

## Discovery of eclogite facies carbonate rocks from the Lindås Nappe, Caledonides, Western Norway

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**ABSTRACT** Eclogite facies carbonate rocks have been discovered associated with the granulite–eclogite transitional rocks within Bergen Arc system, Caledonian Orogen of western Norway. The local occurrences of marbles and calc-silicates are found subparallel to the mafic eclogite facies shear zones on Holsnøy Island. Marbles contain the assemblage calcite ( $\text{Ca}_{0.99}\text{Sr}_{0.01}\text{CO}_3$ ), calcian strontianite ( $\text{Ca}_{0.18-0.44}\text{Sr}_{0.53-0.84}\text{CO}_3$ ), clinopyroxene ( $\text{Jd}_{7-32}$ ), epidote/allanite ( $\text{Ps}_{0-33}$ ), titanite, garnet ( $\text{Alm}_{52-56}\text{Grs}_{28-33}\text{Pyp}_{11-16}$ ), barite ( $\text{Ba}_{0.90-0.99}\text{Sr}_{0.01-0.10}\text{SO}_4$ ), celestine ( $\text{Sr}_{0.67-0.98}\text{Ba}_{0.01-0.23}\text{Ca}_{0.01-0.11}\text{SO}_4$ ), and one apparently homogeneous grain of intermediate composition ( $\text{Ba}_{0.49}\text{Ca}_{0.01}\text{Sr}_{0.50}\text{SO}_4$ ). Adjacent eclogites have clinopyroxene with similar jadeite contents ( $\text{Jd}_{14-34}$ ) and similar garnet ( $\text{Alm}_{51-60}\text{Grs}_{26-36}\text{Pyp}_{8-14}$ ) compositions. The marbles have high contents of Sr (9500–11000 p.p.m) and Y (115–130 p.p.m). However, low concentrations of some key trace elements (110–160 p.p.m. Ba and <5 p.p.m. Nb) appear to indicate that the marble is not a metamorphosed carbonatite. The  $^{87}\text{Sr}/^{86}\text{Sr}$  ratios range from 0.7051 to 0.7059. Field and petrological relationships suggest that metasomatic reactions and fluids played a significant role in producing and/or modifying the marbles. The breakdown of scapolite in the granulite into carbonates and sulphates during eclogite facies metamorphism may have contributed to the metasomatic formation of the marbles along shear zones.

Fluids involved during subduction are an important catalyst for metamorphism and are recognized to have played a critical role in the localized transformation from granulite to eclogite in the Holsnøy Island area. Thermobarometry indicates 640–690 °C and 18–20 kbar for adjacent eclogites and temperatures of 580–650 °C for the calc-silicates. The marble assemblages are consistent with fluid that is dominantly comprised of  $\text{H}_2\text{O}$  ( $X_{\text{CO}_2} < 0.03$ ) under high-pressure conditions. Phase equilibria of the marbles constrain the  $f\text{O}_2$  of the fluids and imply oxidizing conditions of the deep crustal fluids. At present the source of the fluids remains unresolved. The results provide additional insights into the variable and evolving nature of fluids related to subduction and high-pressure metamorphism.

**Key words:** calcian strontianite; eclogite; fluids; marble; Norway.

### INTRODUCTION

Marbles associated with regional occurrences of eclogites provide a unique perspective for evaluating metamorphic processes and fluids associated with crustal subduction. Such marbles have been described from the Tauern Window within the Austrian Alps (Franz & Spear, 1983), from the Pan-African belt within Mali (Caby, 1994) and associated with coesite-bearing eclogites from the Dabie Mountains in central China (e.g. Wang & Liou, 1991; Schertl & Okay, 1994; Zhang & Liou, 1996). Siliceous dolomites from the Tauern window indicate an early assemblage of diopside, tremolite, dolomite, calcite, quartz and zoisite stable at fluid pressures of 18–25 kbar (Franz & Spear, 1983). Impure marbles from the Dabie Mountains contain occasional coesite in dolomite (Schertl &

Okay, 1994), coesite, quartz pseudomorphs after coesite, aragonite, and calcite pseudomorphs after aragonite occurring as fine-grained inclusions in garnet and omphacite (Wang & Liou, 1993) indicating that both the marbles and associated eclogites experienced ultra-high pressure metamorphism (> 27 kbar). Evaluation of phase equilibria for both the marbles from the Tauern Window and the Dabie Mountain indicates fluid compositions with very low  $X_{\text{CO}_2}$ , similar to those inferred in the associated eclogite facies rocks (e.g. Franz & Spear, 1983; Holland, 1979; Wang & Liou, 1993). The emerging picture of the fluid evolution during high-pressure metamorphism suggests their compositions can be both complex and highly variable (e.g. Yardley, 1997).

Marbles occur associated with the regionally exposed granulite to eclogite transitional rocks in the Lindås

Nappe of the Bergen Arcs system (BAS), western Norway. The eclogites formed locally in the granulites as a result of deformation and the introduction of fluids (Austrheim & Griffin, 1985; Austrheim, 1987, 1990). Calcite and dolomite occur as part of the equilibrium assemblage of a few eclogites (Jamtveit *et al.*, 1990). Local outcrops of marbles are also associated with the eclogites. Marbles from the Liavatn area on Holsnøy Island (Fig. 1) have calcite with average  $\delta^{18}\text{O}$  of  $9.5 \pm 0.6\text{‰}$  and  $\delta^{13}\text{C}$  of  $-4 \pm 0.8\text{‰}$  that are in equilibrium with the silicate minerals (van Wyck *et al.*, 1996). The data are interpreted to reflect retrogression due to the influx of fluids at temperatures of approximately 600 °C. There is a considerable range in estimates of composition of the eclogite fluid phase. Based on phase equilibria in the eclogites (*sensu stricto*) from Holsnøy the fluid composition during eclogite facies metamorphism was estimated to have been  $\text{H}_2\text{O}$ -rich ( $X_{\text{H}_2\text{O}} > 0.75$ ; Jamtveit *et al.*, 1990). Fluid inclusions found in quartz within eclogite facies quartz veins contain abundant  $\text{N}_2$  together with  $\text{CO}_2$ , whereas  $\text{H}_2\text{O}$  was not detected (Andersen *et al.*, 1990). Multiphase brine inclusions in omphacite are interpreted to reflect enrichment of elements incompatible with the silicate structure during consumption of  $\text{H}_2\text{O}$  by eclogite-forming hydration reactions (Svensen *et al.*, 1999, 2001), suggesting the infiltration of water-rich brines. There is no reason to assume the pervasive invasion of a single fluid. The carbonate rocks occur parallel to or along eclogite shear zones in the area. The composition of fluids in equilibrium with the carbonates may not have been the same as that in the mafic eclogite during eclogite facies metamorphism.

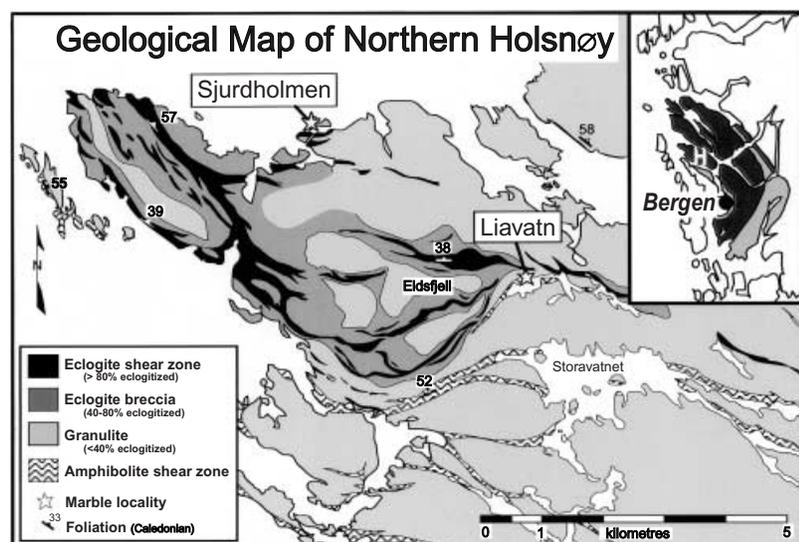
In this contribution, we present initial results from an investigation of carbonate rocks associated with the eclogite facies shear zones on Holsnøy Island within the Lindås Nappe of the BAS. The mineralogy of the carbonates includes unusual high-Sr minerals

including calcian strontianite. We document field and petrological relations of the high-pressure marbles and calc-silicates in tandem with phase equilibria and Sr isotope characteristics. The composition of clinopyroxene in the marbles and the field association demonstrate that the marbles were also subjected to eclogite facies conditions. Investigation into the phase equilibria yields information not only of the  $P$ - $T$  conditions but also the constraints on the variable fluid compositions in equilibrium with the eclogite facies rocks. The results provide insights into the possible fluid mediated metasomatic formation of the marbles concurrent with eclogite facies metamorphism during deep subduction of the continental crust.

## GEOLOGICAL AND TECTONIC FRAMEWORK

The development of the mid-Palaeozoic Scandinavian Caledonian Orogen (SCO) is related to the convergence of Laurentia (Greenland) and Baltica (Scandinavia). The SCO is characterized by a series of thrust nappes generally displaced from west to east over the Baltic Shield during the main phase of continent-continent collision (Roberts & Gee, 1985). Extensive erosion and late to postorogenic extension led to the present tectonic juxtaposition of lowermost crustal and upper crustal rocks (e.g. Andersen & Jamtveit, 1990). Within the SCO of western Norway eclogite facies rocks occur locally in the Western Gneiss Region (WGR) and within the overlying Caledonian Nappes of the Bergen Arcs system (BAS). One of the arcuate Caledonian nappes of the Bergen Arcs, the Lindås Nappe, consists of a granulite facies anorthosite complex and is interpreted to represent a slice of the lowermost continental crust now exposed over *c.* 1000 km<sup>2</sup> (Austrheim & Griffin, 1985).

The Lindås Nappe anorthosite complex records two main phases of deformation and metamorphism. Late



**Fig. 1.** Geological map of the NW Holsnøy Island area, Bergen Arcs, western Norway. Eclogites occur locally within the granulites as diffuse patches and along fractures and shear zones. The major eclogite facies shear zones illustrated on the map form an anastomosing network within the partially eclogitized granulites. Marble and related eclogite samples for this study are from the Liavatn and Sjurdholmen localities.

Proterozoic (Sveconorwegian) garnet–granulite facies metamorphism and deformation occurred at deep crustal levels followed by a second period of deformation and metamorphism during the Caledonian Orogeny that brought the granulites to eclogite facies conditions (Austrheim & Griffin, 1985). However, not all the granulite converted to eclogite. The localized transition from granulite to eclogite facies rocks took place over a few centimetres apparently as a result of deformation and the infiltration of fluids. The eclogites are best preserved on Holsnøy Island, in the north-western part of the Lindås Nappe. There the eclogites formed along fractures within the granulites, along shear zones up to 120 m thick and continuous for several kilometres, and in breccias where angular blocks of granulite are surrounded by anastomosing, well-foliated eclogite (Austrheim, 1987; Klaper, 1991; Boundy *et al.*, 1992). In addition, significant volumes of the partially eclogitized granulites contain eclogites that are apparently unrelated to any high strain zone. Such granulites contain isolated patches of eclogite and are presumably related to the infiltration of fluids along microfractures (Fig. 1). In addition to the map units, several different eclogite occurrences have recently been recognized, though at a scale that prevents their inclusion as map units. Austrheim & Boundy (1994) documented eclogitic pseudotachylytes related to the eclogite facies metamorphism. Marbles and calc-silicates are associated with the eclogites in a few isolated areas (Boundy, 1995; Boundy *et al.*, 1997a).

### FIELD RELATIONS OF THE MARBLES

Marbles occur associated with the eclogites and eclogite facies shear zones on Holsnøy Island. The area consists of granulites, partially eclogitized granulites and eclogites (Fig. 1). On the map scale these eclogites are depicted as major eclogite facies shear zones. Liavatn and Sjurholmen are the two most extensive marble localities discovered thus far (Fig. 1). The marble locality at Liavatn was mapped in detail (1:5000; Fig. 2). Unfortunately, the marbles are not well exposed, and thus the mapping depicts their minimum extent. At the Liavatn locality, marbles occur in a zone that parallels an eclogite facies shear zone, which is variably retrograded to amphibolite facies (Fig. 2). The eclogite shear zone and subparallel marble bodies occur within a large body of granulite facies anorthosites and are associated locally with high Fe–Ti ultramafic lenses. The zone of marbles is approximately 400 m long and up to 50 m thick, although the thickness is quite variable. Despite the poor exposure, it appears that the marble zone is severely attenuated towards the west (Fig. 2). The northern boundary of the marbles is formed by an eclogite zone that lies structurally above the marbles. Locally marble is intercalated with the eclogite shear zone along layers parallel to the foliation (Fig. 3). Retrogressed anorthosite shear zone rocks, termed ‘paper schist’ due to the extremely thin foliation layers, occur along the southern boundary, structurally below the marbles, and also occur locally along the northern boundary. The retrograde shear zones are amphibolite and lower grade (Fig. 2).

The marbles at the Liavatn locality are strongly deformed. At the macroscopic level, the foliation is defined by epidote/pyroxene rich layers within the carbonate matrix (Fig. 3). Generally this foliation is subparallel to the bounding eclogite and retrograde shear zones. In several areas the foliation is folded. Lenses of eclogite occur locally within the marbles and are typically 5–30 cm in diameter. Other mafic layers within the carbonates are comprised of mainly biotite

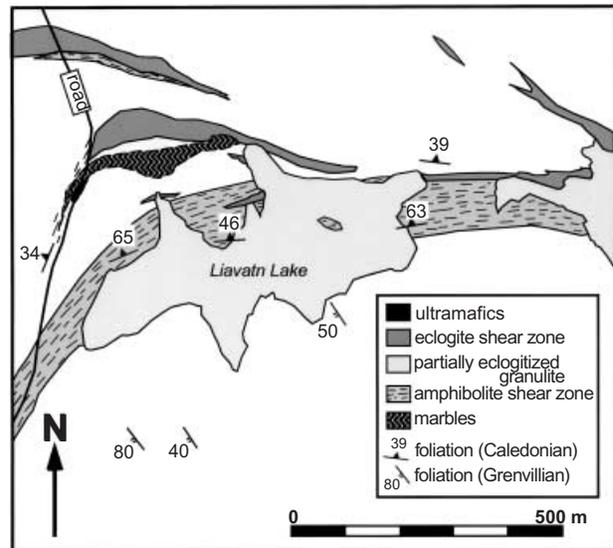


Fig. 2. Detailed map of the Liavatn marble locality on NW Holsnøy Island area. Locally the marble includes layers and boudins of eclogite and anorthosite ‘paper schist’.



