Plenum Press, New York, 1979, pp. 997–1005.

The calculated relative intensities of the multiplet pattern vary significantly with the size of the cluster; e.g. the five central lines for a 8.9 and 10-atom cluster are calculated to be 0.034:0.276:1.000:0.276:0.034, and 0.044:0.311:1.000:0.311:0.044, and 0.056:0.345:0.056, respectively. Because of the small $J_{119\text{Sn}-117\text{Sn}}$ coupling, some overlapping of signals takes place, which favors the smaller ones.

chemical shift-charge correlations have been found in the series of clusters $\text{Sn}_{9-x}\text{Pb}_x^{4-}$ ($x = 0-9$) and Sn_4^{2-} [3]. In addition we have observed that replacement of a L_2Pt group by a L_2Pd group does not change significantly the $\delta(119\text{-Sn})$ of a $\text{L}_2\text{MSn}_9^{4-}$ cluster, and the same is to be expected in the iso-electronic $\text{L}_2\text{MSn}_8\text{TI}^{5-}$ series.

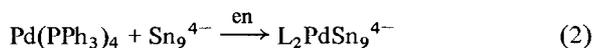
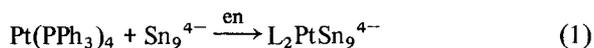
In going a step further in this reasoning we see that Sn_9^{4-} and $\text{Sn}_8\text{TI}^{5-}$ are isoelectronic and $\delta(119\text{-Sn})$ are very close at -1230 ppm and -1167 ppm, so that similar chemical shifts would be expected for $\text{L}_2\text{MSn}_9^{4-}$ and $\text{L}_2\text{MSn}_8\text{TI}^{5-}$ clusters, as is in fact the case ($\cong -750$ ppm for $\text{L}_2\text{MSn}_9^{4-}$ and $\cong -740$ ppm for $\text{L}_2\text{MSn}_8\text{TI}^{5-}$ (see Table I). This is what would be expected by comparison with the reactions of Sn_9^{4-} and Pb_9^{4-} where the initial number of Sn or Pb atoms in the cluster remains constant [6].

The ^{31}P NMR spectra of all solutions are very similar and display only one signal, assigned to free PPh_3 . In a few cases, a very low intensity signal was observed in addition to that of PPh_3 . Based on the results we obtained with the $[\text{L}_2\text{PtSn}_9]^{4-}$ species, we believe that the PPh_3 ligands have been substituted by the less bulky en.

The similarity of ^{119}Sn -NMR chemical shifts for these clusters, and the ^{31}P NMR results for $(\text{PPh}_3)_2\text{PtSn}_9$ [6] indicate that the transition metal-ligand moiety behaves in the same way in all species, contributing two electrons to the bonds within the cluster. Consequently, L_2Pt or L_2Pd moieties, where L is a two electron ligand, are expected.

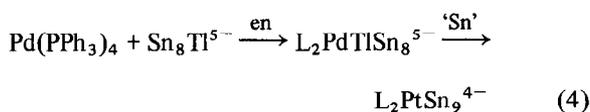
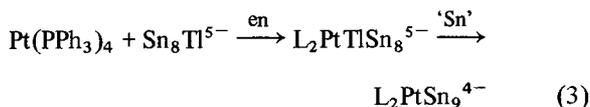
Summary

A transition metal can be incorporated into a homonuclear naked metal cluster if it is bonded to the appropriate ligands, yielding in these cases diheteronuclear cluster species (eqns. 1, 2).



Triheteronuclear cluster species have been obtained when the zero valent transition metal complexes were reacted with $\text{Sn}_8\text{TI}^{5-}$. These species react

further, probably with a Sn-containing phase to yield the diheteronuclear species (faster in the Pt than in the Pd case) (eqns. 3, 4).



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