Fig. 3. Field photograph of marble interlayered with eclogite at the Liavatn locality. The foliation of the marble is defined by epidote and pyroxene-rich layers and parallels that in the adjacent well-foliated eclogite. Part of a Swiss army knife is shown for scale (left edge of photo).

and titanite and may represent retrograded eclogite lenses. Some of the eclogite layers adjacent to the marbles contain veins of quartz, omphacite, clinozoisite and calcite. The veins are discontinuous and are typically on the order of 1 cm thick and 10 cm long. One eclogite sample at Liavatn displays cavities with carbonates and/or clinopyroxene crystals protruding into the cavity. The veins and cavities suggest that fluid pressure approximated load pressure in these rocks.

At Sjurholmen, an island off the NW coast of Holsnøy, the marbles occur in layers from <5 cm to 3 m thick (Fig. 1). Along the shoreline, these outcrops are well exposed. Thin carbonate bands are interlayered with the eclogites and retrogressed eclogites. The eclogites contain centimetre thick, discontinuous quartz–muscovite veins. Metre thick marble layers are parallel to the general foliation in the area. A quartz layer on the order of 50 cm thick lies structurally above the thick marble layers. Eclogites occur as boudins or lenses within the quartz layer. Retrogressed mangerites bound the package of eclogite, marble and quartz layer, both structurally above and below.

The protolith to the eclogites in this area appears to be predominantly mangerite, which at this locality is heavily retrogressed (to amphibolite facies or lower grade). At the macroscopic scale, epidote/pyroxene-rich layers define the foliation of these highly deformed marbles, similar to those at Liavtn. The foliation is generally sub-parallel to that in the adjacent layers, although some folding also occurs.

In many cases the protolith for the eclogites may be deduced from the field relationships, and are typically granulites of anorthositic to gabbroic composition, although jotunite, mangerite and locally ultramafic granulite protoliths have also been found (Austrheim & Griffin, 1985; Austrheim, 1990). For most of the amphibolite and greenschist facies shear zones it is likewise the case that the protolith can be deduced from the relationships observed in the field. However, given the minor and localized occurrence of marbles studied, a protolith is not evident based solely on the field relationships. The Liavtn marble body is located entirely within a large anorthosite

body, comprised of anorthosites, gabbros and, locally, ultramafic rocks.

## DESCRIPTION OF ASSEMBLAGES

Samples were collected from the two localities shown (Fig. 1); mineral assemblages are given in Table 1 and mineral analyses in Tables 2–10. The marbles have a fairly uniform grain size (0.3–1.0 mm). At both the Liavtn and Sjurholmen localities the marbles consist predominantly of calcite along with minor epidote/allanite, titanite, apatite and celestine. Typically, calcite appears homogeneous with sharp grain boundaries. At the microscopic scale some calcite, however, has oriented inclusions of celestine (Fig. 4a). Clinopyroxene occurs in Liavtn marble samples BA9/91 A' (= BA10/92) and as inclusions in garnet porphyroblasts from sample BA21/93. Clinopyroxene shows symplectitic breakdown textures, in part to hornblende. Most of the clinopyroxene appears unaffected by this retrogression. Clinopyroxene in marbles from both the Liavtn and Sjurholmen localities displays bright emerald green pleochroism. Garnet occurs in one of the calc-silicate samples from the Liavtn locality (BA21/93). The subhedral grains of garnet are full of inclusions, mainly carbonate with omphacitic clinopyroxene and epidote. The sample also includes areas of symplectite, which in part is hornblende, but no relict clinopyroxene was found outside of the garnet porphyroblasts in this sample. Epidote and allanite from the marbles show dramatic zonation (Fig. 4b). The eclogite samples from both localities consist of omphacite, garnet, rutile, titanite, calcite and quartz. The omphacite shows symplectitic breakdown textures mainly along the grain boundaries. Some mafic layers within the marbles consist of rutile with titanite overgrowths, garnet and symplectite with relict omphacite and are interpreted to represent a lens of retrograded eclogite within the marble.

**Table 1.** Description of the marble and adjacent eclogite samples.

| Sample                     | Mineral assemblage  | Temperature estimate |                      |
|----------------------------|---|----------------------|----------------------|
|                            |   | cc-dol <sup>1</sup>  | grt-cpx <sup>2</sup> |
| <b>Liavtn locality</b>     |   |                      |                      |
| HA4/90 A                   | marble cc, ep, ttn, ap, cel, ms                             | ≥460 °C              |                      |
| HA5/90                     | marble cc, ep, ttn, ap, cel, ms                             | ≥390 °C              |                      |
| BA9/91 A                   | marble cc, ttn, ep, cel, bar (dol)                          | ≥490 °C              |                      |
| BA9/91 A'                  | marble cc, ttn, omp, ep, cel, bar                           | ≥460 °C              |                      |
| BA9/92 A                   | calc-silicate cc, ep, ms, ttn                               | ≥450 °C              |                      |
| BA9/92B                    | marble cc, ep, ttn, ap, bt, ms (chl)                        | –                    |                      |
| BA21/93                    | calc-silicate cc, ep, grt, ms, rut, ap, omp (symp)          | ≥620 °C              | 580–650 °C           |
| Liavtn                     | eclogite omp, grt, ru, ttn, cc, ms                          | –                    | 640–690 °C           |
| <b>Sjurholmen locality</b> |   |                      |                      |
| Sample                     | Mineral assemblage  |                      |                      |
| BA25/92 A                  | marble cc, cpx, all, ttn, ap, cel, bar, bt, op (hbd)        | ≥350 °C              |                      |
| BA25/92B                   | marble cc, cpx, all, ttn, ap, bt, cel, op, ilm? (hbd)       | ≥440 °C              |                      |
| BA69/93                    | eclogite cpx, grt, all, cc, ttn, ru, ap (symp)              | –                    | 595 °C               |
| BA70/93                    | calc-silicate cpx, cc, ttn, ap, grt, qtz, alb, bar, zrc, ep | ≥280 °C              |                      |

Abbreviations: all, allanite; ap, apatite; bar, barite; bt, biotite; cc, calcite; cel, celestine; chl, chlorite; cpx, clinopyroxene; dol, dolomite; epd, epidote; grt, garnet; hbd, hornblende; ilm, ilmenite; mu, muscovite; omp, omphacite; plg, plagioclase; op, opaque; qz, quartz; ru rutile; symp, symplectite; ttn, titanite; zrc, zircon. () indicates retrogressive phase.

<sup>1</sup> Anovitz & Essene (1987).

<sup>2</sup> Krogh Ravna (2000).

## ANALYTICAL METHODS

The minerals were studied by back-scattered electron (BSE) imaging and analysed (Tables 2–10) with a Cameca CAMEBAX electron microprobe using a wavelength dispersive system and a PAP correction routine at the University of Michigan. The BSE images were used as a guide to detect potential compositional zonation in the phases and to aid in determining the areas of interest for analysis. Some of the BSE images are shown in Fig. 4(a,b). All samples were analysed at an accelerating voltage of 15 kV and a beam current of 10 nA with the exception of some carbonates for which the beam current was reduced to 7 nA to minimize elemental migration. The elements F, Na, Cl and K were measured at the beginning of each analysis to minimize errors associated with their migration. The F and Cl were measured for

**Table 2.** Carbonate analyses.

|                            | HA4/90 A<br>5/1 | HA5/90<br>4/1 | BA9/91 A<br>4/1 | BA9/91 A'<br>1/1 | BA9/92<br>2/2 | BA21/93<br>3/1 | BA21/93<br>2/3* | BA25/92 A<br>1/2 | BA25/92B<br>3/5 | BA70/93<br>3/2 |
|----------------------------|-----------------|---------------|-----------------|------------------|---------------|----------------|-----------------|------------------|-----------------|----------------|
| FeO                        | 1.55            | 1.07          | 1.73            | 1.60             | 1.25          | 3.11           | 3.44            | 1.10             | 1.75            | 0.72           |
| MnO                        | 0.13            | 0.22          | 0.18            | 0.19             | 0.26          | 0.19           | 0.00            | 0.33             | 0.36            | 0.51           |
| SrO                        | 0.65            | 1.14          | 0.80            | 0.69             | 0.00          | 0.00           | 1.21            | 0.00             | 0.24            | 0.00           |
| MgO                        | 1.38            | 0.91          | 1.39            | 1.35             | 1.14          | 2.90           | 3.14            | 0.67             | 0.96            | 0.52           |
| CaO                        | 50.64           | 51.42         | 50.65           | 51.06            | 52.85         | 50.04          | 46.30           | 53.67            | 51.63           | 53.29          |
| CO <sub>2</sub>            | 45.65           | 45.24         | 45.65           | 45.11            | 43.65         | 44.45          | 45.91           | 43.72            | 42.96           | 43.14          |
| Total                      | 100.01          | 100.00        | 100.40          | 100.00           | 99.15         | 100.69         | 100.00          | 99.49            | 97.89           | 98.18          |
| Mg                         | 0.04            | 0.02          | 0.04            | 0.03             | 0.03          | 0.07           | 0.08            | 0.02             | 0.02            | 0.01           |
| Fe                         | 0.00            | 0.00          | 0.02            | 0.00             | 0.02          | 0.04           | 0.05            | 0.02             | 0.02            | 0.01           |
| Mn                         | 0.02            | 0.02          | 0.00            | 0.02             | 0.00          | 0.00           | 0.00            | 0.00             | 0.01            | 0.01           |
| Sr                         | 0.01            | 0.01          | 0.01            | 0.01             | 0.00          | 0.00           | 0.01            | 0.00             | 0.00            | 0.00           |
| Ca                         | 0.93            | 0.95          | 0.93            | 0.93             | 0.94          | 0.89           | 0.81            | 0.96             | 0.94            | 0.95           |
| T (°C) cc-dol <sup>1</sup> | 460             | 390           | 490             | 460              | 450           | 620            | –               | 350              | 440             | 280            |

<sup>1</sup> Anovitz & Essene (1987).

\* included in garnet.

**Table 3.** Orthorhombic carbonate analyses.

| Sample/Point    | BA9/91 A 1 | BA9/91 A 2 | BA9/91 A 4 | BA9/91 A 5 | BA9/91 A 6 | BA9/91 A 7 | BA9/91 A 8 | BA9/91 A 9 |
|-----------------|------------|------------|------------|------------|------------|------------|------------|------------|
| CaO             | 7.12       | 16.20      | 15.41      | 18.20      | 19.16      | 18.96      | 19.47      | 13.86      |
| SrO             | 61.58      | 48.99      | 50.16      | 45.29      | 47.15      | 44.42      | 45.89      | 52.29      |
| MgO             | 0.00       | 0.00       | 0.00       | 0.00       | 0.00       | 0.00       | 0.00       | 0.00       |
| BaO             | 0.05       | 0.09       | 0.03       | 0.00       | 0.00       | 0.00       | 0.00       | 0.00       |
| PbO             | 0.01       | 0.00       | 0.00       | 0.00       | 0.00       | 0.00       | 0.00       | 0.00       |
| FeO             | 0.12       | 0.51       | 0.08       | 0.01       | 0.03       | 0.01       | 0.06       | 0.06       |
| MnO             | 0.04       | 0.02       | 0.01       | 0.00       | 0.00       | 0.00       | 0.00       | 0.00       |
| CO <sub>2</sub> | 30.82      | 34.29      | 34.61      | 37.52      | 33.19      | 37.58      | 34.51      | 34.01      |
| Total           | 99.74      | 100.10     | 100.29     | 101.02     | 99.53      | 100.98     | 99.93      | 100.22     |
| Ca              | 0.18       | 0.37       | 0.36       | 0.40       | 0.44       | 0.42       | 0.44       | 0.32       |
| Sr              | 0.83       | 0.61       | 0.63       | 0.54       | 0.59       | 0.53       | 0.56       | 0.66       |
| Mg              | 0.00       | 0.00       | 0.00       | 0.00       | 0.00       | 0.00       | 0.00       | 0.00       |
| Ba              | 0.00       | 0.00       | 0.00       | 0.00       | 0.00       | 0.00       | 0.00       | 0.00       |
| Pb              | 0.00       | 0.00       | 0.00       | 0.00       | 0.00       | 0.00       | 0.00       | 0.00       |
| Fe              | 0.00       | 0.01       | 0.00       | 0.00       | 0.00       | 0.00       | 0.00       | 0.00       |
| Mn              | 0.00       | 0.00       | 0.00       | 0.00       | 0.00       | 0.00       | 0.00       | 0.00       |
| Str             | 0.82       | 0.61       | 0.64       | 0.57       | 0.57       | 0.56       | 0.56       | 0.67       |
| Ar              | 0.18       | 0.38       | 0.36       | 0.43       | 0.43       | 0.44       | 0.44       | 0.33       |

Selected orthorhombic carbonate analyses from marble sample BA9/91 A. Str = strontianite, Ar = aragonite.

**Table 4.** Diffraction pattern for orthorhombic carbonate in calcite matrix.

| Sin theta | d Å    | Cc Int | Ar Int | hkl |
|-----------|--------|--------|--------|-----|
| 0.3971    | 3.807  | 10     |        | 012 |
| 0.4337    | 3.480  |        | 100    | 111 |
| 0.4462    | 3.381  |        | 60     | 021 |
| 0.4955    | 3.037  | 100    |        | 104 |
| 0.5344    | 2.810  | 5      |        | 006 |
| 0.5934    | 2.523  |        | 20     | 200 |
| 0.5990    | 2.499  | 10     |        | 110 |
| 0.6115    | 2.446  |        | 20     | 112 |
| 0.6388    | 2.339  | 15     |        | 113 |
| 0.7431    | 1.9998 | 7      |        | 202 |
| 0.9063    | 1.6263 | 4      |        | 211 |

Cc Int = intensity of calcite lines; Ar Int = intensity of aragonite lines.

60 s, the LREE, Th, Y and Zr were all measured for 100 s, and all other elements were measured for a count time of 30 s. The carbonates and micas were analysed using a rastered beam (3  $\mu\text{m}^2$ ); all other minerals were analysed with a focused beam in spot mode.

Standards for La, Ce, Pr, Nd, Sm and Gd were pure synthetic phosphates of each element; the Th standard was natural thorite and the Y standard was a synthetic yttrium aluminium garnet (YAG:  $\text{Y}_3\text{Al}_5\text{O}_{12}$ ). The data for LREE were corrected for interelement interferences based on the methods and analyses presented in DeWolf (1993). Wavelength scans were used to detect the presence of minor elements. Relative errors are based on counting statistics for the standards and samples; for epidote and titanite these are *c.* 1% of amount present (OAP) for Si, Al and Ca, and *c.* 5–15% OAP for Sr, Th, Y and LREEs. In general the larger errors are associated with lower abundances of the minor elements and of LREEs.

Whole rock chemistry was analysed using a Philips PW 1400 XRF spectrometer equipped with a Sc-Mo tube at the University of Lausanne, Switzerland. For the Sr isotope analyses small fragments (1–5 mg) of carbonate were separated and dissolved in HCl. Subsequently Sr was separated with 2.5 N HCl from the matrix columns filled with *c.* 15 mL DOWEX AG 50Wx8<sup>®</sup> (200–400 mesh) resin. Strontium was measured on W-single central filaments in the dynamic mode. The sample was loaded with a slurry containing TaF<sub>5</sub>. Total procedural blank concentrations are negligible with less than 50 pg. Replicate measurements of NBS 987 standard yielded a mean <sup>87</sup>Sr/<sup>86</sup>Sr value of 0.710265 ± 20. Fractionation was corrected by

normalizing the Sr-isotope ratios to <sup>86</sup>Sr/<sup>88</sup>Sr = 0.1194. All isotope measurements were performed at the Zentrallaboratorium für Geochronologie at the Universität Münster.

## MINERAL COMPOSITIONS

### Carbonates

Calcite is the major phase in the marbles. Typically calcite within an individual sample is homogenous but there are slight compositional differences between samples (Table 2, Fig. 4a). Calcite contains 0.2–1.2 mole percentage SrCO<sub>3</sub>. Calcite from the Sjurdholmen samples is more enriched in Fe than those from Liavatn. Minor dolomite was found in only one marble sample (BA9/91 A). The dolomite occurs as < 1  $\mu\text{m}$  blebs in the calcite, and thus could not be analysed accurately with the electron microprobe. Carbonate inclusions in garnet from sample BA21/93 contain up to several weight percentage more Mg, Fe, and Sr than the adjacent calcite groundmass (Table 2).

One marble (BA9/91 A) contains minor amounts of a second Ca–Sr rich carbonate that appears as small blebs throughout the groundmass and typically near internal fractures in the rock. The composition of this carbonate is intermediate between aragonite (Ar) and strontianite (Str), ranging from Str<sub>82</sub> to Str<sub>56</sub> (Table 3). Figure 5 is a ternary plot of these carbonates as well as other data from the literature. The compositions from this study confirm a greater extent of solid solution along the CaCO<sub>3</sub>–SrCO<sub>3</sub> binary than was observed by Speer (1983), as was also concluded by Vorob'yev *et al.* (1989) and Pisarskii *et al.* (1998). The carbonate analyses from Vorob'yev *et al.* (1989) and Pisarskii *et al.* (1998) are suspect, however, because they appear to be finely intergrown with other carbonates, some showing complex exsolution textures. Additionally, they have not been adequately characterized structurally.

|                                      | BA9/91 A'<br>cpx1/4 | Liavatn<br>cpx1/1 | BA21/93<br>cpx 5 | BA21/93<br>cpx 9 | BA25/92 A<br>cpx1/2 | BA25/92B<br>cpx3/1 | BA69/93<br>cpx2/3 | BA70/93<br>cpx1/2 |
|--------------------------------------|---------------------|-------------------|------------------|------------------|---------------------|--------------------|-------------------|-------------------|
| SiO <sub>2</sub>                     | 53.21               | 53.08             | 53.67            | 52.79            | 52.32               | 52.74              | 54.28             | 52.61             |
| TiO <sub>2</sub>                     | 0.31                | 0.22              | 0.16             | 0.31             | 0.07                | 0.04               | 0.26              | 0.07              |
| Al <sub>2</sub> O <sub>3</sub>       | 7.88                | 7.92              | 8.27             | 8.66             | 1.94                | 2.20               | 8.29              | 3.70              |
| Cr <sub>2</sub> O <sub>3</sub>       | 0.05                | 0.01              | 0.00             | 0.02             | 0.04                | 0.04               | 0.00              | 0.00              |
| Fe <sub>2</sub> O <sub>3</sub>       | 3.24                | 4.97              | 6.02             | 7.76             | 3.85                | 4.71               | 6.53              | 4.55              |
| FeO                                  | 6.88                | 5.18              | 4.61             | 3.72             | 9.70                | 8.18               | 5.97              | 8.61              |
| MnO                                  | 0.09                | 0.06              | 0.06             | 0.04             | 0.25                | 0.23               | 0.08              | 0.24              |
| MgO                                  | 7.53                | 7.69              | 8.13             | 7.78             | 9.15                | 9.07               | 5.66              | 8.55              |
| CaO                                  | 15.65               | 14.89             | 14.14            | 14.19            | 20.13               | 19.11              | 11.64             | 18.11             |
| Na <sub>2</sub> O                    | 5.06                | 5.53              | 5.84             | 5.95             | 2.28                | 3.03               | 7.35              | 3.38              |
| Total                                | 99.89               | 99.55             | 100.90           | 101.22           | 99.74               | 99.34              | 100.07            | 99.82             |
| Calculated on the basis of 4 cations |                     |                   |                  |                  |                     |                    |                   |                   |
| Si                                   | 1.96                | 1.95              | 1.94             | 1.91             | 1.98                | 1.99               | 1.99              | 1.98              |
| Al <sup>IV</sup>                     | 0.04                | 0.05              | 0.06             | 0.09             | 0.02                | 0.01               | 0.01              | 0.02              |
| Al <sup>VI</sup>                     | 0.30                | 0.29              | 0.29             | 0.28             | 0.07                | 0.09               | 0.34              | 0.14              |
| Fe <sup>+3</sup>                     | 0.09                | 0.14              | 0.17             | 0.21             | 0.11                | 0.13               | 0.18              | 0.13              |
| Ti                                   | 0.01                | 0.01              | 0.00             | 0.01             | 0.00                | 0.00               | 0.01              | 0.00              |
| Cr                                   | 0.00                | 0.00              | 0.00             | 0.00             | 0.00                | 0.00               | 0.00              | 0.00              |
| Mg                                   | 0.41                | 0.42              | 0.44             | 0.42             | 0.52                | 0.51               | 0.31              | 0.48              |
| Fe <sup>+2</sup>                     | 0.21                | 0.16              | 0.14             | 0.11             | 0.31                | 0.26               | 0.18              | 0.27              |
| Mn                                   | 0.00                | 0.00              | 0.00             | 0.00             | 0.01                | 0.01               | 0.00              | 0.01              |
| Ca                                   | 0.62                | 0.59              | 0.55             | 0.55             | 0.82                | 0.77               | 0.46              | 0.73              |
| Na                                   | 0.36                | 0.39              | 0.41             | 0.42             | 0.17                | 0.22               | 0.52              | 0.25              |
| Aug                                  | 0.63                | 0.60              | 0.57             | 0.57             | 0.83                | 0.77               | 0.47              | 0.75              |
| Jd                                   | 0.30                | 0.29              | 0.21             | 0.22             | 0.07                | 0.09               | 0.34              | 0.14              |
| Aeg                                  | 0.06                | 0.10              | 0.17             | 0.21             | 0.10                | 0.13               | 0.18              | 0.11              |

Abbreviations: Aeg = aegirine; Aug = augite; c, core; cpx, clinopyroxene; Jd = jadeite; r, rim.

|                                      | BA21/93<br>gar1/3c | BA21/93<br>gar1/11r | BA21/93<br>gar2/3r | BA21/93<br>gar2/5c | Liavatn<br>gar1/1c | Liavatn<br>gar1/3r | BA 69/93<br>gar1/1c | BA 69/93<br>gar1/2r |
|--------------------------------------|--------------------|---------------------|--------------------|--------------------|--------------------|--------------------|---------------------|---------------------|
| SiO <sub>2</sub>                     | 37.91              | 37.74               | 38.23              | 38.05              | 37.92              | 38.33              | 37.55               | 37.63               |
| TiO <sub>2</sub>                     | 0.38               | 0.08                | 0.10               | 0.46               | 0.14               | 0.07               | 0.05                | 0.05                |
| Al <sub>2</sub> O <sub>3</sub>       | 20.65              | 20.58               | 21.25              | 20.84              | 20.46              | 21.11              | 20.42               | 20.60               |
| Cr <sub>2</sub> O <sub>3</sub>       | 0.01               | 0.02                | 0.02               | 0.03               | 0.00               | 0.00               | 0.00                | 0.00                |
| FeO                                  | 24.64              | 26.08               | 25.43              | 24.31              | 26.30              | 24.99              | 27.99               | 26.36               |
| MnO                                  | 1.17               | 1.04                | 0.87               | 1.04               | 0.98               | 0.82               | 2.86                | 2.79                |
| MgO                                  | 2.94               | 2.99                | 4.17               | 3.79               | 3.05               | 3.55               | 2.09                | 2.14                |
| CaO                                  | 11.66              | 10.76               | 10.07              | 11.38              | 11.24              | 11.18              | 9.57                | 9.87                |
| Total                                | 99.38              | 99.29               | 100.13             | 99.90              | 100.08             | 100.04             | 100.52              | 99.45               |
| Calculated on the basis of 12 oxygen |                    |                     |                    |                    |                    |                    |                     |                     |
| Si                                   | 3.01               | 3.01                | 2.99               | 2.98               | 3.00               | 3.01               | 2.99                | 3.01                |
| Al                                   | 1.93               | 1.93                | 1.96               | 1.93               | 1.91               | 1.95               | 1.92                | 1.94                |
| Ti                                   | 0.02               | 0.00                | 0.01               | 0.03               | 0.01               | 0.00               | 0.00                | 0.00                |
| Cr                                   | 0.00               | 0.00                | 0.00               | 0.00               | 0.00               | 0.00               | 0.00                | 0.00                |
| Mg                                   | 0.35               | 0.35                | 0.49               | 0.44               | 0.36               | 0.42               | 0.25                | 0.26                |
| Fe <sup>+2</sup>                     | 1.63               | 1.74                | 1.66               | 1.59               | 1.74               | 1.64               | 1.87                | 1.76                |
| Mn                                   | 0.08               | 0.07                | 0.06               | 0.07               | 0.07               | 0.05               | 0.19                | 0.19                |
| Ca                                   | 0.99               | 0.92                | 0.84               | 0.96               | 0.95               | 0.94               | 0.82                | 0.85                |
| O                                    | 12.00              | 12.00               | 12.00              | 12.00              | 12.00              | 12.00              | 12.00               | 12.00               |
| Alm                                  | 53.55              | 56.40               | 53.69              | 51.23              | 55.80              | 53.76              | 59.71               | 57.74               |
| Grs                                  | 32.43              | 29.76               | 28.16              | 31.76              | 30.54              | 30.83              | 26.16               | 27.70               |
| Pyr                                  | 11.39              | 11.51               | 16.24              | 14.70              | 11.55              | 13.63              | 7.94                | 8.36                |
| Sps                                  | 2.58               | 2.28                | 1.91               | 2.30               | 2.11               | 1.78               | 6.19                | 6.19                |
| Uv                                   | 0.04               | 0.06                | 0.03               | 0.07               | 0.00               | 0.00               | 0.00                | 0.00                |
| Mg/Mg + Fe                           | 0.18               | 0.17                | 0.23               | 0.22               | 0.17               | 0.20               | 0.12                | 0.13                |

Abbreviations: Alm, almandine; c, core; Grs, grossular; Pyr, pryrope; r, rim; Sps, spessartine; Uv, uvarovite.

**Table 5.** Clinopyroxene analyses.

**Table 6.** Garnet analyses.

Single crystal XRD analysis was performed on a grain of the intermediate carbonate from sample BA9/91 A using a Gandolfi camera to determine cell parameters. Table 4 shows the diffraction pattern for a grain of intermediate composition Str<sub>60</sub>, as well as calcite peaks, which appear in the pattern because the orthorhombic carbonate could not be completely removed from the matrix. The *d*-values for the two strongest lines of the intermediate carbonate (111 &

021) lie between the *d*-values for aragonite and strontianite (Table 4), 60% toward strontianite. The experiments of Carlson (1980) suggest that Str<sub>56</sub>–Str<sub>67</sub> is not stable with calcite at the peak conditions experienced by these marbles. Dasgupta *et al.* (1990) also noted the occurrence of calcian strontianite with compositions of Str<sub>71</sub>–Str<sub>61</sub> in equilibrium with calcite and barytocalcite in metamorphic rocks (Fig. 5) and concluded that they do not fit with experimentally

Table 7. Epidote and allanite analyses.

|                                | HA4/90<br>2/1 | HA5/90<br>3/1 | BA9/91A<br>3/1 | BA9/91A'<br>1/1c | BA9/92A<br>1/2 | BA9/92B<br>1/1 | BA21/93<br>1/1 | BA21/93<br>1/8 | BA25/92A<br>1/1 | BA25/92B<br>1/14r | BA69/93<br>2/1 |
|--------------------------------|---------------|---------------|----------------|------------------|----------------|----------------|----------------|----------------|-----------------|-------------------|----------------|
| SiO <sub>2</sub>               | 36.34         | 36.35         | 37.84          | 37.47            | 38.20          | 38.02          | 34.57          | 36.92          | 32.87           | 34.02             | 35.03          |
| TiO <sub>2</sub>               | 0.10          | 0.17          | 0.36           | 0.33             | 0.11           | 0.24           | 0.28           | 0.34           | 0.23            | 0.28              | 0.25           |
| Al <sub>2</sub> O <sub>3</sub> | 25.60         | 24.56         | 26.53          | 25.70            | 27.43          | 26.63          | 23.38          | 26.59          | 19.42           | 20.14             | 22.05          |
| Fe <sub>2</sub> O <sub>3</sub> | 7.49          | 3.43          | 0.28           | 0.24             | -0.03          | 0.14           | 9.57           | 8.17           | 6.67            | 4.91              | 4.48           |
| FeO                            | 0.82          | 5.15          | 7.46           | 7.91             | 6.64           | 6.21           | 0.00           | 0.00           | 6.34            | 7.52              | 6.60           |
| MnO                            | 0.08          | 0.07          | 0.07           | 0.09             | 0.00           | 0.05           | 0.06           | 0.09           | 0.07            | 0.15              | 0.16           |
| MgO                            | 0.23          | 0.36          | 0.20           | 0.21             | 0.09           | 0.20           | 0.49           | 0.18           | 0.53            | 0.43              | 0.49           |
| CaO                            | 19.97         | 18.89         | 21.86          | 21.75            | 22.98          | 21.70          | 18.01          | 21.49          | 15.55           | 16.71             | 17.40          |
| SrO                            | 2.56          | 1.98          | 1.61           | 1.53             | 2.03           | 1.58           | 2.00           | 2.15           | 0.88            | 0.72              | 1.05           |
| Y <sub>2</sub> O <sub>3</sub>  | 0.17          | 0.10          | 0.00           | 0.00             | 0.06           | 0.00           | 0.00           | 0.00           | 0.00            | 0.02              | 0.10           |
| La <sub>2</sub> O <sub>3</sub> | 1.34          | 2.33          | 0.15           | 0.33             | 0.07           | 0.30           | 2.53           | 0.12           | 3.65            | 2.71              | 2.30           |
| Ce <sub>2</sub> O <sub>3</sub> | 2.18          | 2.99          | 0.52           | 0.67             | 0.00           | 0.50           | 3.62           | 0.29           | 6.71            | 4.89              | 4.22           |
| Pr <sub>2</sub> O <sub>3</sub> | 0.08          | 0.48          | 0.07           | 0.00             | 0.04           | 0.00           | 0.98           | 0.08           | 0.72            | 0.44              | 0.43           |
| Nd <sub>2</sub> O <sub>3</sub> | 0.65          | 0.90          | 0.39           | 0.29             | 0.05           | 0.07           | 1.19           | 0.21           | 2.12            | 1.78              | 2.24           |
| Sm <sub>2</sub> O <sub>3</sub> | 0.00          | 0.10          | 0.07           | 0.01             | 0.00           | 0.00           | 0.56           | 0.04           | 0.36            | 0.60              | 0.33           |
| Gd <sub>2</sub> O <sub>3</sub> | 0.16          | 0.41          | 0.23           | 0.02             | 0.00           | 0.00           | 0.23           | 0.18           | 0.00            | 0.00              | 0.03           |
| ThO <sub>2</sub>               | 0.12          | 0.00          | 0.06           | 0.01             | 0.12           | 0.00           | 0.00           | 0.00           | 0.05            | 0.09              | 0.30           |
| ZrO <sub>2</sub>               | 0.00          | 0.04          | 0.05           | 0.00             | 0.00           | 0.02           | 0.00           | 0.00           | 0.00            | 0.01              | 0.00           |
| F                              | 0.00          | 0.05          | 0.04           | 0.00             | 0.00           | 0.02           | 0.00           | 0.00           | 0.10            | 0.01              | 0.03           |
| Total                          | 97.91         | 98.35         | 97.78          | 96.55            | 97.80          | 95.70          | 97.45          | 96.84          | 96.27           | 95.46             | 97.50          |
| Calc. on basis of 8 cations    |               |               |                |                  |                |                |                |                |                 |                   |                |
| Si                             | 2.97          | 2.99          | 2.99           | 3.00             | 2.99           | 3.05           | 3.00           | 3.00           | 2.93            | 2.99              | 2.98           |
| Ti                             | 0.01          | 0.01          | 0.02           | 0.02             | 0.01           | 0.01           | 0.02           | 0.02           | 0.02            | 0.02              | 0.02           |
| Al                             | 2.46          | 2.38          | 2.47           | 2.43             | 2.53           | 2.52           | 2.39           | 2.55           | 2.04            | 2.09              | 2.21           |
| Fe <sup>+3</sup>               | 0.41          | 0.41          | 0.48           | 0.49             | 0.43           | 0.35           | 0.62           | 0.50           | 0.66            | 0.55              | 0.49           |
| Fe <sup>+2</sup>               | 0.12          | 0.15          | 0.03           | 0.05             | 0.00           | 0.08           | 0.00           | 0.00           | 0.25            | 0.32              | 0.26           |
| Mg                             | 0.01          | 0.00          | 0.00           | 0.01             | 0.00           | 0.00           | 0.06           | 0.02           | 0.01            | 0.01              | 0.01           |
| Mn                             | 0.03          | 0.04          | 0.02           | 0.02             | 0.01           | 0.02           | 0.00           | 0.01           | 0.07            | 0.06              | 0.06           |
| Ca                             | 1.75          | 1.67          | 1.85           | 1.87             | 1.93           | 1.86           | 1.67           | 1.87           | 1.49            | 1.57              | 1.59           |
| Sr                             | 0.12          | 0.09          | 0.07           | 0.07             | 0.09           | 0.10           | 0.10           | 0.10           | 0.05            | 0.04              | 0.05           |
| Y                              | 0.01          | 0.00          | 0.00           | 0.00             | 0.00           | 0.00           | 0.00           | 0.00           | 0.00            | 0.00              | 0.00           |
| REE                            | 0.14          | 0.23          | 0.04           | 0.04             | 0.00           | 0.03           | 0.29           | 0.03           | 0.46            | 0.35              | 0.31           |
| Th                             | 0.00          | 0.00          | 0.00           | 0.00             | 0.00           | 0.00           | 0.00           | 0.00           | 0.00            | 0.00              | 0.01           |
| Zr                             | 0.00          | 0.00          | 0.00           | 0.00             | 0.00           | 0.00           | 0.00           | 0.00           | 0.00            | 0.00              | 0.00           |
| F                              | 0.00          | 0.01          | 0.01           | 0.00             | 0.00           | 0.01           | 0.00           | 0.00           | 0.03            | 0.00              | 0.01           |
| sum oct <sup>1</sup>           | 2.99          | 2.96          | 3.01           | 3.00             | 2.97           | 2.96           | 3.10           | 3.09           | 2.97            | 2.98              | 2.99           |
| sum A <sup>2</sup>             | 2.05          | 2.04          | 1.99           | 2.00             | 2.04           | 1.99           | 2.07           | 2.00           | 2.07            | 2.02              | 2.02           |

<sup>1</sup> sum oct = Al + Fe<sup>3+</sup> + Fe<sup>2+</sup> + Mg + Ti; <sup>2</sup>sum A = Ca + Mn + Ba + Sr + Y + REE + Th + Zr.

determined stability fields. Extrapolation of Carlson's isopleths suggests that strontianite compositions (Str<sub>56</sub>–Str<sub>67</sub>) may become stable with calcite at much lower temperatures and pressures. For example, strontianite with composition Str<sub>60</sub> would be in equilibrium with calcite at approximately 300 °C and 2 kbar. These data imply that carbonate compositions may record late recrystallization during exhumation of the marbles. However, Casey *et al.* (1996) predicted from drop-solution enthalpies that intermediate compositions are thermodynamically unstable and should unmix to form calcic strontianite and strontian aragonite at low temperatures. The miscibility gap should widen if calcite is formed instead of aragonite. Using a  $W = 13.5 \text{ kJ mol}^{-1}$ , the crest of a symmetrical miscibility gap (Str<sub>50</sub>Ar<sub>50</sub>) is predicted to be at about 540 °C. This crest lies very close to our Str<sub>56</sub> composition, and the predicted temperature agrees with our thermometric estimates of peak metamorphic conditions. More experiments are needed in the system CaCO<sub>3</sub>–SrCO<sub>3</sub> to confirm the presence of the solvus and to enable extrapolation of compositions of coexisting calcite and calcian strontianite over a wider range of  $P$ – $T$ .

### Clinopyroxene

Clinopyroxene in the marbles has similar compositions to those in adjacent eclogite samples. Liavtn marbles contain clinopyroxene with compositions of Jd<sub>29–31</sub>, whereas an adjacent sample from an eclogite layer has a composition of Jd<sub>29–32</sub> (Table 5, Fig. 6). Garnet from a Liavtn calc-silicate contains inclusions of omphacitic clinopyroxene (Jd<sub>22</sub>). Marbles from the Sjurholm locality have clinopyroxene that is more diopsidic with only a few mole percent jadeite (Jd<sub>6–14</sub>). Clinopyroxene from an adjacent eclogite sample has similar low jadeite contents (Jd<sub>14–17</sub>), whereas another has a higher jadeite content (Jd<sub>33–34</sub>) (Table 5). No K<sub>2</sub>O was detected in any of the clinopyroxene. Eclogites from the major eclogite facies shear zones have clinopyroxene compositions of Jd<sub>39–49</sub> and the protoliths are gabbroic anorthosites to gabbros (Austrheim & Griffin, 1985; Boundy *et al.*, 1992). The difference in jadeite content may well be a function of bulk composition rather than  $P$ – $T$  conditions.

**Table 8.** Titanite analyses.

| Wt% Oxide                      | HA5/90<br>1/lc | HA5/90<br>2/1 | BA9/91A'<br>2/1 | BA9/91A'<br>2/2r | BA25/92A'<br>1/lc | BA25/92 A'<br>1/2r | BA25/92B<br>1/lc | BA25/92B<br>1/2r | Liavtn<br>1/1 | Liavtn<br>2/1 |
|--------------------------------|----------------|---------------|-----------------|------------------|-------------------|--------------------|------------------|------------------|---------------|---------------|
| SiO <sub>2</sub>               | 30.61          | 30.41         | 30.26           | 29.62            | 30.10             | 29.96              | 30.53            | 30.09            | 30.51         | 30.67         |
| TiO <sub>2</sub>               | 34.65          | 34.15         | 35.02           | 35.06            | 35.18             | 35.28              | 35.88            | 35.81            | 35.78         | 36.24         |
| Al <sub>2</sub> O <sub>3</sub> | 2.52           | 2.70          | 2.47            | 2.29             | 0.94              | 0.76               | 1.43             | 1.38             | 2.32          | 1.67          |
| Fe <sub>2</sub> O <sub>3</sub> | 0.70           | 0.81          | 0.60            | 0.65             | 2.09              | 1.82               | 1.06             | 0.84             | 0.49          | 0.49          |
| MnO                            | 0.05           | 0.03          | 0.05            | 0.05             | 0.06              | 0.04               | 0.01             | 0.03             | 0.02          | 0.04          |
| MgO                            | 0.00           | 0.00          | 0.01            | 0.00             | 0.01              | 0.01               | 0.02             | 0.02             | 0.00          | 0.00          |
| CaO                            | 28.43          | 28.23         | 28.78           | 28.83            | 27.93             | 28.13              | 27.78            | 28.09            | 28.11         | 28.47         |
| BaO                            | 0.00           | 0.00          | 0.00            | 0.00             | 0.00              | 0.00               | 0.00             | 0.00             | 0.00          | 0.00          |
| SrO                            | 0.30           | 0.82          | 0.05            | 0.13             | 0.33              | 0.22               | 0.00             | 0.11             | 0.24          | 0.21          |
| Y <sub>2</sub> O <sub>3</sub>  | 0.48           | 0.00          | 0.03            | 0.04             | 0.05              | 0.06               | 0.29             | 0.18             | 0.00          | 0.00          |
| La <sub>2</sub> O <sub>3</sub> | 0.02           | 0.02          | 0.00            | 0.00             | 0.05              | 0.00               | 0.00             | 0.01             | 0.03          | 0.00          |
| Ce <sub>2</sub> O <sub>3</sub> | 0.00           | 0.00          | 0.11            | 0.09             | 0.22              | 0.31               | 0.16             | 0.08             | 0.12          | 0.21          |
| Pr <sub>2</sub> O <sub>3</sub> | 0.11           | 0.06          | 0.00            | 0.00             | 0.00              | 0.15               | 0.00             | 0.00             | 0.14          | 0.11          |
| Nd <sub>2</sub> O <sub>3</sub> | 0.06           | 0.06          | 0.00            | 0.00             | 0.21              | 0.18               | 0.15             | 0.29             | 0.17          | 0.19          |
| Sm <sub>2</sub> O <sub>3</sub> | 0.06           | 0.03          | 0.00            | 0.00             | 0.04              | 0.05               | 0.18             | 0.12             | 0.00          | 0.00          |
| Gd <sub>2</sub> O <sub>3</sub> | 0.00           | 0.00          | 0.27            | 0.03             | 0.08              | 0.13               | 0.12             | 0.00             | 0.05          | 0.18          |
| ThO <sub>2</sub>               | 0.00           | 0.01          | 0.10            | 0.07             | 0.00              | 0.00               | 0.08             | 0.06             | 0.00          | 0.00          |
| ZrO <sub>2</sub>               | 0.02           | 0.00          | 0.00            | 0.02             | 0.10              | 0.00               | 0.04             | 0.11             | 0.03          | 0.06          |
| F                              | 0.34           | 0.33          | 0.38            | 0.40             | 0.40              | 0.22               | 0.27             | 0.29             | 0.17          | 0.19          |
| Total                          | 98.36          | 97.74         | 98.14           | 97.29            | 97.79             | 97.33              | 98.01            | 97.52            | 98.17         | 98.73         |
| Calc. on basis of 3 cations    |                |               |                 |                  |                   |                    |                  |                  |               |               |
| Si                             | 1.00           | 0.99          | 0.98            | 0.97             | 0.99              | 1.00               | 1.01             | 0.99             | 1.00          | 1.00          |
| Ti                             | 0.85           | 0.84          | 0.86            | 0.86             | 0.87              | 0.88               | 0.89             | 0.89             | 0.88          | 0.89          |
| Al                             | 0.10           | 0.10          | 0.09            | 0.09             | 0.04              | 0.03               | 0.06             | 0.05             | 0.09          | 0.06          |
| Fe <sup>+3</sup>               | 0.02           | 0.02          | 0.01            | 0.02             | 0.05              | 0.05               | 0.03             | 0.02             | 0.01          | 0.01          |
| Mg                             | 0.00           | 0.00          | 0.00            | 0.00             | 0.00              | 0.00               | 0.00             | 0.00             | 0.00          | 0.00          |
| Mn                             | 0.00           | 0.00          | 0.00            | 0.00             | 0.00              | 0.00               | 0.00             | 0.00             | 0.00          | 0.00          |
| Ca                             | 0.99           | 0.99          | 1.00            | 1.01             | 0.99              | 1.00               | 0.98             | 0.99             | 0.99          | 1.00          |
| Ba                             | 0.00           | 0.00          | 0.00            | 0.00             | 0.00              | 0.00               | 0.00             | 0.00             | 0.00          | 0.00          |
| Sr                             | 0.01           | 0.02          | 0.00            | 0.00             | 0.01              | 0.00               | 0.00             | 0.00             | 0.00          | 0.00          |
| Y                              | 0.01           | 0.00          | 0.00            | 0.00             | 0.00              | 0.00               | 0.01             | 0.00             | 0.00          | 0.00          |
| REE                            | 0.00           | 0.00          | 0.00            | 0.00             | 0.01              | 0.01               | 0.01             | 0.01             | 0.01          | 0.01          |
| Th                             | 0.00           | 0.00          | 0.00            | 0.00             | 0.00              | 0.00               | 0.00             | 0.00             | 0.00          | 0.00          |
| Zr                             | 0.00           | 0.00          | 0.00            | 0.00             | 0.00              | 0.00               | 0.00             | 0.00             | 0.00          | 0.00          |
| F                              | 0.03           | 0.03          | 0.04            | 0.04             | 0.04              | 0.02               | 0.03             | 0.03             | 0.02          | 0.02          |
| sum A                          | 1.01           | 1.01          | 1.01            | 1.02             | 1.00              | 1.02               | 1.00             | 1.01             | 1.00          | 1.01          |
| sum oct                        | 0.96           | 0.96          | 0.97            | 0.97             | 0.96              | 0.96               | 0.97             | 0.97             | 0.98          | 0.97          |

**Table 9.** Celestine and barite analyses.

|                                      | HA4/90<br>cel 1/1 | BA9/91 A<br>cel 1/6 | BA9/91 A<br>cel 2/3 | BA9/91 A<br>int 3/7 | BA9/91 A'<br>cel 2/2 | BA9/91 A'<br>bar 2/2 | BA25/92 A<br>cel 1/1a | BA25/92 A<br>bar 4/1 | BA25/92B<br>cel 5/2 | BA25/92B<br>bar 3/1 |
|--------------------------------------|-------------------|---------------------|---------------------|---------------------|----------------------|----------------------|-----------------------|----------------------|---------------------|---------------------|
| SO <sub>3</sub>                      | 43.78             | 40.76               | 43.74               | 40.36               | 44.53                | 33.72                | 41.15                 | 33.43                | 40.24               | 33.04               |
| TiO <sub>2</sub>                     | 0.00              | 0.00                | 0.00                | 0.00                | 0.06                 | 0.01                 | 0.00                  | 0.00                 | 0.00                | 0.00                |
| SrO                                  | 48.55             | 57.16               | 51.59               | 23.65               | 55.59                | 2.98                 | 35.94                 | 4.82                 | 35.49               | 2.21                |
| PbO                                  | 0.00              | 0.00                | 0.00                | 0.00                | 0.00                 | 0.00                 | 0.00                  | 0.00                 | 0.00                | 0.00                |
| CaO                                  | 3.19              | 0.46                | 0.09                | 0.32                | 0.22                 | 0.10                 | 3.25                  | 0.17                 | 2.20                | 0.92                |
| BaO                                  | 5.30              | 0.35                | 3.03                | 34.36               | 1.52                 | 62.01                | 18.55                 | 59.09                | 21.54               | 64.38               |
| FeO                                  | 0.04              | 0.06                | 0.04                | 0.01                | 0.03                 | 0.39                 | 0.14                  | 0.00                 | 0.07                | 0.01                |
| MnO                                  | 0.03              | 0.00                | 0.00                | 0.00                | 0.00                 | 0.00                 | 0.00                  | 0.00                 | 0.00                | 0.00                |
| Total                                | 100.88            | 98.80               | 98.49               | 98.70               | 101.95               | 99.21                | 99.03                 | 97.52                | 99.55               | 100.56              |
| Calculated on the basis of 2 cations |                   |                     |                     |                     |                      |                      |                       |                      |                     |                     |
| S                                    | 0.99              | 0.98                | 1.01                | 1.02                | 1.00                 | 0.98                 | 0.99                  | 0.98                 | 0.98                | 0.95                |
| Ti                                   | 0.00              | 0.00                | 0.00                | 0.00                | 0.00                 | 0.00                 | 0.00                  | 0.00                 | 0.00                | 0.00                |
| Sr                                   | 0.85              | 1.06                | 0.92                | 0.46                | 0.97                 | 0.07                 | 0.67                  | 0.11                 | 0.67                | 0.05                |
| Pb                                   | 0.00              | 0.00                | 0.00                | 0.00                | 0.00                 | 0.00                 | 0.00                  | 0.00                 | 0.00                | 0.00                |
| Ca                                   | 0.10              | 0.03                | 0.00                | 0.01                | 0.01                 | 0.00                 | 0.11                  | 0.01                 | 0.08                | 0.04                |
| Ba                                   | 0.06              | 0.01                | 0.04                | 0.45                | 0.02                 | 0.94                 | 0.23                  | 0.90                 | 0.27                | 0.96                |
| Fe                                   | 0.00              | 0.00                | 0.00                | 0.00                | 0.00                 | 0.01                 | 0.00                  | 0.00                 | 0.00                | 0.00                |
| Mn                                   | 0.00              | 0.00                | 0.00                | 0.00                | 0.00                 | 0.00                 | 0.00                  | 0.00                 | 0.00                | 0.00                |
| O                                    | 3.97              | 4.02                | 3.95                | 3.91                | 4.01                 | 3.96                 | 3.97                  | 3.96                 | 3.96                | 3.90                |
| Celestine                            | 83.7              | 96.4                | 95.8                | 50.0                | 97.5                 | 6.6                  | 66.0                  | 10.7                 | 65.6                | 4.7                 |
| Barite                               | 6.2               | 0.9                 | 4.2                 | 48.9                | 1.8                  | 93.0                 | 23.0                  | 88.6                 | 26.9                | 91.7                |
| Anhydrite                            | 10.2              | 2.7                 | 0.0                 | 1.0                 | 0.7                  | 0.4                  | 11.0                  | 0.7                  | 7.5                 | 3.6                 |

cel = celestine, bar = barite, int = intermediate.

**Table 10.** Biotite analysis.

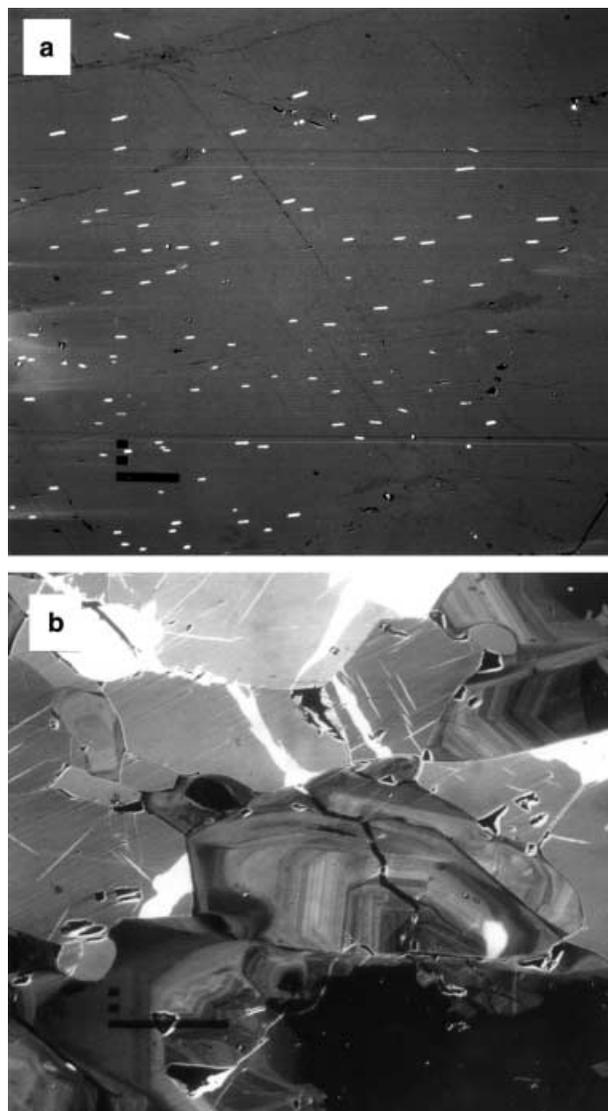
| Wt% Oxide                      | BA 25/92a.1 | BA 25/92a.2 |
|--------------------------------|-------------|-------------|
| Al <sub>2</sub> O <sub>3</sub> | 13.12       | 13.27       |
| FeO                            | 23.47       | 23.21       |
| MnO                            | 0.14        | 0.21        |
| MgO                            | 9.95        | 10.22       |
| CaO                            | 0.26        | 0.31        |
| BaO                            | 0.38        | 0.39        |
| Na <sub>2</sub> O              | 0.14        | 0.09        |
| K <sub>2</sub> O               | 8.82        | 8.31        |
| F                              | 0.01        | 0.01        |
| Cl                             | 0.07        | 0.07        |
| Total                          | 94.70       | 94.83       |
| Calc. on basis of 24 oxygen    |             |             |
| Si                             | 6.25        | 6.26        |
| Ti                             | 0.26        | 0.27        |
| Al                             | 2.66        | 2.67        |
| Fe <sup>2+</sup>               | 3.37        | 3.31        |
| Mn                             | 0.02        | 0.03        |
| Mg                             | 2.55        | 2.60        |
| Ca                             | 0.05        | 0.06        |
| Ba                             | 0.03        | 0.03        |
| Na                             | 0.05        | 0.03        |
| K                              | 1.93        | 1.81        |
| F                              | 0.01        | 0.01        |
| Cl                             | 0.02        | 0.02        |
| X <sub>Phl</sub>               | 0.43        | 0.44        |
| X <sub>Ann</sub>               | 0.57        | 0.56        |

### Garnet

Garnet was found in only one of the carbonate samples, Liavatn sample BA21/93. The compositions vary from Alm<sub>52–56</sub>Grs<sub>28–33</sub>Pyp<sub>11–16</sub>Sps<sub>2</sub> (Table 6) and show weak prograde zonation. Garnet from one of the adjacent eclogite layers displays similar compositions of Alm<sub>51–56</sub>Grs<sub>31–35</sub>Pyp<sub>10–14</sub>Sps<sub>2</sub> excluding the core of one large garnet analysed, which may be relict from earlier granulite facies metamorphism (sample Liavatn; Table 6). The eclogite sample from Sjurholmen (BA69/93) has similar garnet compositions of Alm<sub>58–60</sub>Grs<sub>26–28</sub>Pyp<sub>8</sub>Sps<sub>6</sub> (Fig. 7). In contrast, eclogite facies garnet from the major shear zones is typically less almandine rich and much more pyropic (Alm<sub>40</sub>Grs<sub>21</sub>Pyp<sub>38</sub>Sps<sub>1</sub>) for the anorthosite protolith and Alm<sub>47</sub>Grs<sub>26</sub>Pyp<sub>26</sub>Sps<sub>1</sub> for the gabbro protolith; Boundy *et al.*, 1992).

### Epidote/Allanite

The Liavatn marbles contain REE-bearing epidote, whereas in the Sjurholmen marbles it is more properly termed allanite (Table 7). In the investigated marbles, the epidote and allanite are dramatically zoned with respect to REEs, for example varying from <1–5 wt percentage total REE oxides in a single grain. No consistent zonation pattern was discovered. Within an individual sample, some grains are zoned with REE-rich cores to REE-poor rims while other display the opposite zonation (Table 7; Fig. 4b). In other grains the zonation in REEs does not vary consistently or smoothly from core to rim. The Sr contents also vary, from 1.5 to 4.0 wt % SrO in epidote from the Liavatn marbles and from 0.3 to 1.1 wt.% SrO in allanite from the

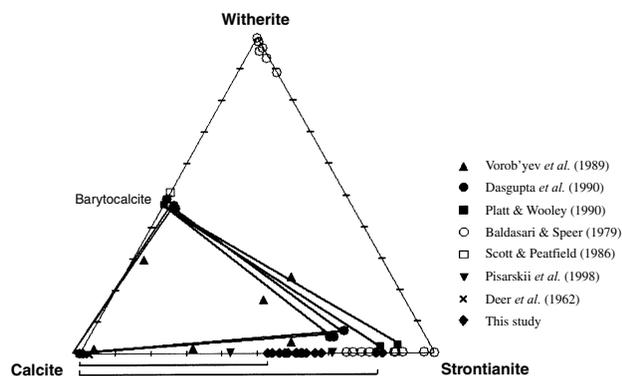


**Fig. 4.** Back-scattered electron microscope (BSE) images of samples used in the study. (a) Calcite (grey) with aligned inclusions of celestine (bright elongate needles) (from sample BA25/92 A, Sjurholmen locality). The two lighter gray square patches in the calcite are the result of beam damage during microprobe analysis. The scale bar is 100  $\mu\text{m}$  (b) Zoning in allanite from sample BA69/93, Sjurholmen locality. Lighter areas correspond to more REE enrichment, whereas the darkest areas of the allanite are the least REE enriched. The zoning shows several sequences of REE enrichment during growth from core to rim. The scale bar is 100  $\mu\text{m}$ .

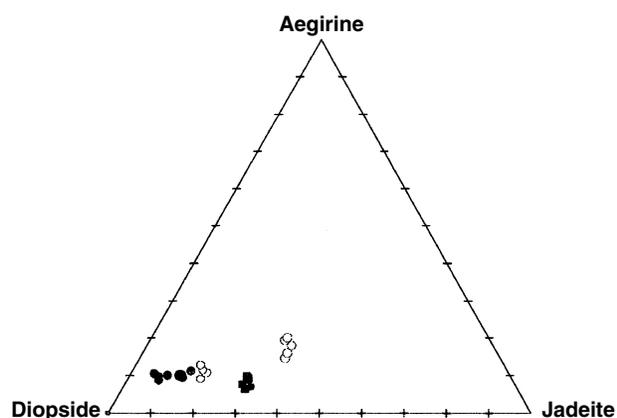
Sjurholmen marbles. Zoisite crystals from the eclogite shear zones have similar Sr contents and have up to 2 wt.% REE oxides (Boundy *et al.*, 1992). One of the eclogite samples from the Sjurholmen locality (BA69/93) has allanite similar to the adjacent marbles.

### Titanite

The titanite analysed from the marbles is typically unzoned (Table 8). It contains variable amounts of



**Fig. 5.**  $\text{CaCO}_3$  –  $\text{BaCO}_3$  –  $\text{SrCO}_3$  ternary plot of orthorhombic carbonates from sample BA9/91 A demonstrating the extent of solid solution along the  $\text{CaCO}_3$  –  $\text{SrCO}_3$  binary. Tie lines have been drawn between coexisting compositions. Filled circles represent carbonates from the Sausar group, India, metamorphosed at 660 °C and 6 kbar (Dasgupta *et al.*, 1990). Filled, inverted triangles represent travertine deposit analyses from Pisarskii *et al.* (1998).

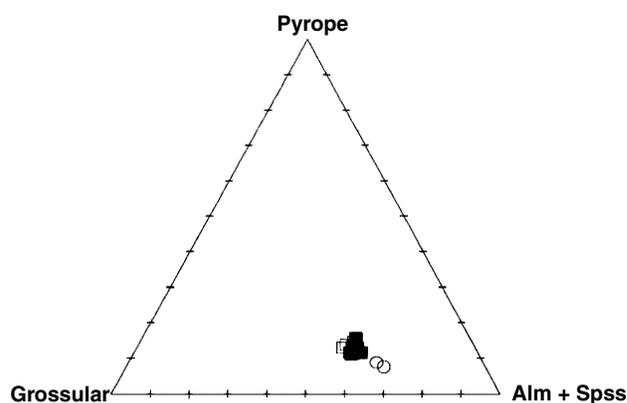


**Fig. 6.** Clinopyroxene compositions. Marble samples are shown as solid symbols and eclogite samples are shown as open symbols. The marble and eclogite samples from the Liavtn locality (■, □) are more jadeite rich than most of the Sjurdholmen locality (●, ○) samples.

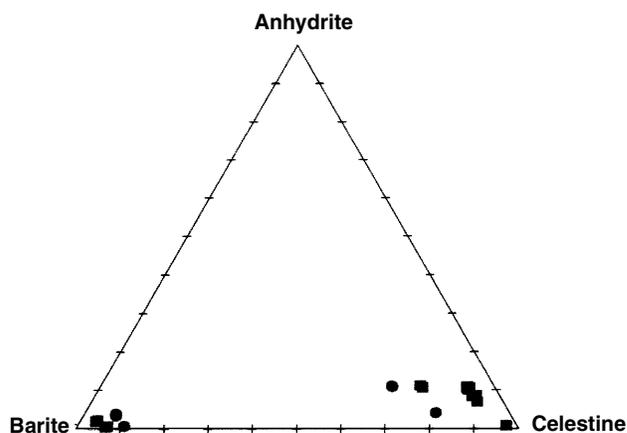
$\text{Al}_2\text{O}_3$  (1–3 wt.%), with low contents of F (0.3 wt.%) and REE (<1 wt.%).

#### Celestine and barite

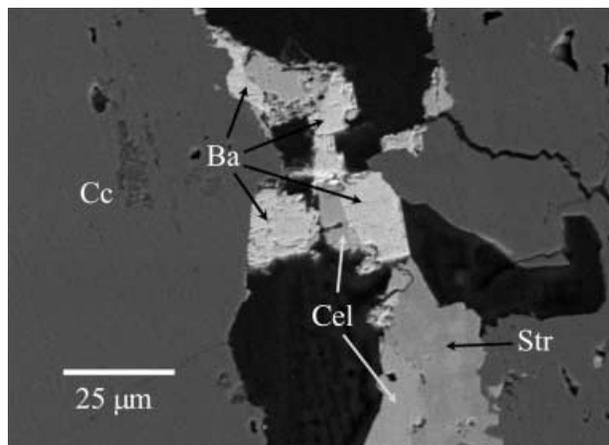
The marble assemblages include celestine and a few have coexisting celestine and barite (Table 9). These are plotted in Fig. 8 and show two populations of barite and barite-celestine with solid solution of minor  $\text{CaSO}_4$ . One sample (BA9/91 A) contains numerous clusters of barite ( $\text{Ba}_{0.49}\text{Ca}_{0.01}\text{Sr}_{0.50}\text{SO}_4$ ) and celestine ( $\text{Sr}_{0.96}\text{Ba}_{0.04}\text{SO}_4$ ) in contact with one another (Fig. 9). A sharp grain boundary between the two phases is suggestive of a possible miscibility gap. Prieto *et al.* (1997) concluded that intermediate compositions of  $(\text{Ba},\text{Sr})\text{SO}_4$  are rare because of the extremely low



**Fig. 7.** Garnet compositions plotted as pyrope – grossular – almandine + spessartine. Marble samples are shown as solid symbols and eclogite samples are shown as open symbols. Liavtn locality (■, □) and Sjurdholmen locality (●, ○) samples.



**Fig. 8.** Sulphate mineral compositions plotted as barite-celestine-anhydrite. Liavtn locality (■, □) and Sjurdholmen locality (●, ○) samples.



**Fig. 9.** BSE image of celestine (Cel) and barite (Ba), and strontianite (Str) in a matrix of calcite (Cc) from marble sample BA9/91 A. The fairly sharp boundary is consistent with the presence of a miscibility gap between celestine and barite.

solubility of BaSO<sub>4</sub> relative to SrSO<sub>4</sub>, which leads to bimodal nucleation of these phases when precipitated from a fluid and that intermediate compositions are therefore only seen if the reacting composition is within a very narrow range. During metamorphic crystallization one would, however, expect to find a single composition, not a wide range in compositions for a single rock in the absence of a miscibility gap. In this study, the analytical data are rather scattered (Fig. 8). Whether a miscibility gap could extend to eclogitic *P–T* conditions, or whether some of the sulphate could have been a high pressure polymorph (Liu & Bassett, 1986), is difficult to evaluate without direct experiments.

### Biotite

Biotite, analysed for one marble sample, BA25/92 A (Table 10), shows typical metamorphic compositions, with moderate Ti (*c.* 2 wt%),  $X_{Ann} = 0.56$  and  $X_{Phl} = 0.44$ . Biotite has very low Cl and F, typically  $0.07 \pm 0.01$  wt% Cl and  $0.01–0.08 \pm 0.05$  wt% F. This corresponds to only small amounts of halogens in the hydroxyl site (0.01 mol percentage).

### WHOLE ROCK CHEMISTRY

The whole rock chemistry has been determined for two marble samples from Liavatn (Table 11). Given the high Sr content of the calcite and the presence of other phases, such as celestine, it is not surprising that the marbles have high bulk-rock Sr contents (9,500–11 000 p.p.m.; Table 11). The high Sr and Y contents are suggestive of a possible magmatic origin (Woolley & Kempe, 1989), although Sr enrichment is also possible in some sedimentary environments as well. However, low concentrations of some other key trace elements, for example Ba (110–160 p.p.m.) and Nb (<5 p.p.m.), suggests that the calc-silicate marble is not likely of magmatic origin (a former carbonatite). Since there is no independent evidence for a protolith from the field data, it is difficult to assess the extent of possible metasomatism during the cycling through the orogenic events. The surrounding granulite and eclogites of anorthositic to gabbroic compositions yield Sr contents of *c.* 1000 p.p.m., Ba contents of *c.* 500 p.p.m. (Rockow *et al.*, 1997). An eclogite from Sjurholmen (mangerite protolith?) reveals Sr contents of *c.* 698 p.p.m and unusually high Ce (3078 p.p.m.), Zr (4122 p.p.m.) Y (502 p.p.m.) and Nb (108 p.p.m.) contents (Svensen *et al.*, 2001).

### Sr isotopes

The <sup>87</sup>Sr/<sup>86</sup>Sr ratios of four of the marble samples were determined at the University of Muenster. Samples from the Liavatn area have variable <sup>87</sup>Sr/<sup>86</sup>Sr ratios, ranging from 0.7051 to 0.7059 (Table 12). The sample analysed from the Sjurholmen locality yielded a <sup>87</sup>Sr/<sup>86</sup>Sr ratio of 0.7058. These <sup>87</sup>Sr/<sup>86</sup>Sr values are

**Table 11.** XRF whole rock analyses.

| Sample no:                     | Liavatn<br>HA4/90 A | Liavatn<br>BA9/91 A |
|--------------------------------|---------------------|---------------------|
| Major elements (wt.%)          |                     |                     |
| SiO <sub>2</sub>               | 0.95                | 0.85                |
| TiO <sub>2</sub>               | 0.04                | 0.02                |
| Al <sub>2</sub> O <sub>3</sub> | 0.69                | 0.69                |
| Fe <sub>2</sub> O <sub>3</sub> | 1.91                | 2.22                |
| MnO                            | 0.20                | 0.19                |
| MgO                            | 1.24                | 1.58                |
| CaO                            | 51.21               | 51.12               |
| Na <sub>2</sub> O              | 0.00                | 0.00                |
| K <sub>2</sub> O               | 0.00                | 0.00                |
| P <sub>2</sub> O <sub>5</sub>  | 1.71                | 0.10                |
| H <sub>2</sub> O               | 1.23                | 0.65                |
| CO <sub>2</sub>                | 38.84               | 41.20               |
| Cr <sub>2</sub> O <sub>3</sub> | 0.00                | 0.00                |
| NiO                            | 0.00                | 0.00                |
| Total                          | 98.02               | 98.62               |
| Trace elements (p.p.m.)        |                     |                     |
| Ba                             | 161                 | 112                 |
| Rb                             | 13                  | 11                  |
| Sr                             | 11082               | 9533                |
| Pb                             | 25                  | 22                  |
| Th                             | 13                  | 12                  |
| U                              | 42                  | 35                  |
| Nb                             | <5                  | <5                  |
| La                             | 63                  | 46                  |
| Ce                             | 138                 | 101                 |
| Nd                             | 87                  | 70                  |
| Y                              | 130                 | 115                 |
| Zr                             | 33                  | 28                  |
| V                              | <3                  | <3                  |
| Cr                             | <3                  | <3                  |
| Ni                             | <2                  | <2                  |
| Co                             | 14                  | 11                  |
| Cu                             | 24                  | 20                  |
| Zn                             | <3                  | <3                  |
| Ga                             | <1                  | <1                  |
| Hf                             | <2                  | <2                  |
| S                              | 936                 | 696                 |

**Table 12.** Isotopic analyses.

| Sample    | Locality | Locality   | <sup>87</sup> Sr/ <sup>86</sup> Sr | Error |
|-----------|----------|------------|------------------------------------|-------|
| HA4 90/4a | marble   | Liavatn    | 0.705256                           | 9     |
| HA4 90/4b | marble   | Liavatn    | 0.705134                           | 10    |
| HA5 90    | marble   | Liavatn    | 0.705860                           | 10    |
| BA9/91    | marble   | Liavatn    | 0.705079                           | 10    |
| BA25 92–4 | marble   | Sjurholmen | 0.705816                           | 8     |

more characteristic of crustal than mantle materials. The variability of ratios, even between samples from the same general locality, likely indicates that they have been modified from their original values or were heterogeneous from the beginning. Metamorphism, the influx of fluids, and metasomatism could variably alter the <sup>87</sup>Sr/<sup>86</sup>Sr ratios. The marbles may well have experienced more than one phase of metamorphism and retrogression (Boundy, 1995; van Wyck *et al.*, 1996). Unfortunately, the nature and extent of metasomatism in altering the isotopic ratios is not known. McCaig *et al.* (2000) documented infiltration of fluids rich in Sr can significantly effect the <sup>87</sup>Sr/<sup>86</sup>Sr composition of carbonate rocks. Based on the field relations, whole rock chemistry and isotopic ratios, the precursor to the

marbles remains unresolved at this time although the Sr isotopes are strongly suggestive of a crustal source (or modified crustal material).

### *P*–*T*–*X* CONDITIONS OF MARBLE METAMORPHISM

#### Thermobarometry

Eclogites from Holsnøy Island (Lindås Nappe) are estimated to have formed at about 650–700 °C at pressures of 15–21 kbar (Austrheim & Griffin, 1985; Jamtveit *et al.*, 1990; Boundy *et al.*, 1992). However, heretofore the equilibrium *P*–*T* conditions of the marbles have not been evaluated. The calcite–dolomite thermometer of Anovitz & Essene (1987) was used to place lower boundary constraints on the metamorphic temperature conditions experienced by the marbles. Since the Mg content of the calcite is commonly reduced by retrograde exsolution of dolomite and/or recrystallization, temperature estimates based on the calcite–dolomite solvus often represent a minimum temperature of formation (Essene, 1983). Furthermore, in the absence of dolomite, the calcite compositions can only be used to estimate minimum temperatures. Calcite and dolomite were found together in only one sample in this study (BA9/91 A, Liavatn locality). However, the texture is suggestive of exsolution, and therefore the coexistence of the two phases under equilibrium conditions cannot be assumed. Temperature estimates using the calcite–dolomite solvus are consistent within a given sample. No discernible intra or intercrystalline variation within a given sample was found (Table 2). For the Liavatn marble, the temperature estimates range from 460 to 490 °C, with one sample recording about 390 °C (Table 1). One marble from the Sjurholmen locality yielded estimates of about 420 °C.

Temperature estimates were obtained from intercalated calc-silicate rocks and adjacent eclogite samples using the garnet–clinopyroxene thermometer of Krogh Ravna (2000) and assuming pressures of 20 kbar. The temperatures calculated using Krogh Ravna (2000) are typically lower than those estimated using the thermometer of Krogh (1988; Table 13). For the eclogite sample from Liavatn, garnet and clinopyroxene yield temperatures of about 640–690 °C, the calc-silicate garnet with included omphacite from Liavatn 580–650 °C, and the eclogite sample from the Sjurholmen locality about 595 °C (Table 13). Previous temperature estimates for the Holsnøy eclogites range from 650 to 700 °C (Austrheim & Griffin, 1985; Jamtveit *et al.*, 1990; Boundy *et al.*, 1992). The lower temperature estimates from marbles and calc-silicates and the Sjurholmen eclogite may indicate re-equilibration at lower temperatures during retrogression.

The presence of omphacite with jadeite contents of up to 30% implies that marble formed under high

**Table 13.** Garnet – pyroxene thermometry of eclogites and calc-silicates.

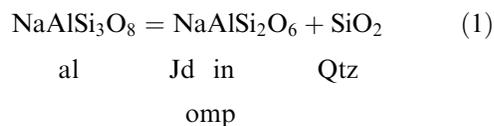
| Sample   | grt  | cpx  | $K_D$ | $T^\circ\text{C}$ (K88) | $T^\circ\text{C}$ (K00) | (K00-K88)<br>$\Delta T$ |
|--|------|------|-------|-------------------------|-------------------------|-------------------------|
| Liavatn eclogite sample                                    |      |      |       |                         |                         |                         |
| Liavatn  | 1/3r | 2/2r | 10.9  | 671                     | 644                     | –27                     |
| Liavatn  | 1/4r | 1/2r | 10.5  | 687                     | 663                     | –24                     |
| Liavatn  | 2/2r | 4/2r | 11.6  | 669                     | 640                     | –29                     |
| Liavatn  | 3/2r | 4/3r | 9.3   | 721                     | 693                     | –28                     |
| Liavatn  | 4/3r | 3/2r | 11.9  | 671                     | 639                     | –32                     |
| Averages:  |      |      | 10.8  | 684                     | 656                     | –28                     |
| Sjurholmen eclogite sample                                 |      |      |       |                         |                         |                         |
| BA69/93  | 1/1c | 2/1c | 13.3  | 586                     | 597                     | 11                      |
| BA69/93  | 1/2r | 2/4r | 14.3  | 583                     | 592                     | 8                       |
|  |      |      | 13.8  | 585                     | 594                     | 10                      |
| Liavatn calc-silicate sample – garnets with cpx inclusions |      |      |       |                         |                         |                         |
| BA21/93 c  | 9    | 3    | 13.1  | 628                     | 586                     | –42                     |
| BA21/93 c  | 2    | 9    | 12.2  | 629                     | 592                     | –37                     |
| BA21/93 c  | 5    | 12   | 13.8  | 621                     | 577                     | –44                     |
| BA21/93 c  | 4    | 7    | 9.5   | 688                     | 650                     | –38                     |
| Averages:  |      |      | 12.2  | 642                     | 602                     | –40                     |

K88 = Krogh (1988) @20 kbar.

K00 = Krogh Ravna (2000) @20 kbar.

r = rim; c = core.

pressure conditions. However, in the absence of both quartz and albite, the jadeite content is not useful in constraining peak pressures (Holland, 1980). The adjacent eclogite sample from the Liavatn locality also has omphacite with Jd<sub>30</sub>, and the assemblage includes coexisting quartz. Therefore, using the reaction

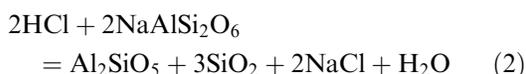


at the estimated average temperature of 660 °C yields a minimum pressure of about 13 kbar in the absence of albite (Holland, 1980). The field association with eclogites and the jadeite content of clinopyroxene from the marbles strongly suggests that the marbles experienced eclogite facies conditions.

#### Fluid composition

Fluid compositions associated with high-pressure metamorphism of the marbles from Holsnøy Island are of interest in order to compare with those from the eclogites. The activity of fluid components in a mixed CO<sub>2</sub>–H<sub>2</sub>O fluid can vary dramatically in high-pressure rocks. Austrheim (1990) illustrated variable activities of components of infiltrating fluids during the eclogitization of granulite facies ultramafites from Holsnøy Island. Svensen *et al.* (2001) showed a mafic eclogite from Sjurholmen may have equilibrated with H<sub>2</sub>O-rich fluids containing high levels of halogens. At elevated pressures, dissociated salts can dramatically reduce *a*H<sub>2</sub>O and increase SiO<sub>2</sub> solubility (Shmulovitch & Graham, 1999). Thus, further evaluation of the fluids associated with the marbles and the surrounding eclogites is critical in evaluating the phase equilibria.

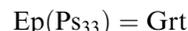
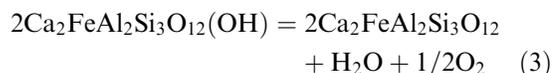
The halogen component of the fluid has been investigated by undertaking Cl and F analysis of biotite on the electron microprobe, giving averages of  $0.07 \pm 0.01$  wt percentage for Cl and  $0.01 \pm 0.05$  wt percentage for F Table (10). These values are less than a third of those reported by Svensen *et al.* (2001) for Cl and one tenth of those for F in biotite with the highest halogen content from WGR eclogite samples. However, halogen contents of hydrous silicates in eclogites from Sjurholmen area are below the EMP detection limit (Svensen *et al.*, 2001). Therefore the increased halogens in the Sjurholmen area were only evidenced locally in the evolved fluid inclusions. This suggests higher  $a_{\text{H}_2\text{O}}$  since Cl prefers to remain in solution and not enter into hydrous minerals under high  $a_{\text{H}_2\text{O}}$ . Using the experiments of Munoz & Swenson (1981) and the recalibration of Munoz (1992), the fluid in equilibrium with the biotite is estimated to have  $f_{\text{HCl}} < 0.021$ . This value is an upper limit assuming  $f_{\text{H}_2\text{O}} = \text{unity}$ , but nonetheless gives a *maximum*  $f_{\text{HCl}}$ . The log K of the following reaction yields an upper limit for the activity of NaCl:



This reaction gives a maximum activity of NaCl = 0.019 at 600 °C and 20 kbar in the absence of kyanite and quartz and using the above HCl, H<sub>2</sub>O, and jadeite components. The mole fraction of NaCl has been calculated using the  $a/X$  relations experimentally determined by Aranovich & Newton (1996) and yields values of  $X_{\text{NaCl}} = 0.023$ . This halogen component in the fluid is much lower than that of the brines suggested by Svensen *et al.* (2001) in other Norwegian eclogites.

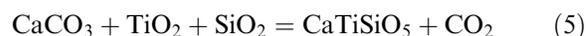
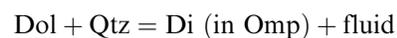
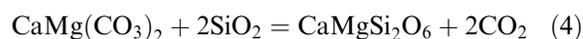
The effect of dissolved SiO<sub>2</sub> on  $a_{\text{H}_2\text{O}}$  in the peak metamorphic fluid may be estimated from the experiments of Manning (1994), who found 2.87 wt.% SiO<sub>2</sub> ( $X_{\text{SiO}_2} = 0.009$ ) dissolved in H<sub>2</sub>O at 600 °C and 20 kbar in the presence of quartz. Using Raoult's law ( $a_{\text{H}_2\text{O}} = X_{\text{H}_2\text{O}}$ ) for small amounts of dilution leads to a reduction in the activity of H<sub>2</sub>O by a factor of only 0.991. This provides a maximum estimate of dilution by SiO<sub>2</sub> for the Holsnøy marbles, which lack quartz. In the absence of an excess fluid phase, the  $T$ - $X$  equilibria generally move to lower temperature with decreasing fluid pressures but will not strongly affect the inferences of a fictive fluid composition. The presence of hydrous minerals such as biotite in the marble severely limits any reduction in fluid pressure (Edwards & Essene, 1988). The use of  $T/(X_{\text{H}_2\text{O}} + X_{\text{CO}_2})$  diagrams is an excellent approximation for these rocks, which have low salinities as recorded by their Cl contents and only minor solution of SiO<sub>2</sub> in the fluid phase.

The  $f_{\text{O}_2}$  was calculated for the garnet-bearing calc-silicate sample (BA21/93) using the method of Donohue & Essene (2000), which utilizes coexisting garnet and epidote via the reaction:

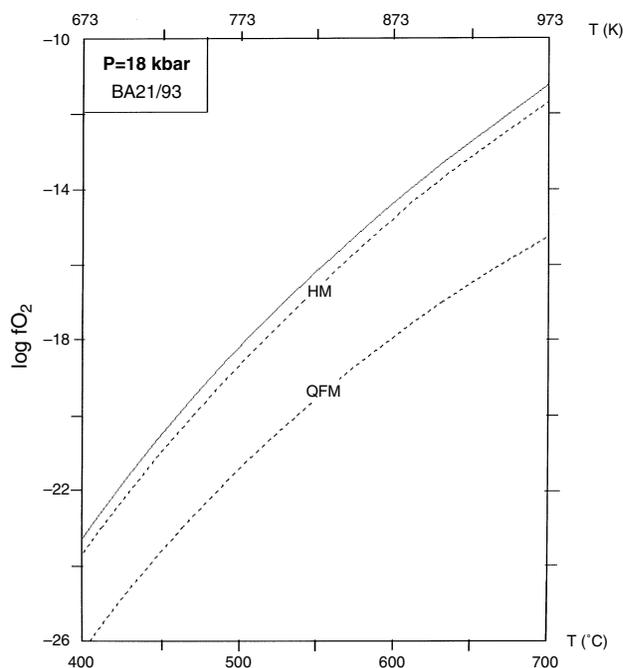


The results indicate fairly oxidizing conditions during metamorphism, approximately 0.3–0.5 log  $f_{\text{O}_2}$  units above the hematite–magnetite buffer over the calculated range of 400–700 °C at a fixed pressure of 18 kbar (Fig. 10). These conditions would not permit a significant CH<sub>4</sub> component in the fluid.

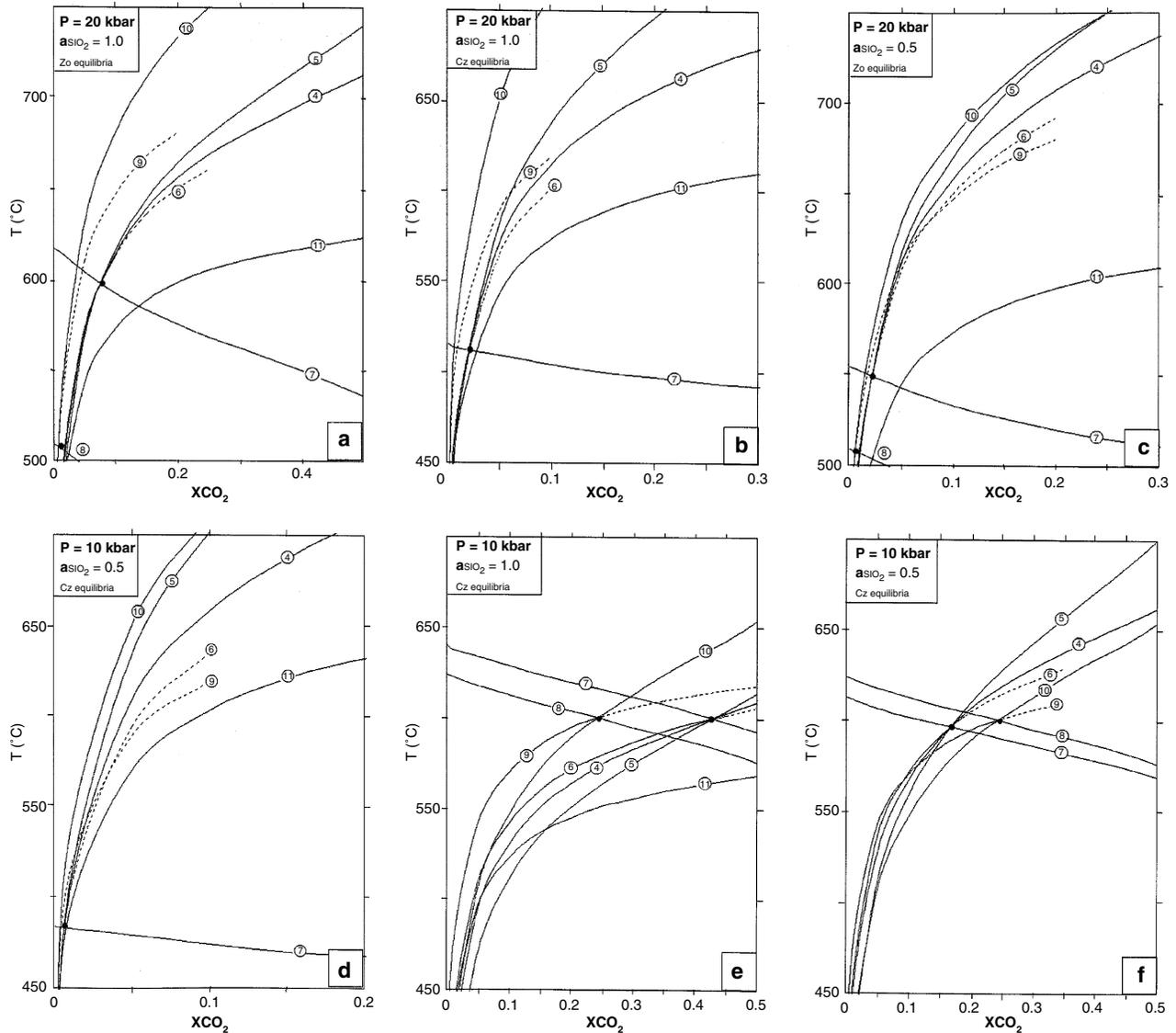
We suggest that an H<sub>2</sub>O–CO<sub>2</sub> binary is a good approximation for the composition of the fluid for the Holsnøy marbles. For samples with coexisting quartz, upper limits on  $X_{\text{CO}_2}$  can be calculated at the estimated  $P$ - $T$  using reactions like:



Using reaction (4) for similar composition clinopyroxene from eclogite facies marbles from the Dabie Mountains in Central China, Wang & Liou, 1993) calculated very low  $X_{\text{CO}_2}$  values (<0.03). The Dabie marbles and eclogites are estimated to have formed



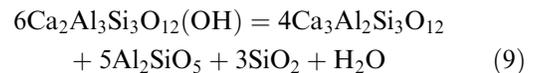
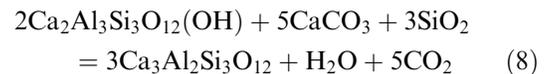
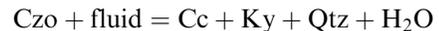
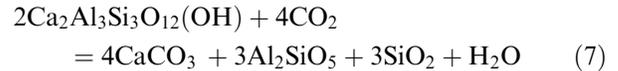
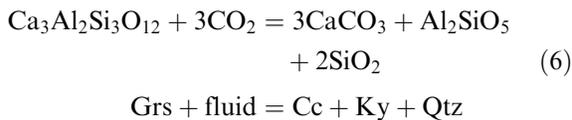
**Fig. 10.** Log  $f_{\text{O}_2}$  -  $T$  diagram calculated for equation 14 and garnet - epidote compositions from sample BA21/93 (solid line). Shown as references are the solid oxygen buffer curves (dashed lines) for hematite - magnetite (HM) and quartz - fayalite - magnetite (QFM). Pressure was fixed at 18 kbar for this calculation. An uncertainty of  $\pm 5$  kbar generates  $\pm 1.0$  log units in  $f_{\text{O}_2}$  (Donohue & Essene, 2000).

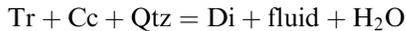
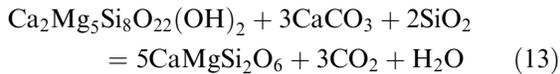
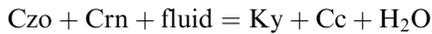
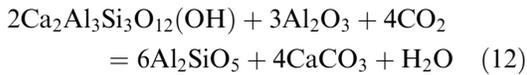
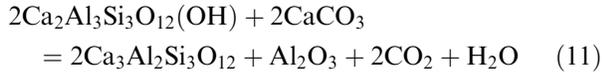
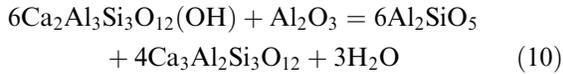


**Fig. 11.** (a–f)  $T$ - $X_{\text{CO}_2}$  diagrams for calc-silicate sample BA21/93. Solid lines represent stable reactions while dashed lines show metastable extensions. Diagrams 11a–d were calculated at 20 kbar using zoisite (Zo) thermodynamic data for 11a & 11c and clinozoisite (Cz) data for 11b & 11d. Figure 11(c,d) shows the effects of reduced silica activity on the locus of reactions in each system. Figure 11(e,f) shows the effects of reduced pressure on the equilibria of Fig. 11(b,d). The limiting reactions for zoisite (Zo), clinozoisite (Cz), and grossular (Grs) are labelled with the appropriate abbreviations.

at similar temperatures to the Holsnøy eclogites, although at higher *minimum* pressures (>25 kbar) based on the occurrence of coesite.

$T$ - $X_{\text{CO}_2}$  equilibria have been constructed for one calc-silicate sample, BA21/93 (Fig. 11a–f) using the following dehydration and decarbonation reactions corrected for solid solutions:



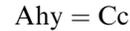


All  $T$ - $X_{\text{CO}_2}$  equilibria were calculated using THERMOCALC (Holland & Powell, 1998) and the equation of state for the fluid was taken directly from this data set. Reactions involving epidote were calculated for both clinozoisite and zoisite because the crystal structure present in the sample is unknown and the thermodynamic properties for the two species are significantly different (e.g. Robie & Hemingway, 1995; Holland & Powell, 1998). Figure 11(a,b) shows the  $T$ - $X_{\text{CO}_2}$  relations for sample BA21/93 at 20 kbar and an activity of silica ( $a_{\text{SiO}_2}$ ) = 1. Reactions 6, 7, 8 and 9 form an invariant point involving the phases (clino)zoisite–grossular–calcite–kyanite–quartz– $\text{H}_2\text{O}$ – $\text{CO}_2$ . Reactions 10, 11, and 12 (the corundum analogues of reactions 7, 8, 9) also form an invariant point at lower temperature and lower  $X_{\text{CO}_2}$ . Reactions 7 and 9 limit the epidote stability in the absence of kyanite. Reaction 6 is the lower stability for grossular. The reaction space for BA21/93 is therefore limited to the region below reaction 9 and left of reaction 8, although it is possible that the assemblage is located directly on the invariant point. This implies that at 20 kbar and  $a_{\text{SiO}_2}$  = 1, the  $X_{\text{CO}_2}$  must be <0.08 for the zoisite equilibria or <0.025 for the clinozoisite equilibria regardless of the temperature estimate. If the  $a_{\text{SiO}_2}$  is reduced to 0.5 from 1.0 (Fig. 11c,d), the quartz-present invariant point for both the zoisite and clinozoisite reactions shifts to even more  $\text{H}_2\text{O}$  rich conditions. Since quartz has not been found in sample BA21/93, it is possible that the  $a_{\text{SiO}_2}$  is indeed <1.0.

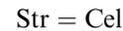
Lowering the pressure at which these calculations are performed has the effect of driving the loci of the invariant points towards much more  $\text{CO}_2$ -rich conditions (Fig. 11e,f). While the pressure estimates of these rocks (18–21 kbar) seems reasonable based on the data for the associated eclogites, the equilibrium pressure of these rocks is not uniquely known. If the Liavatn marbles did equilibrate at significantly lower pressures

than the adjacent eclogites, then the accompanying fluid would have been much more  $\text{CO}_2$  rich. However, the presence of abundant epidote in the marbles indicates high  $X_{\text{H}_2\text{O}}$  conditions, implying that the marbles may have indeed equilibrated at high pressures.

If sufficient limits can be placed on the  $X_{\text{CO}_2}$  and  $f(\text{O}_2)$  the buffered equilibria



(solid solution in barite or celestine)



(solid solution in calcite)

could then provide upper limits on  $\text{SO}_2$ ,  $\text{SO}_3$ , and sulphur fugacity ( $f_{\text{S}_2}$ ), once the  $a/X$  relations of Ca–Ba–Sr sulphates and carbonates have been established.

## DISCUSSION

### Conditions of metamorphism

Marbles and calc-silicates from the BAS area subparallel Caledonian shear zones and appear to be thoroughly affected by the Caledonian deformation and metamorphism. Their intimate occurrence with eclogites and the jadeite content of the clinopyroxene indicates that the marbles experienced the same high-pressure metamorphism. The marbles may even have formed from the precipitation from fluids generated during the high-pressure metamorphism and metasomatism of the granulites. For comparison, using the same thermobarometers employed by Boundy *et al.* (1992), the adjacent eclogite from the Liavatn locality yields temperatures of 670–720 °C and minimum pressures of 13 kbar. These estimates are similar to the temperatures and pressures estimated in Boundy *et al.* (1992) ( $T = 670$  °C,  $P_{\text{min}} = 15$  kbar) for eclogites in the NW Holsnøy Island area. The data from the Liavatn calc-silicate sample with garnet and omphacite yields temperatures of 580–650 °C. These lower temperatures are consistent with those estimated from C isotopes, which are interpreted to represent retrograde temperatures of approximately 600 °C related to the influx of fluids (Van Wyck *et al.*, 1996). Marbles, calc-silicates and some of the associated eclogites record lower temperatures than peak high-pressure eclogite facies temperatures in the area, which may indicate some retrogression to lower temperatures.

### Fluid evolution during high-pressure metamorphism

Fluid released during subduction is an important catalyst for metamorphism, as documented in the BAS (e.g. Austrheim, 1987; Jamtveit *et al.*, 1990). The

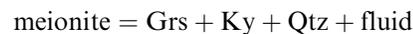
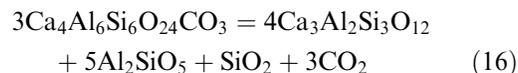
chemistry of such fluids could be continuously modified by dehydration and hydration processes, as fluids are released and consumed during metamorphic reactions at depth (e.g. Svensen *et al.*, 2001). Eclogite formation in the area resulted from the localized infiltration of fluids along fractures and shear zones (e.g. Austrheim, 1987). This may not have been a single pervasive invasion of a single fluid. From a study of a mafic eclogite from Sjurholm, Svensen *et al.* (1999) suggested that infiltrating H<sub>2</sub>O-rich brines caused eclogitization and that their composition was modified during hydration. Phase equilibrium data from the carbonates support that the marbles and calc-silicates were in equilibrium with a CO<sub>2</sub>-poor fluid during eclogite facies metamorphism under relatively oxidizing conditions ( $X_{\text{CO}_2} < 0.08$  at 20 kbar). The results presented here are consistent with previous investigations of phase equilibria of the eclogites from Holsnøy Island, which estimate H<sub>2</sub>O dominated fluid compositions ( $\geq 0.75$ ; Jamtveit *et al.*, 1990). The fluids in equilibrium with the carbonates does not appear to be a (water-rich) brine, as indicated by Svensen *et al.* (1999), suggesting the variable nature of the fluids during high pressure eclogite metamorphism of the area. Different composition fluids may have infiltrated the area at depth and the fluid compositions evolved during high-pressure metamorphic reactions progressed. As hydrous fluids are consumed during high-pressure eclogite metamorphism, the remaining evolved fluids could become more saline (Svensen *et al.*, 1999).

Evaluation of similar marbles from other areas that experienced high-pressure metamorphism also suggests H<sub>2</sub>O dominated fluids (e.g. Franz & Spear, 1983; Wang & Liou, 1993). For example, in marbles of high pressure and ultrahigh pressure metamorphism from the Dabie Shan, the  $X_{\text{CO}_2} < 0.03$  (Wang & Liou, 1993) and Su-Lu terrane the  $X_{\text{CO}_2} = 0.001$  (Kato *et al.*, 1997). High-pressure calc-silicates from Bavaria yield  $X_{\text{CO}_2} \leq 0.03$  (Klemd *et al.*, 1994). Very low  $X_{\text{CO}_2}$  values have also been reported for high-pressure eclogite facies metamorphic rocks in the Alps (e.g. Holland, 1979; Franz & Spear, 1983; Castelli, 1991). Explanations for the very low  $X_{\text{CO}_2}$  values may include: (1) lack of progress for decarbonation reactions during subduction to about 100 km; (2) dominance of dehydration reactions producing substantial free H<sub>2</sub>O, which are common during subduction-zone metamorphism (e.g. Peacock, 1992); and (3) almost complete fractionation of the CO<sub>2</sub> component at high pressure into the solid phases (Molina & Poli, 2000).

The degree of external vs. internal buffering of the fluid is variable depending upon fluid–rock ratios. Despite the influx of free fluids in the rocks at high pressure, the isotopic heterogeneities are best reconciled with local buffering and limited mobility of the fluids, at least during the final stage of equilibration of the Holsnøy eclogites (Boundy *et al.*, 1997b). Limited fluid migration has also been suggested by Svensen

*et al.* (2001). They interpreted multiphase brine inclusions in omphacite to reflect enrichment of elements incompatible with the silicate structure during consumption of H<sub>2</sub>O by eclogite-forming hydration reactions. Similarly, data from high-pressure eclogites from subduction zones supports limited fluid/rock ratios (Selverstone *et al.*, 1992; Getty & Selverstone, 1994). Because most dewatering and devolatilization during tectonic burial occurs prior to eclogite facies metamorphism, the volumetric fluid–rock ratio in eclogites should be generally low at subduction depths of > 45 km (Getty & Selverstone, 1994).

For the Holsnøy Island localities, it is at present not known whether the CO<sub>2</sub>-poor fluid compositions reflect local buffering or a regional external source. Eclogites formed as a result of deformation and infiltration of fluids into relatively anhydrous Precambrian granulite facies rocks (e.g. Austrheim & Griffin, 1985; Austrheim, 1987; Austrheim *et al.*, 1997). This implies a fluid mobility on scales at least exceeding those of the precursors. The lack of nearby sources of extensive hydrous phases suggests the fluid source (for the initial retrogression) is not local. Breakdown of scapolite in the surrounding granulites during high-pressure eclogite facies metamorphism, however, would imply a local, limited CO<sub>2</sub> source for the marbles, and could be controlled by reactions such as:



(Moecher & Essene, 1990).

#### Origin of the marbles: metasomatic interactions with fluids?

Currently a definitive origin of the marbles remains controversial, yet the data available suggest some intriguing possibilities. Marbles from the Liavtna area have calcite with average  $\delta^{18}\text{O}$  of  $9.5 \pm 0.6\text{‰}$  and  $\delta^{13}\text{C}$  of  $-4 \pm 0.8\text{‰}$  and are in equilibrium with other silicate minerals at temperatures of about 600 °C (van Wyck *et al.*, 1996). The authors interpreted the carbon isotope signature of the fluid as being best modelled as derived from dewatered sediments. The Sr isotopes, while variable, are more characteristic of a continental crust rather than mantle component. The variability in Sr isotopes may well reflect metasomatism and/or retrograde metamorphism of the marbles from initial crustal values. If the metasomatic breakdown of minerals in granulites, such as scapolite, contributed to the formation of the eclogite facies marbles and calc-silicates this could account for some of the variability and apparent crustal isotopic signatures. The infiltration of fluids along shear zones can dramatically effect the <sup>87</sup>Sr/<sup>86</sup>Sr ratios. McCaig *et al.* (2000) documented infiltration of brine rich fluids rich in Sr can significantly effect the <sup>87</sup>Sr/<sup>86</sup>Sr composition of carbonate rocks and accounts

for much of the variability in the isotopic ratios. Metasomatic skarns in the Grenville have significantly elevated Sr contents and the partitioning Sr between scapolite and calcite could potentially significantly modify the Sr isotopic system (Shaw *et al.*, 1960).

A metasomatic origin of the marbles and calc-silicates is plausible based on the field relations and is also consistent with the petrological observations. The marbles and calc-silicates occur along or subparallel to Caledonian shear zones. The shear zones served as channel ways for infiltrating hydrous fluids during eclogite facies metamorphism and localized retrogression (Austrheim, 1987; Jamtveit *et al.*, 1990). Some of the surrounding granulites, particularly with anorthositic precursors, contain scapolite. Textural evidence indicates carbonates and Ca-Ba-Sr sulphates forming during breakdown of  $\text{SO}_4\text{-CO}_2$  scapolite in granulites during eclogite facies metamorphism. These observations suggest that the formation of minor carbonates in the eclogites (*sensu stricto*) from Holsnøy resulted from the breakdown of scapolite in the surrounding granulites. One intriguing possibility is that this breakdown of scapolite in the granulites during high-pressure metamorphism could have contributed to the formation of the localized marbles and calc-silicates along the shear zones. This would imply that the marbles themselves are a product of metasomatic reactions and precipitated from a fluid along shear zones during the high-pressure metamorphism.

#### Eclogite facies carbonates from the Lindås Nappe: insights into metamorphism of the subducted crust

Large intact areas of crust on Earth are currently being subducted to depths of 50 km or more. Continental crust is currently being subducted to great depths beneath the Himalayas, particularly evident in the Pamir region in NW Himalayas (Burtman & Molnar, 1993), depths at which high-pressure and ultrahigh pressure metamorphism is expected to occur. Exposures of the arrested stages of eclogite metamorphism in the mafic rocks and now the discovery of eclogite facies marbles from Holsnøy from the Lindås Nappe in the Caledonian Orogen of western Norway provides clues to metamorphic processes and possible fluid evolution as the crust is subducted and recycled. The Lindås Nappe anorthosite complex and related rocks represent a slice of lowermost crust that was subducted to high-pressure eclogite facies conditions. However, the timing of high-pressure metamorphism is directly related to the infiltration and access of fluids in the subducted crust (Austrheim & Envik, 1997; Bingen *et al.*, 2001). The marbles reflect the same high-pressure metamorphic evolution and retain well-preserved eclogitic assemblages. In contrast to some of the mafic eclogites in the area, which equilibrated with evolved saline fluids (Svensen *et al.*, 1999, 2001), fluids in equilibrium with the marbles and calc-silicates appear to be  $\text{H}_2\text{O}$ -rich oxidizing fluids under eclogite facies

conditions. The marbles and related calc-silicates may represent the reaction breakdown products of the  $\text{SO}_4\text{-CO}_2$  scapolites in the granulites during high-pressure metamorphism. The Sr-rich sulphates and carbonate products could then have been concentrated along Caledonian shear zones in relations to the infiltration of fluids. Further quantification of such a metasomatic origin of the eclogite facies carbonates is the subject on continued investigations.

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